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How hydrothermal synthesis improves the synthesis of (Zr,Ce)SiO₄ solid solutions

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ABSTRACT:

Although ZrSiO_4 is the most well-known compound in the zircon-structured family (space group $I4_1/amd$), the experimental conditions for preparing pure and well-crystallized phases which are doped with a tetravalent element via hydrothermal synthesis has never been clearly discussed in the literature. With the aim to answer this question, the experimental conditions of preparation of ZrSiO_4 and $(\text{Zr,Ce})\text{SiO}_4$ were investigated in order to synthesize well-crystallized and pure phases. A multiparametric study has been carried out using soft hydrothermal conditions, with variables including reactant concentration, initial pH of the reactive medium, and duration of the hydrothermal treatment. Pure ZrSiO_4 was obtained through hydrothermal treatment for 7 days at 250°C , within a large acidity range ($1.0 \leq \text{pH} \leq 9.0$) and starting from $C_{\text{Si}} \approx C_{\text{Zr}} \geq 0.2 \text{ M}$. As hydrothermally prepared zircon structured phases can be both hydrated and hydroxylated, its annealed form was also studied after heating to 1000°C . Based on these results, the synthesis of $(\text{Zr,Ce})\text{SiO}_4$ solid solutions were also investigated. The optimal hydrothermal conditions to acquire pure and crystallized phases were obtained in 7 days at 250°C with initial $\text{pH} = 1$ and concentration of the reactants equal to $0.2 \text{ mol}\cdot\text{L}^{-1}$. This led to $\text{Zr}_{1-x}\text{Ce}_x\text{SiO}_4$ solid solutions with the incorporated Ce content up to 40 mol.%. Samples were characterized by multiple methods, including lab and synchrotron PXRD, IR and Raman spectroscopies, SEM, and TGA. Moreover, it was found that these phases were thermally stable in air up to at least 1000°C .

KEYWORDS: zircon, silicate, wet chemistry route, hydrothermal synthesis

1. INTRODUCTION

Although the hydrothermal synthesis of the zircon group (space group $I4_1/amd$) silicates has long been known, with the first hydrothermal synthesis of $ZrSiO_4$ carried out by Von Chrustchoff in 1892 [1], their preparation continues to arouse a great interest, in particular the preparation solid solutions. Nature has long derived zircon structured silicates, in the form of minerals, which exhibit a plethora of solid solutions in both two crystallographic sites. The compositional end-members of these minerals include, zircon ($ZrSiO_4$), hafnon ($HfSiO_4$), stetindite ($CeSiO_4$), thorite ($ThSiO_4$), and coffinite ($USiO_4$), which can form either in primary igneous (i.e., magmatic) or secondary hydrothermal (i.e., epithermal) depositional environments. Elements which can form zircon-type solid solution in the dodecahedral site (M-site) tend to be high-field strength elements, including group III (Sc and Y), group IV (Ti, Zr, and Hf), lanthanides (Ln, La – Lu), and early actinides (An, Ac – Pu). Anthropogenic conditions have also led to the formation of solid solutions with a zircon-type structure, with $(Zr,An)SiO_4$, being observed in corium lavas coming from Chernobyl nuclear accident [2], [3]. Moreover, as a part of Fukushima Daichi nuclear core contained MOx fuels, the presence of $(Zr,Pu)SiO_4$ phases in the corium lavas has been suggested [4]. Understanding properties of such solid solutions is thus crucial for predicting the behavior of corium in the case of severe accident as well as for the determination of remediation methods in such accidental situations. On the other hand, because of the highly refractory character of zircon-type materials to dissolution, the use of $(Zr,An)SiO_4$ solid solutions has been considered as a potential radwaste matrix for the long-term storage of actinides [5]–[7]. A complete description of zircon-based solid solutions is thus mandatory to improve our understanding of how zircon phases form in lavas during a nuclear accident or how they behave in the long-term storage of dismantled nuclear weapons in a geological repository [6], [7].

While the synthesis of pure silicate end-members of the zircon group $MSiO_4$ (with M = Zr, Ce, Hf, Th, Pa, U, Np, Pu, Am) [1], [8]–[13], and a few solid solutions, including $(Zr,Hf)SiO_4$ [14]–[16][17] and $(Th,U)SiO_4$ [18]–[22], a significant knowledge gap still remains in the scientific literature on the preparation of several solid solution systems. This is demonstrated by that synthetic $(Zr,An)SiO_4$ and $(Zr,Ce)SiO_4$ solid solutions appears to be much more difficult [23]–[26][27]. Indeed, the syntheses by dry chemical processes did not allow to produce solid solutions with high concentration An/Ce incorporation rates in the zircon host matrix. Thus, the solubility limits of actinides in the zircon structure using solid-state synthesis were found to be low (7–10 wt.%, 7 ± 1 wt.%; 4 ± 1 wt.% and 2 ± 1 wt.% for Pu, Ce, U and

Th, respectively) [23], [28], [29]. However, such dry route syntheses did not allow the formation of solid solutions as rich as that formed in the corium lavas for (Zr,U)SiO₄ (up to 12.9 wt.%) [2]. Finally, simulations performed on the thermodynamic properties of such solid solutions have shown that actinide-enriched solid solutions are enthalpically metastable [26], [30], [31], which however, contradicts to experimental calorimetry [21], [22].

Several studies were recently carried out by Estevenon *et al.* on the hydrothermal synthesis of zircon-type silicates (HfSiO₄, CeSiO₄, ThSiO₄, and PuSiO₄) and contributed to major advances in the understanding of these synthetic routes [9]–[12], [32], [33]. The aim of this work is to take advantage of this recent knowledge in order to develop the synthesis of (Zr,Ce)SiO₄ solid solutions with higher concentration (mol.%) of Ce than already reported [23], [24]. This solid solution system was selected on the basis that Ce is often viewed as a non-radioactive analog of Pu due to their similarity in physico-chemical properties (i.e., ionic radii and hydrolysis constants) [34]–[37]. Therefore, through understanding the synthetic conditions of (Zr,Ce)SiO₄ solid solution one can gain important insight towards synthesizing (Zr,Pu)SiO₄ solid solution with high mol.% Pu, further allowing for the derivation of thermodynamic parameters.

2. MATERIALS AND METHODS

2.1. Syntheses

2.1.1. Preparation of stetindite CeSiO₄

Stetindite (CeSiO₄) sample was synthesized by hydrothermal method from Ce(III)-silicate solid precursor (A-Ce₂Si₂O₇), slightly adapting the protocol recently developed by Estevenon *et al.* [9]. A mixture containing 1.5 g of CeO₂ (Sigma Aldrich, particle size < 5 μm) and 0.571 g of SiO₂ (Sigma Aldrich, 10-20 nm) was mechanically milled (30 Hz, 1 hour) in a Retsch MM 200 vibration mill mixer apparatus using a tungsten carbide milling jar. In this mixture, a silicate excess of 9 mol.% was thus considered to ensure the formation of the silicate phase. This mixture was pelletized by uniaxial pressing (500 MPa) at room temperature and then heated at 1350 °C for 9 hours under reducing atmosphere (Ar–4% H₂) to form the Ce(III)-silicate solid precursor, A-Ce₂Si₂O₇ (space group: *P*4₁). In a 23 mL Teflon lined Parr reactor, 200 mg of A-Ce₂Si₂O₇ was put in contact with 4 mL of 0.75 mol·L⁻¹ HNO₃ solution (prepared by dilution of ACS grade 70% HNO₃, Sigma Aldrich). Moreover, an additional excess of 8 mol.% of silicate was introduced to the solution as Na₂SiO₃·5H₂O (8.8 mg) in order to avoid the formation of CeO₂. The pH of the solution was then adjusted to 6.5 ± 0.1 with a freshly prepared 8 mol·L⁻¹ NaOH solution (from ACS grade NaOH pellets, Sigma Aldrich). This mixture was then hydrothermally treated for 7 days at 150°C under air atmosphere using a Parr autoclave (Parr 4749 acid digestion vessel). After cooling, the final solid was separated from the supernatant solution by centrifugation at 12 000 rpm, washed twice with deionized water and once with ethanol and finally dried overnight in an oven at 60°C.

2.1.2. Preparation of zircon : ZrSiO₄

The preparation of zircon was inspired from the protocols described by Valero [38] and by Estevenon *et al.* for the HfSiO₄ synthesis [10], [11]. Na₂SiO₃·5H₂O (Sigma-Aldrich, ≥ 95%) and ZrO(NO₃)₂·2H₂O (Sigma-Aldrich, ≥ 99%) were mixed together with a 5 mol.% silicate excess and dissolved in 1 mol·L⁻¹ HNO₃ solution (prepared by dilution of Sigma Aldrich ACS grade solution of 70% HNO₃). The pH was then adjusted to the expected value thanks to a freshly prepared 8 mol·L⁻¹ NaOH solution (obtained from Sigma Aldrich ACS grade 98 % NaOH).

All of the prepared mixtures were introduced in 23 mL Teflon lined containers. The containers were put in Parr autoclaves (Parr 4749) then placed in an oven to reach hydrothermal

conditions at 250°C under autogenous pressure for varying holding times. The final cooling to room temperature was done within one hour. Then, the precipitates were separated from the supernatant by centrifugation for 12 min at 12 000 rpm, washed twice with deionized water and once with ethanol, and finally dried overnight in an oven at 60°C.

2.1.3. Preparation of $Zr_{1-x}Ce_xSiO_4$ solid solutions

All of the reagents used were supplied by Sigma-Aldrich. $Na_2SiO_3 \cdot 5H_2O$ ($\geq 95\%$), $ZrO(NO_3)_2 \cdot 2H_2O$ ($\geq 99\%$) and $(NH_4)_2Ce(NO_3)_6$ ($\geq 98.5\%$) were used as aqueous silicate, zirconium and cerium precursors respectively. A $1 \text{ mol} \cdot \text{L}^{-1}$ HNO_3 solution was prepared by dilution of Sigma Aldrich ACS grade solution of 70% HNO_3 . A $8 \text{ mol} \cdot \text{L}^{-1}$ $NaOH$ solution was freshly prepared from Sigma Aldrich ACS grade $NaOH$ (98 %) before the experiments.

The syntheses of $Zr_{1-x}Ce_xSiO_4$ solid solutions were performed on the basis of the conditions previously determined for each end-members, $CeSiO_4$ and $ZrSiO_4$, described above. Aqueous zirconium, cerium and silicate solutions were prepared by dissolving $ZrO(NO_3)_2 \cdot 2H_2O$, $(NH_4)_2Ce(NO_3)_6$ and $Na_2SiO_3 \cdot 5H_2O$ in $1 \text{ mol} \cdot \text{L}^{-1}$ HNO_3 in order to obtain a reacting mixture with a silicate excess of 5 mol.%, aiming at obtaining Zr and Ce oxide-free samples. The global cation concentration ($C_{Zr(IV)} + C_{Ce(IV)}$) was set to $0.2 \text{ mol} \cdot \text{L}^{-1}$ with varying zirconium and cerium concentrations depending on the targeted composition. The pH of the mixture was then adjusted to about $pH = 1$ with the help of $8 \text{ mol} \cdot \text{L}^{-1}$ $NaOH$.

All of the prepared mixtures were introduced in 23 mL Teflon lined containers (Parr 4749). The conditions of the hydrothermal treatment were fixed to $t = 7$ days and $T = 250^\circ\text{C}$, under autogenous pressure. The final cooling to room temperature was done within one hour. Then, the precipitates were separated from the supernatants by centrifugation at 12 000 rpm for 12 min, washed twice with deionized water and once with ethanol, and finally dried overnight at 60°C in an oven. Some of the obtained powders were finally annealed at 1000°C in a muffle furnace for 10 hours.

2.2. Characterization of the prepared samples

PXRD data were collected on the resulting powders using the Bruker D8 advance diffractometer equipped with a LynxEye detector and using $Cu \text{ K}\alpha$ radiation ($\lambda = 1.54184 \text{ \AA}$) in a reflection geometry (parallel beam). Patterns were recorded between 5° and 90° (2θ) with steps of 0.019° and a total counting time of 2.5 to 3 hours per sample. Pure silicon was used as

a standard material to extract the instrumental function. Then, the collected data were refined by the Rietveld method using the Fullprof suite package [39]. During the refinements, different profile and structure parameters were allowed to vary, such as the zero shift, unit cell parameters, scale factor, and overall displacement factor. However, the occupancy of each site was fixed to the calculated values. In order to facilitate the diffractograms comparison, the background has been subtracted and the intensity divided by the maximum.

In order to ensure the identification of all of the phases present in annealed samples (especially to avoid the misidentification of secondary oxide phases), High Resolution Powder X-Ray Diffraction have been performed at the MARS beamline of the SOLEIL Synchrotron on the dedicated CX2 end-station. The CX2 end-station is equipped with a four-circle diffractometer that we used in transmission mode. Before being collected by a set of 24 scintillation detectors, the X-ray beams diffracted by the samples are highly collimated using a multi-crystal analyzer stage made of two banks of Ge(111) analyzer crystals mounted on the 2θ arm. Diffraction patterns were obtained after the normalization of the diffraction signals coming from all the 24 detectors. This signal normalization was done using a specific procedure developed by the beamline [40]. It is well-known that this geometry associated with a capillary sample holder is one of the most efficient ways to record very high-resolution diffraction patterns. The samples were placed in 0.6 mm Kapton capillary tubes (1.5 cm long). The powder was sealed in the tube by epoxy resin. The detector geometry parameters were calibrated using the NIST standard LaB₆ (SRM 660b). The HR-XRD end-station was set at 17.0 keV ($\lambda = 0.72959 \text{ \AA}$) after calibration at the yttrium K-edge energy ($E = 17,038 \text{ keV}$) with an incident collimated X-ray beam sizing about $300 \times 1000 \text{ \mu m}^2$ ($V \times H$). Samples were spun in the incoming X-ray beam to average a potential texturing effect. XRD patterns were collected in the 0 to 115° 2θ range, with an angular step of 0.002° and a counting time of 0.1 s per frame in a continuous scan mode. Data were treated using the TOPAS software [41]. XRD pattern fitting was done using Le Bail model, considering no sample texture, no strain, and the structure factors given by the atomic standard.

Raman spectra were recorded with a Horiba-Jobin Yvon Aramis device equipped with an edge filter and a Nd:YAG laser (532 nm) that delivered 60 mW at the sample surface. In order to avoid any laser-induced degradation of the compound, the power was turned down by the means of optical filters. The laser beam was then focused on the sample using an Olympus BX 41 microscope with an $\times 50$ LMP objective, resulting in a spot area of $\sim 1 \text{ \mu m}^2$ and a power of 39 mW. For each spectrum, measurements were performed in the $100\text{--}4000 \text{ cm}^{-1}$ range for

dwelt time of 1 to 30 s. Four scans were recorded and averaged for each analyzed area in order to minimize the instrumental error. In order to facilitate their comparison, the spectra were normalized in the [0–1] range.

FTIR spectra were recorded with a Perkin-Elmer FTIR Spectrum 100 device in the 300 – 4000 cm^{-1} range. Powdered samples were deposited on the surface of an ATR crystal without any prior preparation. The spectra collected in such operating conditions exhibited a resolution lower than 4 cm^{-1} . Five scans were performed and merged for each sample to average the instrumental error. In order to facilitate their comparison, the spectra were divided by their maximum intensity.

SEM observations were directly conducted using a FEI Quanta 200 electronic microscope on small powdered samples without any prior preparation such as metallization. The electronic microscope was equipped either with an Everhart-Thornley Detector (ETD) or a Back-Scattered Electron Detector (BSED), under high vacuum conditions with a low accelerating voltage (2 kV). These conditions were chosen in order to create a beam deceleration effect that allowed the collection of high resolution images.

Thermogravimetric analyses measurements were conducted on Mettler Toledo TGA/DSC 3+ apparatus. After recording a baseline using an empty alumina crucible, about 20 mg of powdered sample were placed into another alumina crucible of the same type, then weight loss was measured during a heat treatment up to 1000 °C with a rate of 5°C·min⁻¹ under air. The hydration level of the prepared samples prior calcination was thus determined.

3. RESULTS AND DISCUSSION

3.1. Optimization of the zircon ZrSiO_4 preparation

Even if the hydrothermal synthesis of ZrSiO_4 was often mentioned in the literature [29], [38], [42]–[49], [50]–[54], there are still an important details concerning the description of the optimal experimental conditions which are often lacking. Overall, the formation of ZrSiO_4 under hydrothermal conditions, even if always successful, is known to be slow, leading to poorly crystallized samples. Additionally, the impact of crucial parameters known to affect the synthesis of chemical analogues MSiO_4 (M = Hf, Ce, Th, U, Pu), such as the initial pH, the reagent concentrations and the hydrothermal treatment duration has never been evaluated. Temperature was fixed to 250°C as, according to literature [10], [11], [38], [42], [44], its only role was to increase the reaction kinetics.

3.1.1. Effect of the pH of the starting solution

Considering a starting mixture with Zr concentration of $0.2 \text{ mol}\cdot\text{L}^{-1}$ and a Si/Zr molar ratio of 1.05, the effect of the pH of the starting mixture on the nature of the resulting precipitate was followed between $1.0 \text{ mol}\cdot\text{L}^{-1} \text{ HNO}_3$ and $\text{pH}_{\text{ini}} = 11.7$. Hydrothermal treatments were performed at 250°C for 20 days in order to limit the constraints related to the crystallization kinetics of ZrSiO_4 . The whole set of the samples prepared over the studied initial pH range were characterized by PXRD (**Figure 1**).

From these results, ZrSiO_4 was formed on a wide range of acidity, from a starting mixture in $1.0 \text{ mol}\cdot\text{L}^{-1} \text{ HNO}_3$ solution up to an initial pH value of about 10.0. A small amount of ZrO_2 (few wt.%, space group $P2_1/c$) was quantified only for the samples prepared with $\text{pH}_{\text{ini}} = 2.0$ and 3.1 (syntheses (3) and (4)). Nevertheless, for higher pH values (syntheses (11) and (12)), the precipitation of an unidentified phase was also evidenced. Contrary to the experiments performed by Valero [38], the formation of ZrSiO_4 was observed for an initial pH close to neutrality. However, the prepared silicates exhibit a lower crystallinity than those obtained from higher or lower pH values. Comparing our results with those obtained by Valero, it is likely that nanoscale zircon was obtained in his study rather than an amorphous phase when pH values close to neutrality.

Unit cell parameters were determined by Rietveld refinements of PXRD patterns for all the ZrSiO_4 samples prepared between $1.0 \text{ mol}\cdot\text{L}^{-1} \text{ HNO}_3$ and $\text{pH}_{\text{ini}} = 9.0$. The obtained values did not allow observing any significant variation of the unit cell parameters with respect to the initial pH of the reactive media. Indeed, the unit cell volume reached a maximum value of $262.7(7) \text{ \AA}^3$ ($a = 6.628(5) \text{ \AA}$ and $c = 5.980(6) \text{ \AA}$) for $\text{pH}_{\text{ini}} = 3.1$ and a minimum value of $261.1(7) \text{ \AA}^3$ ($a = 6.612(5) \text{ \AA}$ and $c = 5.973(6) \text{ \AA}$) for $1.0 \text{ mol}\cdot\text{L}^{-1} \text{ HNO}_3$ (**Table SI 2**). No real trend was observed between these two extrema. Such unit cell volumes were found to be slightly bigger than the reference value obtained on sample prepared by dry chemistry route at high temperature (i.e. 260.8 \AA^3 ($a = 6.604 \text{ \AA}$ and $c = 5.979 \text{ \AA}$) [42]. According to recent results published on zircon-based compounds, this difference could be due to the insertion of molecular water or hydroxide groups in the zircon structure [10], [11], [55]. It may be correlated to non-ideal crystallization in the soft hydrothermal conditions considered.

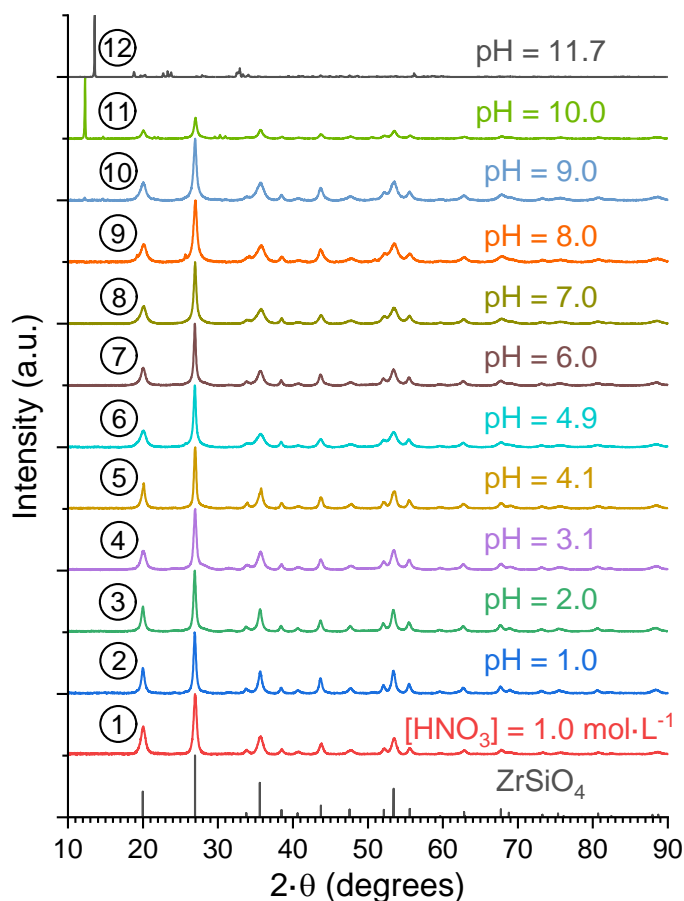


Figure 1. PXRD patterns obtained for ZrSiO_4 samples prepared under hydrothermal conditions (20 days, $T = 250^\circ\text{C}$) starting with zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ and for various initial pH values: $1.0 \text{ mol}\cdot\text{L}^{-1} \text{ HNO}_3$ (1), pH = 1.0 (2), pH = 2.0 (3), pH = 3.1 (4), pH = 4.1 (5), pH = 4.9 (6), pH = 6.0 (7), pH = 7.0 (8), pH = 8.0 (9), pH = 9.0 (10), pH = 10.0 (11) and pH = 11.7 (12).

3.1.2. Effect of the hydrothermal treatment duration

In order to observe the impact of the hydrothermal treatment duration on the formation of ZrSiO_4 , several experiments were also performed at 250°C with an initial pH value equal to 1 and a zirconium concentration of $0.2 \text{ mol}\cdot\text{L}^{-1}$. Holding times were fixed to 1, 7 and 20 days. From PXRD analyses (**Figure 2**), it is clear that single phase ZrSiO_4 was always formed showing that the duration of the hydrothermal treatment did not affect the stability of the obtained phase. However, the longer the hydrothermal treatment duration, the better the crystallinity of the final product. This later point supports the data reported in the literature from which the kinetics of crystallization of ZrSiO_4 appears to be slower compared to that of HfSiO_4 or ThSiO_4 [10], [11].

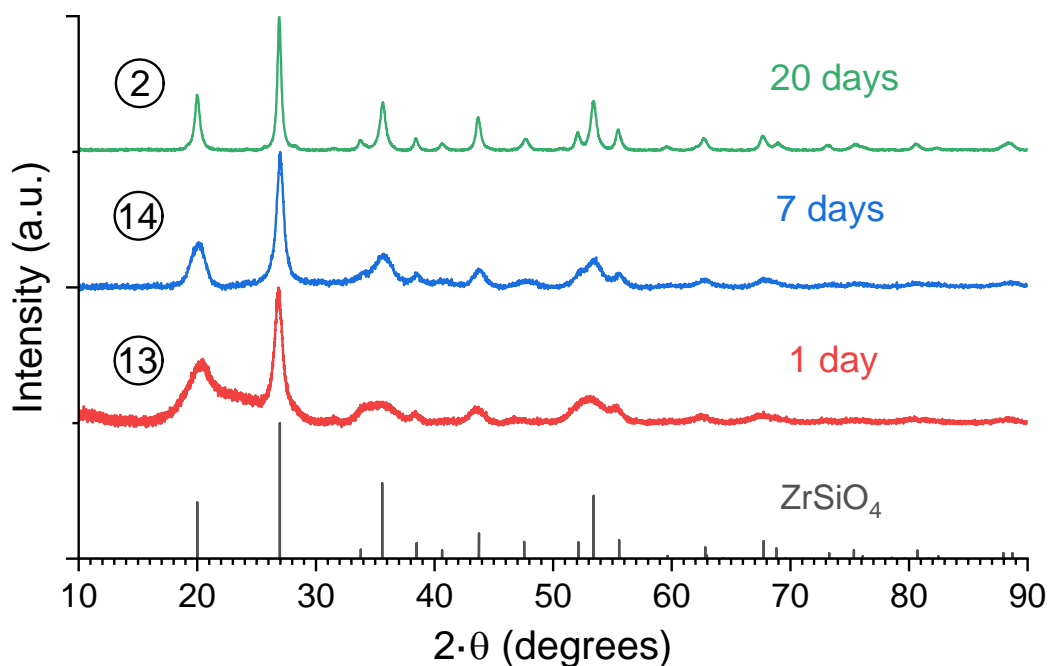


Figure 2. PXRD patterns obtained for ZrSiO_4 samples prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, $\text{pH} = 1$) starting with zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ and for holding time of 1 day (13), 7 days (14) and 20 days (2).

Moreover, based on Rietveld refinements (**Figure 3**), extension of the duration of the hydrothermal treatment led to a small, but noticeable, change of the unit cell volume and crystallite size. Indeed, the unit cell volume and the crystallite size reached $V = 260.9(2) \text{ \AA}^3$ ($a = 6.611(3) \text{ \AA}$ and $c = 5.9711(4) \text{ \AA}$) and $16 \pm 1 \text{ nm}$, respectively, after 1 day compared to $V = 262.52(2) \text{ \AA}^3$ ($a = 6.6265(2) \text{ \AA}$ and $c = 5.9784(3) \text{ \AA}$) and $24 \pm 1 \text{ nm}$ after 20 days.

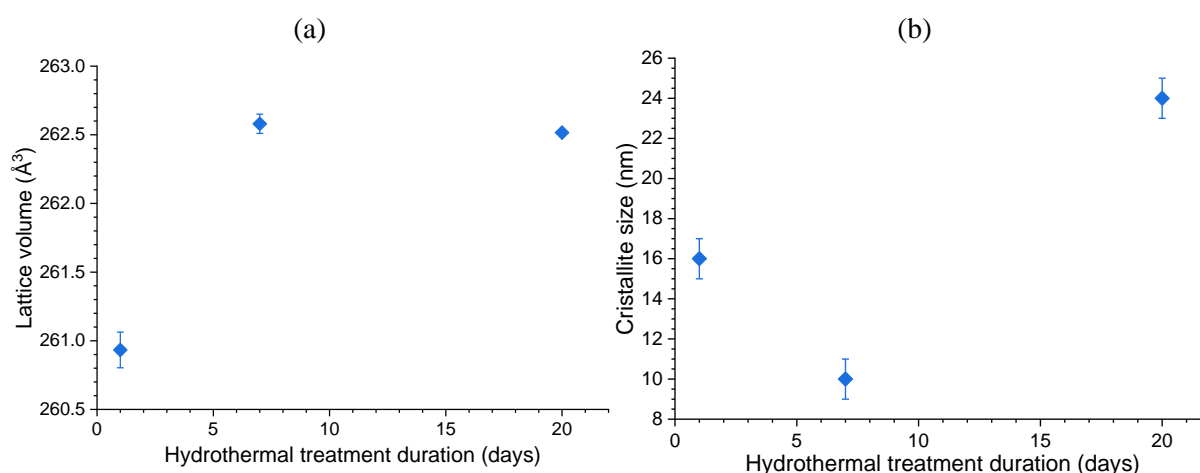


Figure 3. Unit cell volume (a) and average crystallite size (b) determined by Rietveld refinement for ZrSiO_4 samples prepared under hydrothermal conditions ($T = 250^\circ\text{C}$) starting with zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ and at $\text{pH}_{\text{ini}} = 1$ with hydrothermal treatment duration of 1 day (13), 7 days (14) and 20 days (2).

3.1.3. Effect of the starting zirconium and silicate concentrations

Since the reagent concentrations were crucial parameters for the synthesis of a number of zircon-group compounds, it was also suspected that the concentrations of zirconium and silicate in the starting mixture affected the formation of ZrSiO_4 by varying the saturation index of the solution with respect to zircon. With this objective, several syntheses were performed under hydrothermal conditions with various starting zirconium and silicate concentrations. The impact of the zirconium and silicate concentrations was followed between $0.05 \text{ mol}\cdot\text{L}^{-1}$ and $1.0 \text{ mol}\cdot\text{L}^{-1}$, keeping constant the Si/Zr molar ratio to 1.05 the initial pH value to 1, and the conditions of the hydrothermal treatment ($T = 250 \text{ }^\circ\text{C}$, 7 days). From the results summarized in **Figure 4**, pure ZrSiO_4 was obtained for $C_{\text{Zr}} \geq 0.2 \text{ mol}\cdot\text{L}^{-1}$ whereas the formation of ZrO_2 was observed as a byproduct for lower concentration such as $C_{\text{Zr}} = 0.05 \text{ mol}\cdot\text{L}^{-1}$. From these observations, one can conclude that silicate concentration in solution was not sufficient to fully counterbalance the hydrolysis of zirconium in the less concentrated media. Therefore, the formation of ZrSiO_4 was favored for $C_{\text{Zr}} \geq 0.2 \text{ mol}\cdot\text{L}^{-1}$. Moreover, the starting concentration significantly affected the crystallization state of the final sample (**Figure 5**). Especially, decreasing the zirconium and silicate ions concentration in the reactive medium induced an increase of the crystallite size of ZrSiO_4 (from 7 to 14 nm when increasing the starting concentration from $1.0 \text{ mol}\cdot\text{L}^{-1}$ to $0.05 \text{ mol}\cdot\text{L}^{-1}$). This behavior is similar to what has been mentioned in the literature for ThSiO_4 and HfSiO_4 [10], [11]. It may be attributed to the faster kinetics of nucleation in the more concentrated reactive media, which led to favor nucleation process instead of crystallite growth. Simultaneously, the unit cell volume increased significantly for higher starting concentrations in the reactive medium. Indeed, it reached $V = 265.5(2) \text{ \AA}^3$ ($a = 6.654(1) \text{ \AA}$ and $c = 5.996(2) \text{ \AA}$) for starting concentrations of $1.0 \text{ mol}\cdot\text{L}^{-1}$. This result might be explained by a higher amount of water/hydroxide groups inserted in the zircon phase for the syntheses performed for high zirconium and silicate starting concentrations.

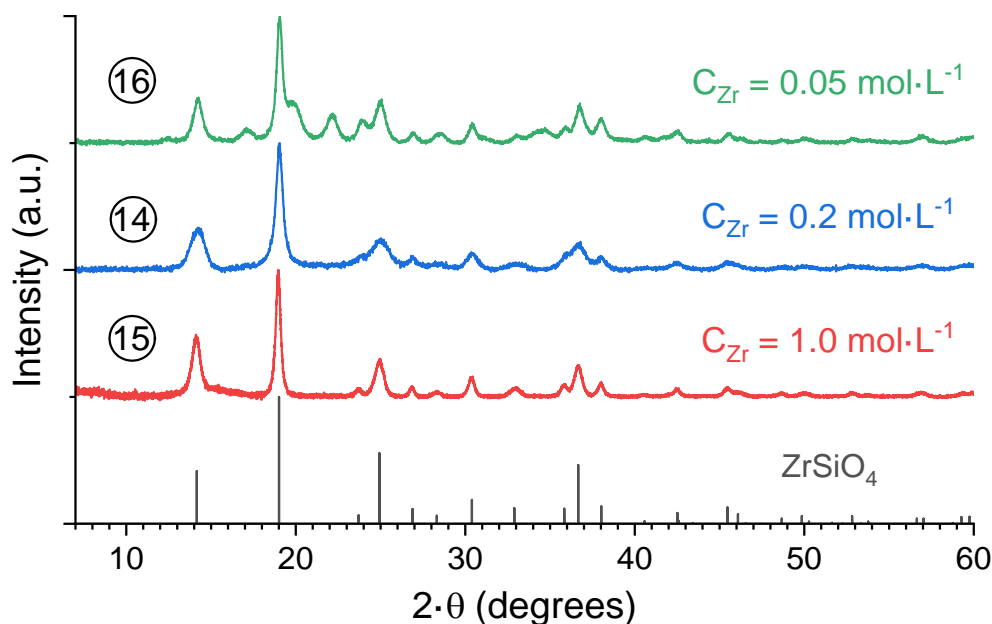


Figure 4. PXRD patterns obtained for ZrSiO_4 samples prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days) at $\text{pH} = 1.0$ starting with zirconium and silicate concentrations of $1.0 \text{ mol}\cdot\text{L}^{-1}$ (15), $0.2 \text{ mol}\cdot\text{L}^{-1}$ (14) and $0.05 \text{ mol}\cdot\text{L}^{-1}$ (16).

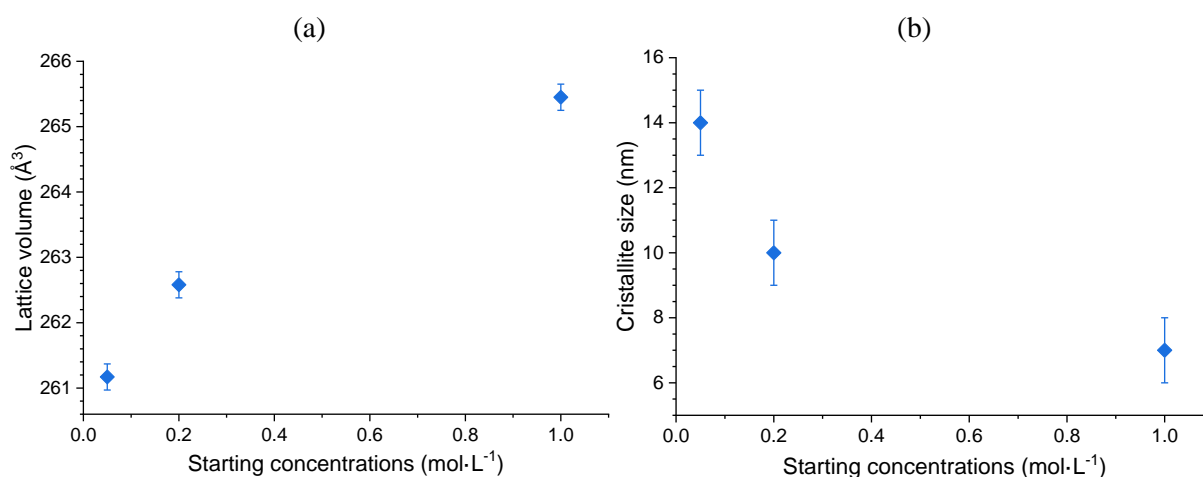


Figure 5. Unit cell volume (a) and average crystallite size (b) determined by Rietveld refinement for ZrSiO_4 samples prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1$) starting with zirconium and silicate concentrations of $1.0 \text{ mol}\cdot\text{L}^{-1}$ (15), $0.2 \text{ mol}\cdot\text{L}^{-1}$ (14) and $0.05 \text{ mol}\cdot\text{L}^{-1}$ (16).

3.1.4. Additional characterization of annealed ZrSiO_4 sample

Natural zircons are known to be highly hydrated and hydroxylated [43], [56], like several other zircon-group compounds such as coffinite USiO_4 [55], [57], [58]. It therefore appeared important to characterize ZrSiO_4 samples in the pristine state (washed and dried from hydrothermal synthesis) and after annealing at 1000°C . The characterized ZrSiO_4 sample was

prepared under hydrothermal conditions at 250°C for 7 days, at pH = 1 and starting from zirconium and silicate concentrations of 0.2 mol·L⁻¹.

First, SEM micrographs were recorded for both ZrSiO₄ samples. None of the zircon-type samples exhibited a well-defined morphology. It was composed of very fine nanoparticles gathered in bigger aggregates of few micrometers in size (**Figure SI 1**). These observations were in good agreement with what was reported in the literature for ZrSiO₄, although the morphology of this phase appeared to be highly dependent on the precursors used as well as on the initial pH of the starting mixture [42], [59], [60].

TGA analysis (**Figure 6**) revealed that ZrSiO₄ was thermally stable up to 1000 °C. The TGA profile obtained was comparable to the data available in the literature for synthetic zircon specimens despite different synthetic conditions [43], [56]. Due to the high total mass loss measured, i.e. 10.5 wt.%, the sample was found to be highly hydrated and hydroxylated. A first mass loss (4.5 wt.%) was observed around 100°C. It was assigned to the elimination of water molecules adsorbed on the surface of the sample. A second mass loss (≈ 4 wt.%) was observed up to about 480 °C. It could be explained by the significant release of the “structural” water confined into the [001] channel, more strongly interacting with the structure of the sample, as suggested by Strzelecki *et al.* [55], [61]. Then, the additional mass loss (≈ 2 wt.%) observed up to 1000 °C, may be due to the continuous release of the most strongly bonded water molecules among with the hydroxyl groups. These results agreed well with the results recently obtained by Guo *et al.* [62] for TGA analyzes of synthetic coffinite USiO₄ samples and by Estevenon *et al.* [10], [11] for HfSiO₄ and ThSiO₄ samples. However, it is known that hydration of synthetic or natural ZrSiO₄ samples is erratic and almost always higher than for other silicates of the zircon-group [43], [56], [63].

PXRD patterns of pristine and annealed ZrSiO₄ sample confirmed the thermal stability up to 1000°C (**Figure 6**). The zircon-type structure (space group *I4₁/amd*) was preserved and no trace of oxide formation such as ZrO₂ was evidenced. The unit cell volume decreased from $V = 262.6(1) \text{ \AA}^3$ ($a = 6.622(1) \text{ \AA}$ and $c = 5.988(1) \text{ \AA}$) to $V = 260.8(1) \text{ \AA}^3$ ($a = 6.5995(5) \text{ \AA}$ and $c = 5.987(1) \text{ \AA}$) after the thermal treatment, again evidencing the release of the confined water molecules and hydroxyl groups. It was also noted that the heat treatment allowed the growth of crystallites, from size of $10 \pm 1 \text{ nm}$ to $28 \pm 1 \text{ nm}$.

Pristine and annealed samples were characterized by SEM. However, no characteristic change of morphology was observed (**Figure SI 1**).

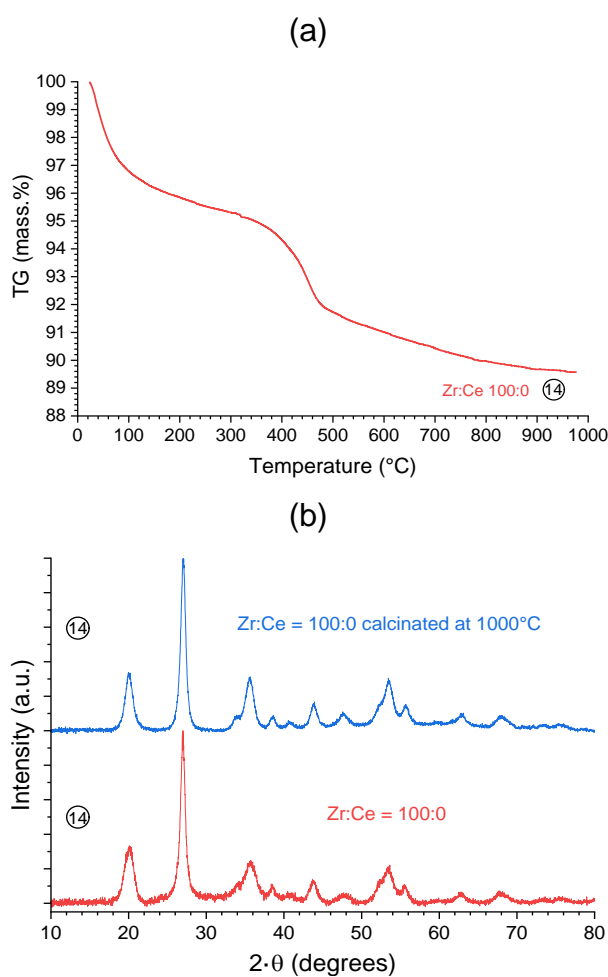


Figure 6. TGA analysis (a) and PXRD patterns (b) recorded for a pristine and annealed (after TGA, 1000°C) ZrSiO₄ sample (14) prepared under hydrothermal conditions (T = 250°C, 7 days) at pH = 1.0 starting with zirconium and silicate concentrations of 0.2 mol·L⁻¹.

The analysis by IR and Raman spectroscopies (**Figure 7**) recorded for pristine and annealed ZrSiO₄ samples confirmed the previous interpretations. Regarding the Raman spectra, seven vibration bands were evidenced in addition to vibrational bands observed below 300 cm⁻¹, which were attributed to external modes. Then, four vibration bands located near 440 cm⁻¹, 610 cm⁻¹, 970 cm⁻¹ and 1000 cm⁻¹ belonged to the SiO₄ internal modes ν_2 , ν_4 , ν_1 and ν_3 , respectively. Now focusing on the IR spectra of the pristine zircon sample, the water bending mode band, even with a low intensity, was observed near 1550 cm⁻¹. Moreover, stretching mode of hydroxyl groups, near 2930 cm⁻¹, was clearly visible in the spectrum. Finally, a wide band ranging between 3700 – 3100 cm⁻¹ was assigned to the stretching mode band of hydroxide. Thus, the presence of these three vibration bands supported the hydrated and hydroxylated character of ZrSiO₄, as often mentioned in the literature for several zircon-type compounds

[43], [55], [56], [64]. Especially, annealing step at 1000 °C led to the disappearance of these bands, showing the full dehydration and dehydroxylation of the zircon sample. This result agreed well with that suggested from TGA analysis.

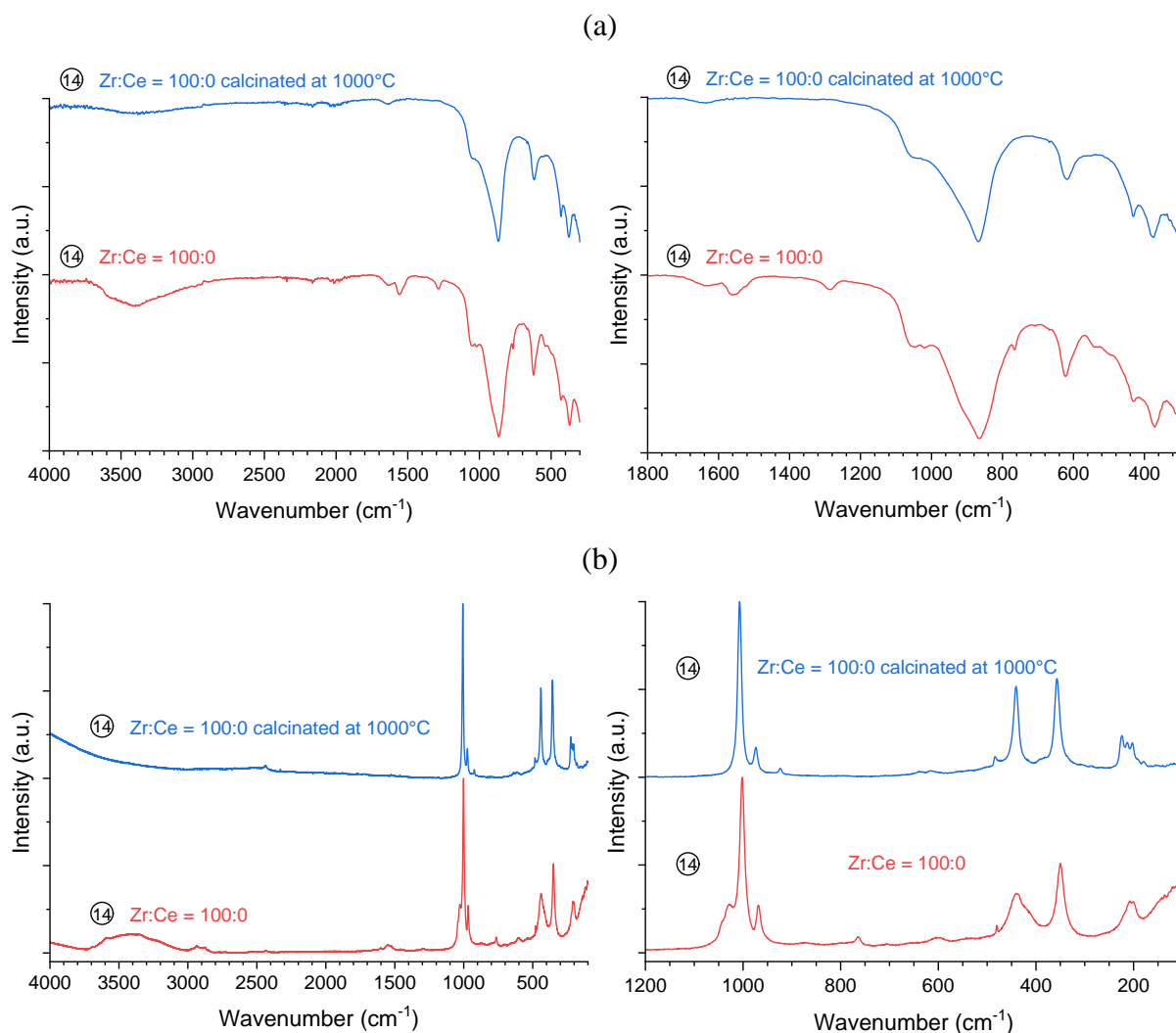


Figure 7. Infrared (a) and Raman (b) spectra recorded in the 4000 – 300 cm^{-1} and 4000 – 100 cm^{-1} , respectively, for pristine (14) and annealed ZrSiO_4 sample prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1.0$) starting with zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$.

3.2 Hydrothermal synthesis of $(\text{Zr,Ce})\text{SiO}_4$ solid solutions

The experimental optimal conditions considered for the synthesis of $(\text{Zr,Ce})\text{SiO}_4$ solid solutions were fixed on the basis of those determined for both ZrSiO_4 and CeSiO_4 end-members. Syntheses were thus performed through hydrothermal treatment at 250°C for 7 days starting with silicate and global $\text{Zr} + \text{Ce}$ concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ in nitric acid ($\text{pH} = 1$).

Such conditions appeared to be the best compromise to prepare oxide-free and relatively well-crystallized silicates.

3.2.1. PXRD characterization

Results of PXRD analysis showed the formation of $(\text{Zr,Ce})\text{SiO}_4$ solid solutions for Ce incorporation concentrations ranging from 5 mol.% to about 40 mol.% (**Figure 8**). Indeed, PXRD patterns revealed the presence of all XRD lines associated to the tetragonal $I4_1/amd$ zircon-type structure, without any extra peak. Moreover, a progressive shift towards the low angles was observed when increasing the Ce concentration up to 50 mol.%. With sample of Ce 50 mol.% showing secondary phase (CeO_2), XRD results suggested that the solubility limit of Ce in the ZrSiO_4 matrix is above 40 mol.%, which is significantly higher than that reported on synthetic specimens prepared by dry chemistry route [23], [24]. Furthermore, a low additional shift was obtained between the samples containing 50 mol.% and 60 mol.% of Ce. It suggested that the limit of cerium incorporation was obtained in that range using this synthesis route. This limit is found to be much higher than those reported in the literature using dry routes of synthesis. Considering that the Ce end-member formation is not possible under such synthesis conditions (see the multiparametric synthesis of CeSiO_4 in [33]), this also underlines the important role of zirconium in the co-precipitation of Ce(IV) as a silicate compound.

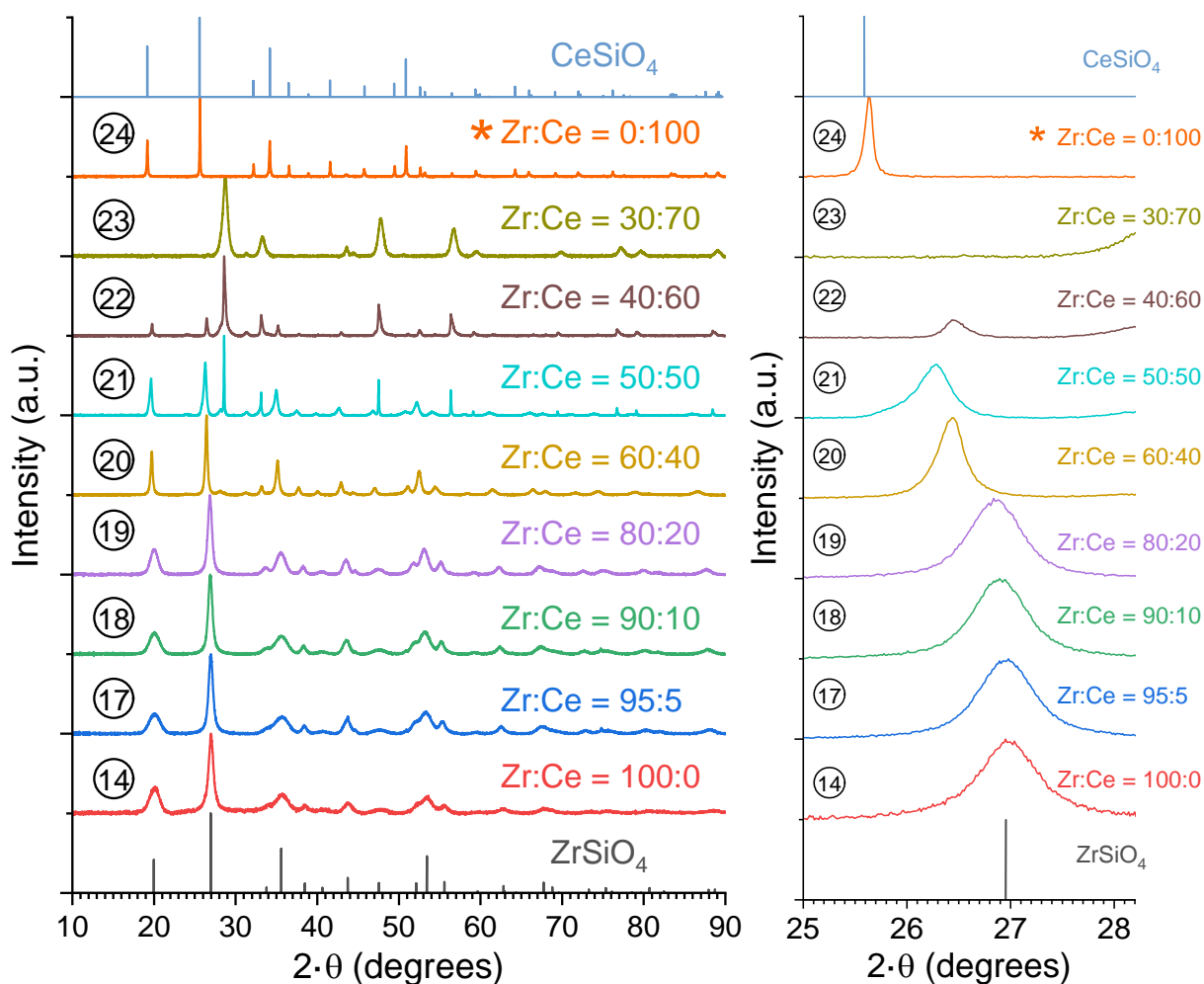


Figure 8. PXRD patterns recorded for pristine $(\text{Zr,Ce})\text{SiO}_4$ solid solutions with various chemical compositions prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1.0$) starting with Zr + Ce and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ for Zr:Ce = 100:0 (14), 95:5 (17), 90:10 (18), 80:20 (19), 60:40 (20), 50:50 (21), 40:60 (22) and 30:70 (23). * Reference CeSiO_4 sample ($T = 150^\circ\text{C}$, 7 days, $\text{pH} = 6.5$) starting with cerium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ (24).

Since ZrSiO_4 dehydration process appeared to be progressive up to temperature of 1000°C (**Figure 6a**), samples calcination have been carried out at 800°C and 1000°C in order to obtain $(\text{Zr,Ce})\text{SiO}_4$ water-free samples and to compare the impact of the water loss on the silicate structure. Rietveld refinements were thereafter carried out on all pristine solid solutions (i.e. from 5 to 60 mol.% of Ce) as well as on samples annealed at 800°C and 1000°C to study their thermal stability along with the effect of dehydration on the values of unit cell parameters (**Table 1** and **Figure 9**). Independently of hydration, a and c unit cell parameters were found to be very close to that expected values if ideal solid solutions was assumed. Indeed, the unit cell volume increased almost linearly as a function of the Ce molar content, following closely Vegard's law, which is in good agreement with the progressive substitution of Zr by Ce within

the structure. The small deviation from Vegard's law could be explained by the presence of ZrO_2 or/and CeO_2 in the mixture with silicate, affecting the Ce content of the solid solution, or non-ideal volume of mixing.

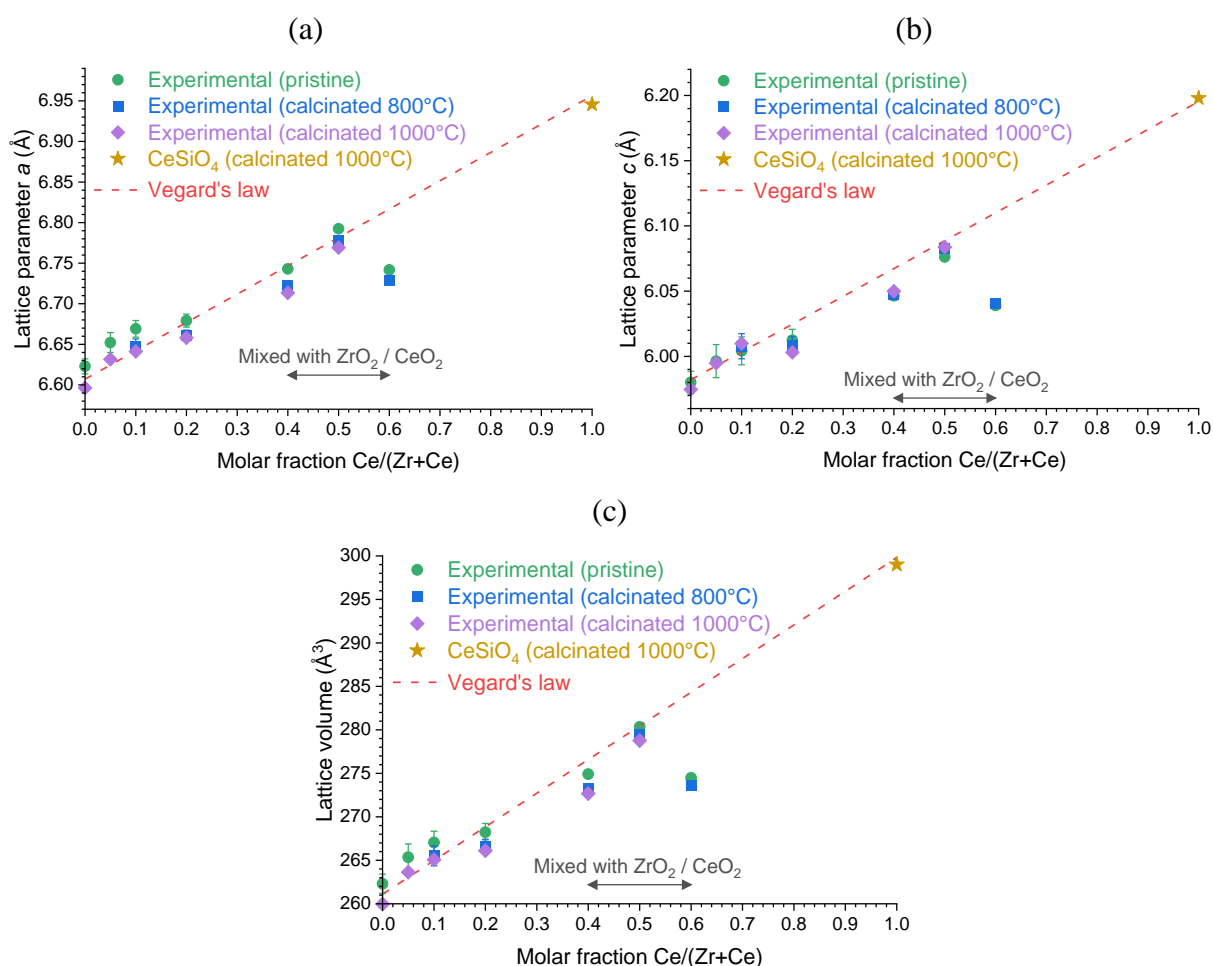


Figure 9. Unit cell parameters a (a) and c (b) and volume (c) obtained by Rietveld refinements performed from PXRD patterns of pristine $(\text{Zr,Ce})\text{SiO}_4$ solid solutions and annealed samples at 800°C and 1000°C . * Reference CeSiO_4 sample wasn't obtained in the same chemical conditions as the $(\text{Zr,Ce})\text{SiO}_4$ samples.

Regarding to the annealing of the pristine samples, a small decrease of the unit cell volume was first observed. Dehydration led to the contraction of the lattice predominantly in the direction perpendicular to $[001]$ (*i.e.* a and b axes). Indeed, water molecules are expected to be located in the channels parallel to $[001]$, which could explain that the c axis was not significantly affected [55], [57]. A similar trend has been observed between 800°C and 1000°C and allows to confirm that the mass loss at high temperature could still be explained by water molecule elimination. The solid solutions were found to be thermally stable up to at least

1000°C. In order to improve the detection of secondary phases and the determination of unit cell parameters, the samples annealed at 1000°C were characterized by high resolution PXRD at the MARS beamline of SOLEIL facility (**Figure 10**). The measurement were performed at just below the yttrium K-edge energy ($E = 17.0$ keV, $\lambda = 0.72959$ Å), in order to facilitate the direct comparison with PXRD patterns obtained at the Lab's scale performed at the copper K_{α} radiation ($\lambda = 1.54184$ Å), all of the PXRD patterns have been plotted in scattering vector, according to the Bragg's law, in supporting information (**Figures SI 2 to SI 7**). Single-phase solid solutions were obtained up to 20 mol.% of Ce. For higher Ce incorporation content, the presence of ZrO_2 as secondary phase was confirmed after annealing at 1000°C (**Figure 10**).

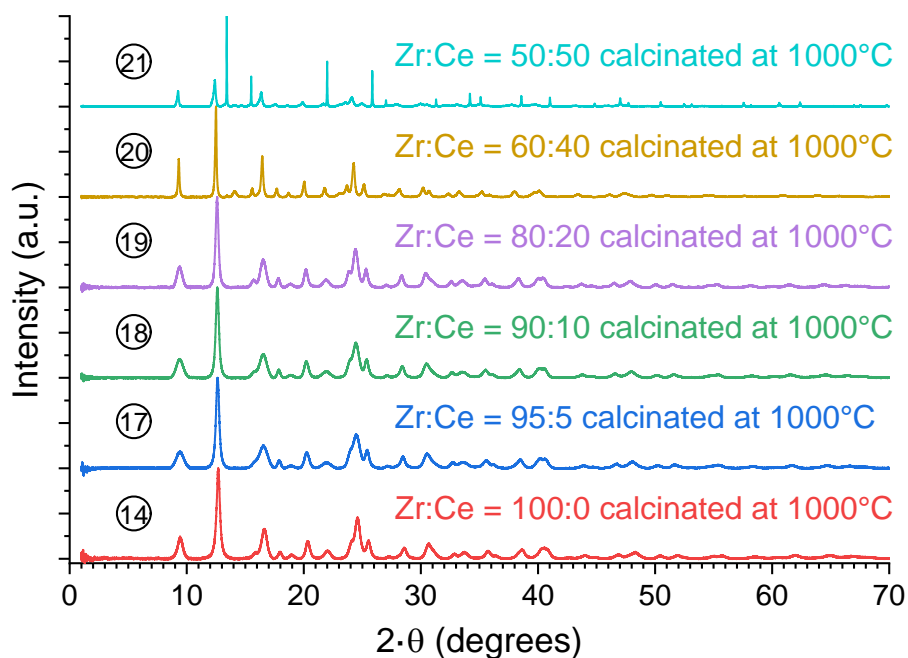


Figure 10. Synchrotron PXRD patterns recorded for $(Zr,Ce)SiO_4$ solid solutions annealed at 1000°C with Zr:Ce = 100:0 (14), Zr:Ce = 95:5 (17), Zr:Ce = 90:10 (18), Zr:Ce = 80:20 (19), Zr:Ce = 60:40 (20) and Zr:Ce = 50:50 (21) prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1.0$) starting with Zr + Ce and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$.

Table 1 Unit cell parameters and volume obtained by Rietveld refinements performed from PXRD patterns of pristine (Zr,Ce)SiO₄ solid solutions and annealed samples at 800°C or 1000°C.

	$\frac{\text{Ce}}{\text{Zr} + \text{Ce}}$	Pristine sample			Annealed at 800°C			Annealed at 1000°C		
		<i>a</i> (Å)	<i>c</i> (Å)	<i>V</i> (Å ³)	<i>a</i> (Å)	<i>c</i> (Å)	<i>V</i> (Å ³)	<i>a</i> (Å)	<i>c</i> (Å)	<i>V</i> (Å ³)
(14)	0	6.623(9)	5.980(9)	262.3(1)	-	-	-	6.6046(3)	5.9976(4)	261.62(2)
(17)	0.05	6.652(12)	5.996(13)	265.3(15)	-	-	-	6.6335(7)	6.0127(6)	264.58(7)
(18)	0.1	6.669(10)	6.004(11)	267.1(13)	6.648(9)	6.008(10)	265.5(11)	6.6451(4)	6.0239(5)	266.00(3)
(19)	0.2	6.679(8)	6.012(8)	268.2(10)	6.661(7)	6.008(7)	266.6(8)	6.6642(3)	6.0200(4)	267.36(3)
(20)	0.4	6.743(3)	6.046(3)	274.9(4)	6.723(2)	6.048(2)	273.3(3)	6.7188(2)	6.0581(3)	273.48(2)
(21)	0.5	6.792(1)	6.076(1)	280.3(1)	6.778(1)	6.083(2)	279.4(2)	6.7594(5)	6.0985(7)	278.64(4)
(22)	0.6	6.742(2)	6.039(3)	274.4(3)	6.729(2)	6.041(3)	273.6(3)	-	-	-
(24)	1	6.9564(3)	6.1953(4)	299.8(1)	-	-	-	6.9452(2)	6.1978(2)	298.95(2)

3.2.2. Raman and IR characterization

All the prepared samples were characterized by Raman and IR spectroscopies (**Figure 11**). The features associated to the SiO₄ modes were observed on the spectra. The symmetric ν_1 and antisymmetric ν_3 stretching modes were located near to 970 and 1000 cm⁻¹ and to 870 and 1070 cm⁻¹ by Raman and IR spectroscopies, respectively. Moreover, the ν_2 symmetric bending mode was observed by both Raman and IR around 440 cm⁻¹ while the ν_4 antisymmetric mode was only clearly distinguishable around 620 cm⁻¹ in the IR spectra.

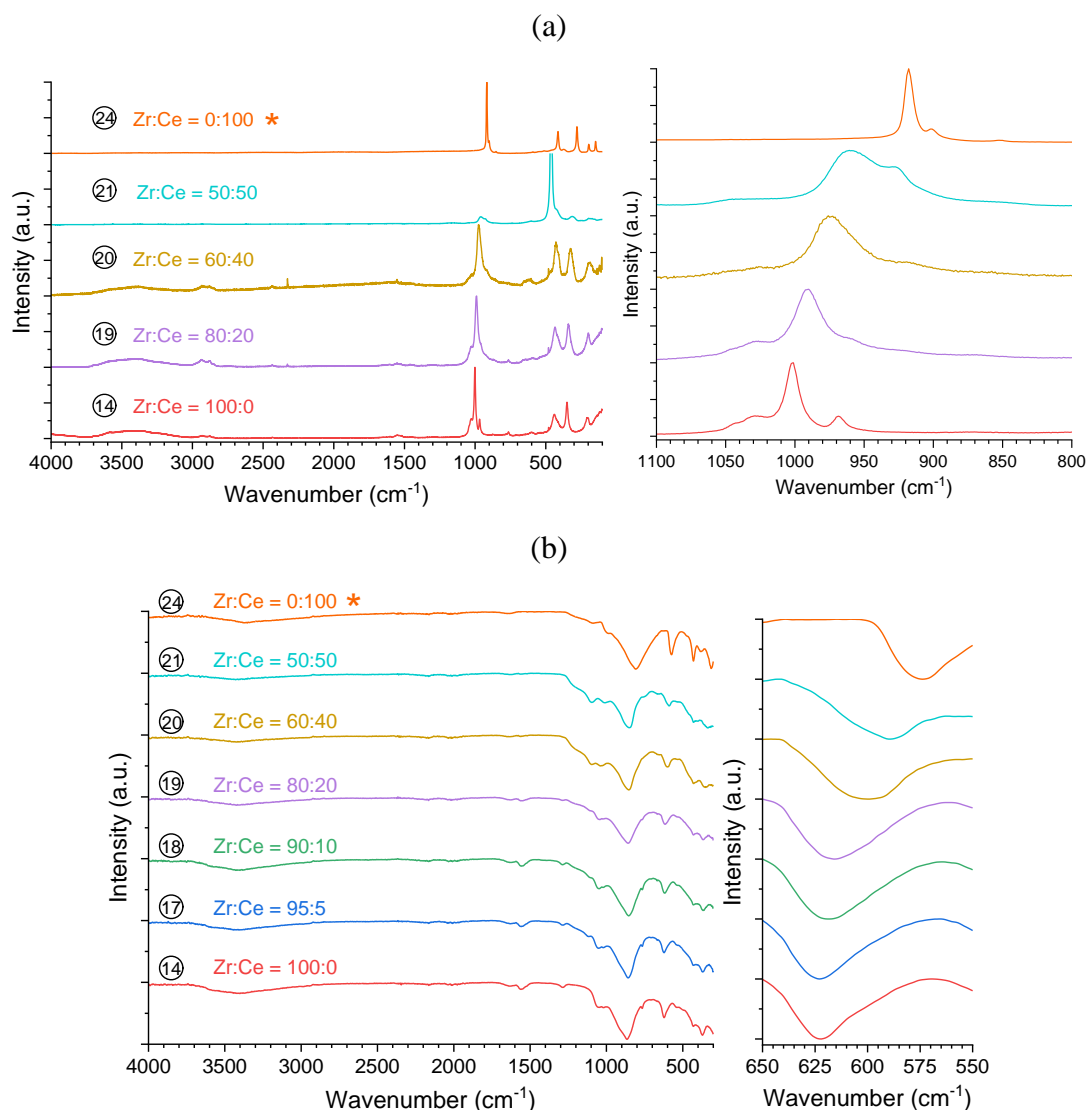


Figure 11. Raman (a) and IR (b) spectra recorded for pristine (Zr,Ce)SiO₄ solid solutions with Zr:Ce = 100:0 (14), Zr:Ce = 95:5 (17), Zr:Ce = 90:10 (18), Zr:Ce = 80:20 (19), Zr:Ce = 60:40 (20), Zr:Ce = 50:50 (21) prepared under hydrothermal conditions (T = 250°C, 7 days, pH = 1.0) starting with Zr + Ce and silicate concentrations of 0.2 mol·L⁻¹. * Reference CeSiO₄ sample (T = 150°C, 7 days, pH = 6.5) starting with cerium and silicate concentrations of 0.2 mol·L⁻¹ (24).

Additionally, as previously observed for ZrSiO₄, Raman and IR spectra suggested the presence of remaining water and/or hydroxyl groups into the pristine samples through the observation of associated weak bands (**Figure 11**). All these bands did not appear anymore after annealing at 1000°C, showing the dehydration/dehydroxylation of the heated samples (**Figure 12**).

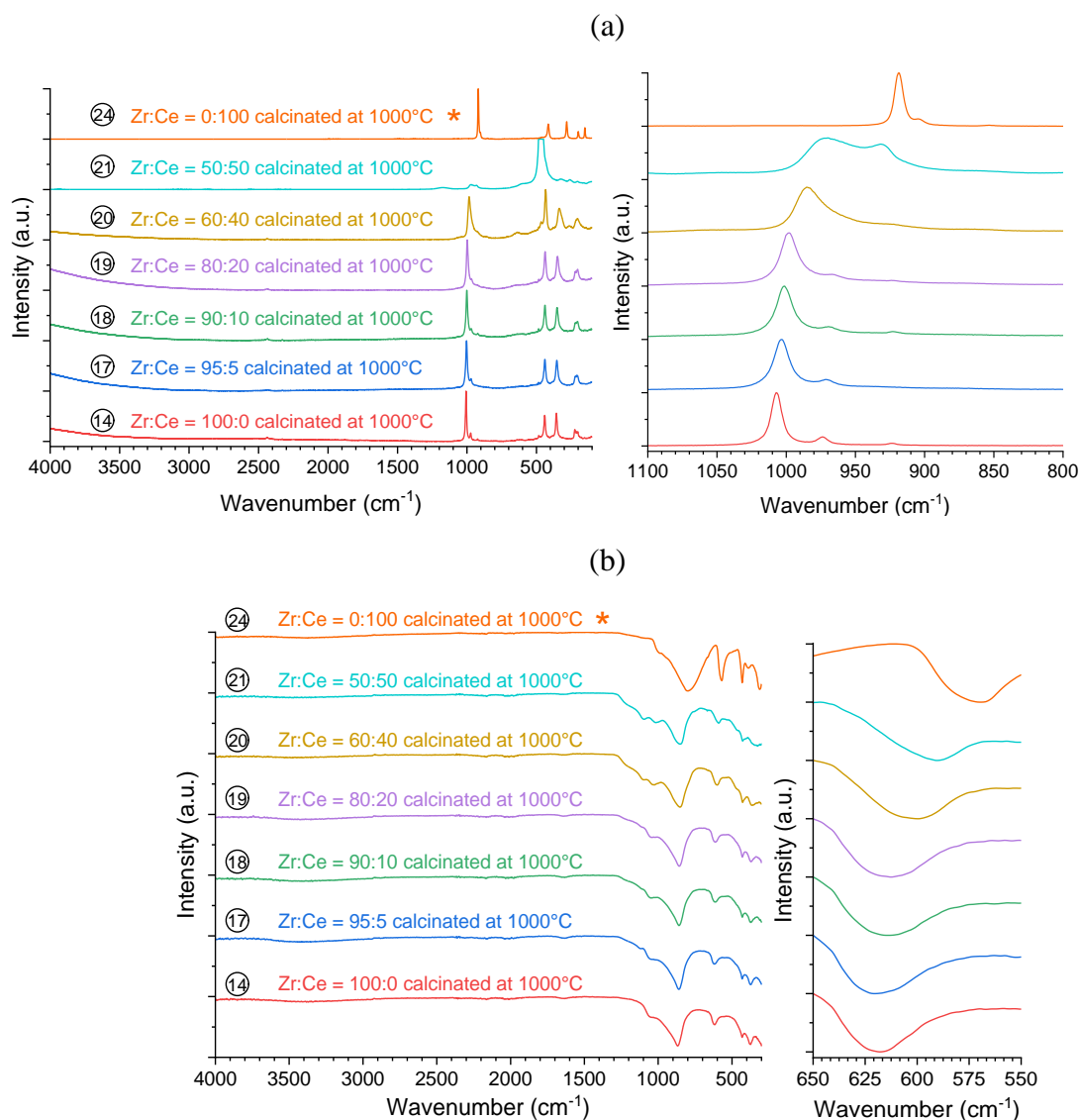


Figure 12. Raman (a) and IR (b) spectra recorded for annealed (Zr,Ce)SiO₄ solid solutions with Zr:Ce = 100:0 (14), Zr:Ce = 95:5 (17), Zr:Ce = 90:10 (18), Zr:Ce = 80:20 (19), Zr:Ce = 60:40 (20), Zr:Ce = 50:50 (21) prepared under hydrothermal conditions (T = 250°C, 7 days, pH = 1.0) starting with Zr + Ce and silicate concentrations of 0.2 mol·L⁻¹. * Reference CeSiO₄ sample (T = 150°C, 7 days, pH = 6.5) starting with cerium and silicate concentrations of 0.2 mol·L⁻¹ (24).

The variations of the vibration bands observed by Raman (ν_2 and ν_3) and IR (ν_1 and ν_4) were plotted for samples annealed at 1000°C versus the chemical composition (**Figure 13** for ν_3 and ν_4 , **Figure SI 8** for ν_1 and ν_2). The obtained results are in good agreement with those from PXRD. Indeed, as observed by PXRD, both ν_3 and ν_4 variations followed an almost linear trend which suggested the cerium incorporation increases in ZrSiO₄.

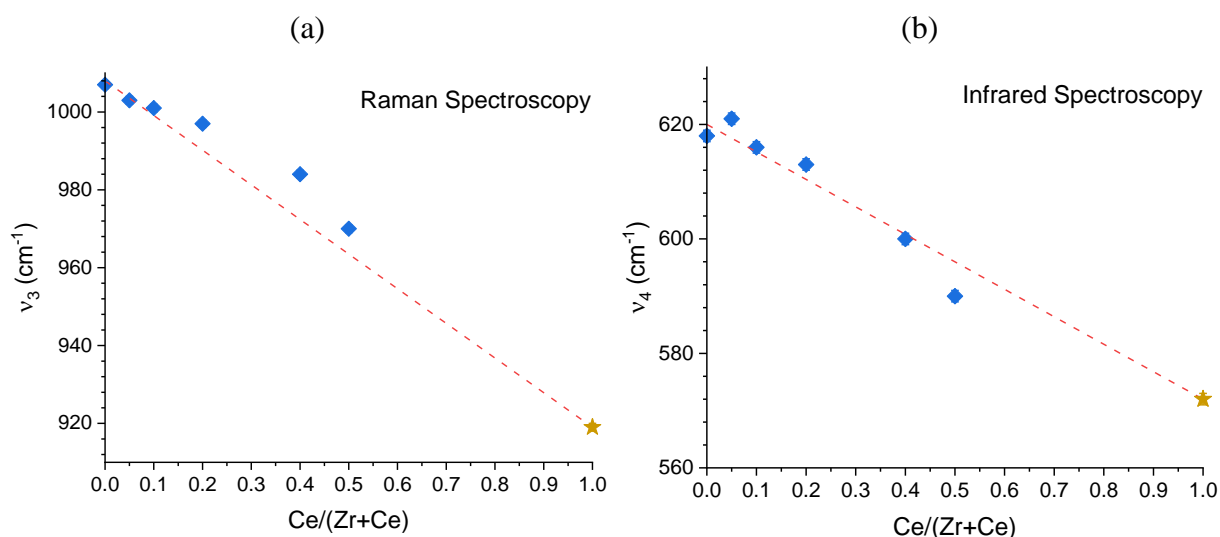


Figure 13. Variation of the ν_3 (a) and ν_4 (b) observed by Raman and IR spectroscopy for $(\text{Zr,Ce})\text{SiO}_4$ (annealed samples at 1000°C) versus the expected chemical composition. * Reference CeSiO_4 sample wasn't obtained in the same chemical conditions as the $(\text{Zr,Ce})\text{SiO}_4$ samples.

3.2.3. Thermogravimetric Analyses

$(\text{Zr,Ce})\text{SiO}_4$ samples were characterized by TGA (**Figure 14**). From these results, it is clear that the weight loss has a negatively correlated dependence on the expected Ce incorporation content. For Ce content below 40 mol.%, TGA curves were close to that recorded for pure ZrSiO_4 (**Figure 6a**). The dehydration/dehydroxylation process occurred in three steps for heating temperatures almost independent of the composition. Moreover, the TGA curve of 50 mol.% of Ce agrees well with that of pure CeSiO_4 . In this case, the dehydration mechanism occurred in only two steps with a weight loss in the same order of magnitude as CeSiO_4 (i.e. ≈ 3.5 wt.%) [65]. This is consistent with the fact that pure ZrSiO_4 is more prone to hydration and hydroxylation than pure CeSiO_4 [55]. Indeed, the equivalent water content of pure hydrothermally synthesized $(\text{Zr,Ce})\text{SiO}_4$ phases (below 20%) might be quantified from the TGA results leading to chemical formulas: $\text{ZrSiO}_4 \cdot 1.2\text{H}_2\text{O}$, $\text{Zr}_{0.95}\text{Ce}_{0.05}\text{SiO}_4 \cdot 1.1\text{H}_2\text{O}$, $\text{Zr}_{0.9}\text{Ce}_{0.1}\text{SiO}_4 \cdot 1.2\text{H}_2\text{O}$ and $\text{Zr}_{0.8}\text{Ce}_{0.2}\text{SiO}_4 \cdot 1.0\text{H}_2\text{O}$ whereas for pure hydrothermally synthesized CeSiO_4 , the equivalent water content measured led to the chemical formula $\text{CeSiO}_4 \cdot 0.43\text{H}_2\text{O}$ [55].

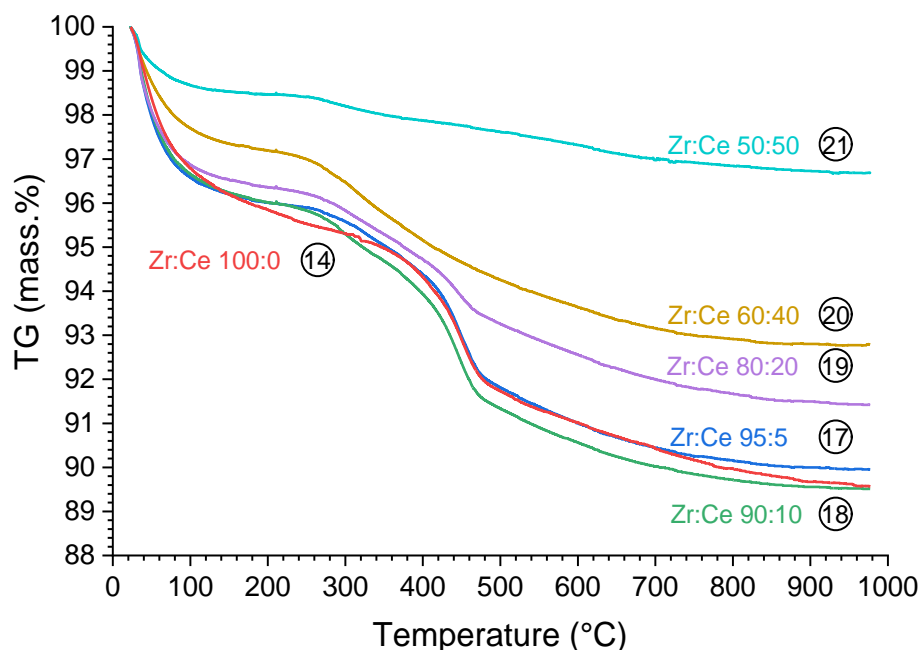


Figure 14. TGA analyses recorded for $(\text{Zr,Ce})\text{SiO}_4$ solid solutions with Zr:Ce = 100:0 (14), Zr:Ce = 95:5 (17), Zr:Ce = 90:10 (18), Zr:Ce = 80:20 (19), Zr:Ce = 60:40 (20) and Zr:Ce = 50:50 (21) prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1.0$) starting with Zr + Ce and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$.

3.2.4. Observations by SEM

SEM micrographs were recorded for several pristine $(\text{Zr,Ce})\text{SiO}_4$ samples (**Figure 15**). The prepared samples did not exhibit a well-defined morphology as they were composed of very fine nanoparticles gathered in bigger aggregates of few micrometers in size, as observed for ZrSiO_4 (**Figure 15a**). On the contrary, CeSiO_4 grains exhibited a well defined morphology with the formation of square-based bipyramids characteristic of this zircon-type materials [9]. Such differences can result from the distinct conditions of synthesis considered for ZrSiO_4 and $(\text{Zr,Ce})\text{SiO}_4$, on the one hand, and CeSiO_4 , on the other hand.

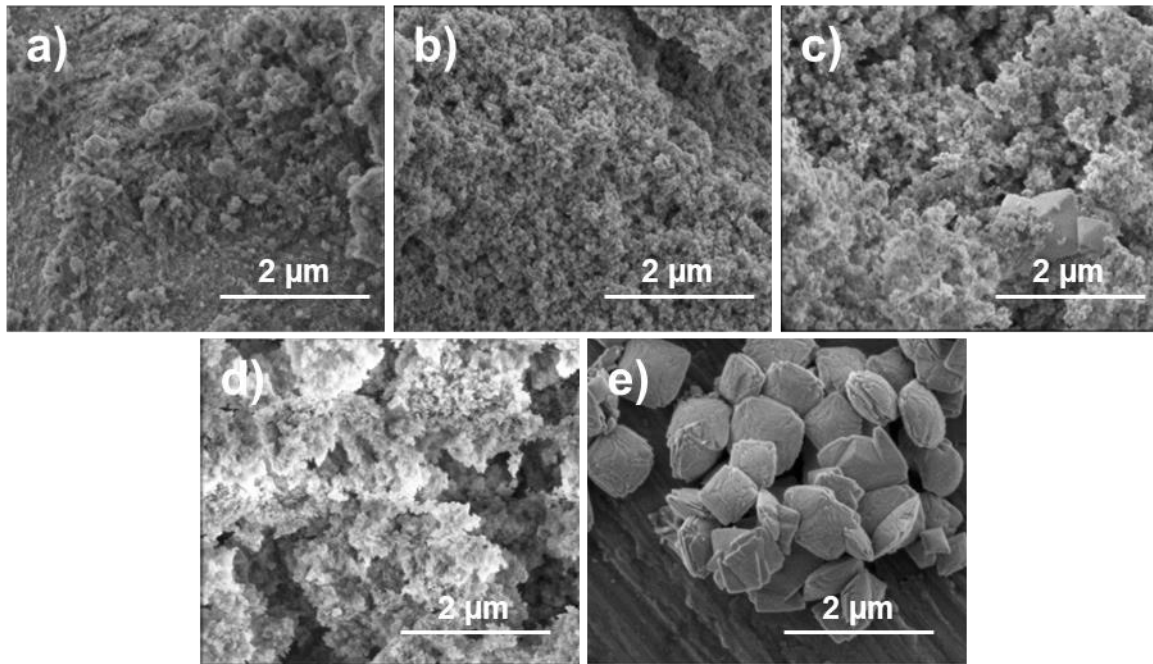


Figure 15. SEM micrographs recorded for $(\text{Zr,Ce})\text{SiO}_4$ solid solutions with **a)** Zr:Ce = 100:0 (14), **b)** Zr:Ce = 60:40 (20), **c)** Zr:Ce = 50:50 (21), **d)** Zr:Ce = 30:70 (23) prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1.0$) starting with Zr + Ce and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$. **e)** Reference CeSiO_4 sample ($T = 150^\circ\text{C}$, 7 days, $\text{pH} = 6.5$) starting with cerium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ (24).

4. CONCLUSION

A multiparametric study of ZrSiO_4 synthesis under hydrothermal conditions has been performed at 250°C in nitric acid media by varying several experimental parameters, including pH, the hydrothermal treatment duration and the zirconium and silicate concentrations in the reactive media. The results allowed us to determine an appropriate set of hydrothermal conditions to prepare single phased samples, which are listed below:

- the initial pH ranged from $[\text{HNO}_3] = 1.0 \text{ mol}\cdot\text{L}^{-1}$ to $\text{pH} = 9.0$;
- the longer is the hydrothermal treatment, the better is the crystallization state and $\Delta t = 7$ days being a good compromise;
- the zirconium and silicate concentration was over $0.2 \text{ mol}\cdot\text{L}^{-1}$ and a Si/Zr molar ratio of 1.05.

ZrSiO_4 samples obtained with a pH close to the neutrality appeared less crystallized than those obtained for higher or lower pH values more likely explained by the formation of nanoscale zircon than an amorphous phase formation. The ZrSiO_4 sample prepared in the

optimal conditions did not exhibit a well-defined morphology, was found to be highly hydrated and hydroxylated with a thermal stability up to 1000°C.

On the basis of the optimal experimental conditions previously determined for ZrSiO₄, the formation of (Zr,Ce)SiO₄ solid solution by hydrothermal treatment was studied. Thus, the synthesis tests were performed with a cation (Zr + Ce) concentration of 0.2 mol·L⁻¹ in a nitric acid media, a pH value equal to 1 and a hydrothermal treatment at 250°C for 7 days. In these conditions, well crystallized and oxide-free solid solutions were obtained up to 20 mol.% of Ce incorporation. Moreover, the experiment performed to reach higher incorporation content led to mixture with ZrO₂ and/or CeO₂ phases but suggested that the solubility limit of Ce in the ZrSiO₄ matrix can be reached almost 40 mol.%, well above the values usually obtained for (Zr,Ce)SiO₄ solid solutions [23], [24]. Solid solutions with higher cerium incorporation rate might be obtained by using chemical route closer to that allowing the formation of the pure CeSiO₄ end-member [9], [33]. However, due to the importance of the cerium redox variation in the CeSiO₄ synthesis process, the Ce substitution by a pure tetravalent reactant, like Zr(IV), could be difficult.

Such as in the case of ZrSiO₄, the (Zr,Ce)SiO₄ samples did not exhibit a well-defined morphology, and were found to be highly hydrated (raw formula of pure (Zr,Ce)SiO₄ phases: ZrSiO₄·1.2H₂O, Zr_{0.95}Ce_{0.05}SiO₄·1.1H₂O, Zr_{0.9}Ce_{0.1}SiO₄·1.2H₂O, Zr_{0.8}Ce_{0.2}SiO₄·1.0H₂O), with dehydrated phases thermally stable up to 1000°C regardless of Ce content.

These results are particularly interesting in the context of studies that could be carried out on (Zr,An)SiO₄ solid solutions that formed under accidental conditions and could allow the evaluation of the physical properties of these solid solutions. In this perspective, a complementary calorimetry study is underway in order to learn more about the thermodynamic properties of (Zr,Ce)SiO₄ solid solutions and would be a first step in the perspective of studying later (Zr,U)SiO₄ and (Zr,Pu)SiO₄ solid solutions.

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Notes

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How hydrothermal synthesis improves the synthesis of (Zr,Ce)SiO₄ solid solutions

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SUPPORTING INFORMATION

Table SI 1. Synthesis parameters for the presented (Zr,Ce)SiO₄ hydrothermal syntheses.

Label	Zr:Ce mol ratio	C _{Zr} (mol·L ⁻¹)	C _{Ce} (mol·L ⁻¹)	pH ^{initial}	T (°C)	Δt (days)	Final phase
(1)				[H ₃ O ⁺] = 1 mol·L ⁻¹			ZrSiO ₄
(2)				1.0			ZrSiO ₄
(3)				2.0			ZrSiO ₄ + ε ZrO ₂
(4)				3.1			ZrSiO ₄ + ε ZrO ₂
(5)				4.1			ZrSiO ₄
(6)	100:0	0.20	0	4.9	250	20	ZrSiO ₄
(7)				6.0			ZrSiO ₄
(8)				7.0			ZrSiO ₄
(9)				8.0			ZrSiO ₄
(10)				9.0			ZrSiO ₄
(11)				10.0			ZrSiO ₄ + another phase
(12)				11.7			another phase
(13)	100:0	0.20	0	1.0	250	1	ZrSiO ₄
(14)						7	ZrSiO ₄
(15)	100:0	1.0	0	1.0	250	7	ZrSiO ₄
(16)		0.05					ZrSiO ₄ + ZrO ₂
(17)	95:5	0.19	0.01				(Zr,Ce)SiO ₄
(18)	90:10	0.18	0.02				(Zr,Ce)SiO ₄
(19)	80:20	0.16	0.04				(Zr,Ce)SiO ₄
(20)	60:40	0.12	0.08				(Zr,Ce)SiO ₄ + ZrO ₂
(21)	50:50	0.10	0.10	1.0	250	7	(Zr,Ce)SiO ₄ + ZrO ₂ + CeO ₂
(22)	40:60	0.08	0.12				(Zr,Ce)SiO ₄ + ZrO ₂ + CeO ₂
(23)	30:70	0.06	0.14				ZrO ₂ + CeO ₂
(24)	0:100	0	0.20	6.5	150	7	CeSiO ₄

Table SI 2. Unit cell parameters and volume obtained by Rietveld refinements made from the PXRD patterns of pristine ZrSiO₄ samples prepared under hydrothermal conditions ($\Delta t = 20$ days, $T = 250^\circ\text{C}$) with starting zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ and with various initial pH values.

Synthesis	pH _{ini}	<i>a</i> (Å)	<i>c</i> (Å)	<i>V</i> (Å ³)
(1)	[HNO ₃] = $1.0 \text{ mol}\cdot\text{L}^{-1}$	6.612(5)	5.973 (6)	261.1(7)
(2)	1.0	6.627(4)	5.979(4)	262.6(5)
(3)	2.0	6.627(3)	5.982(4)	262.7(4)
(4)	3.1	6.628(5)	5.980(6)	262.7(7)
(5)	4.1	6.623(4)	5.972(5)	262.0(5)
(6)	4.9	6.625(7)	5.972(8)	262.1(9)
(7)	6.0	6.625(5)	5.975(6)	262.2(7)
(8)	7.0	6.621(7)	5.968(8)	261.6(9)
(9)	8.0	6.618(7)	5.970(8)	261.5(9)
(10)	9.0	6.614(6)	5.974(7)	261.3(8)

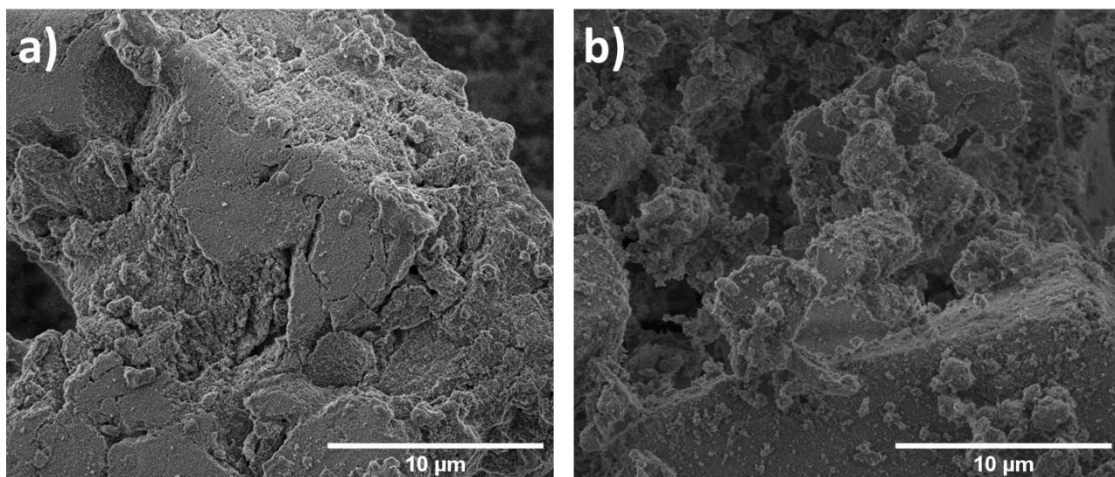


Figure SI 1. SEM micrographs recorded for **a)** pristine (14) and **b)** annealed (1000°C) ZrSiO_4 sample prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1.0$) starting with zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$.

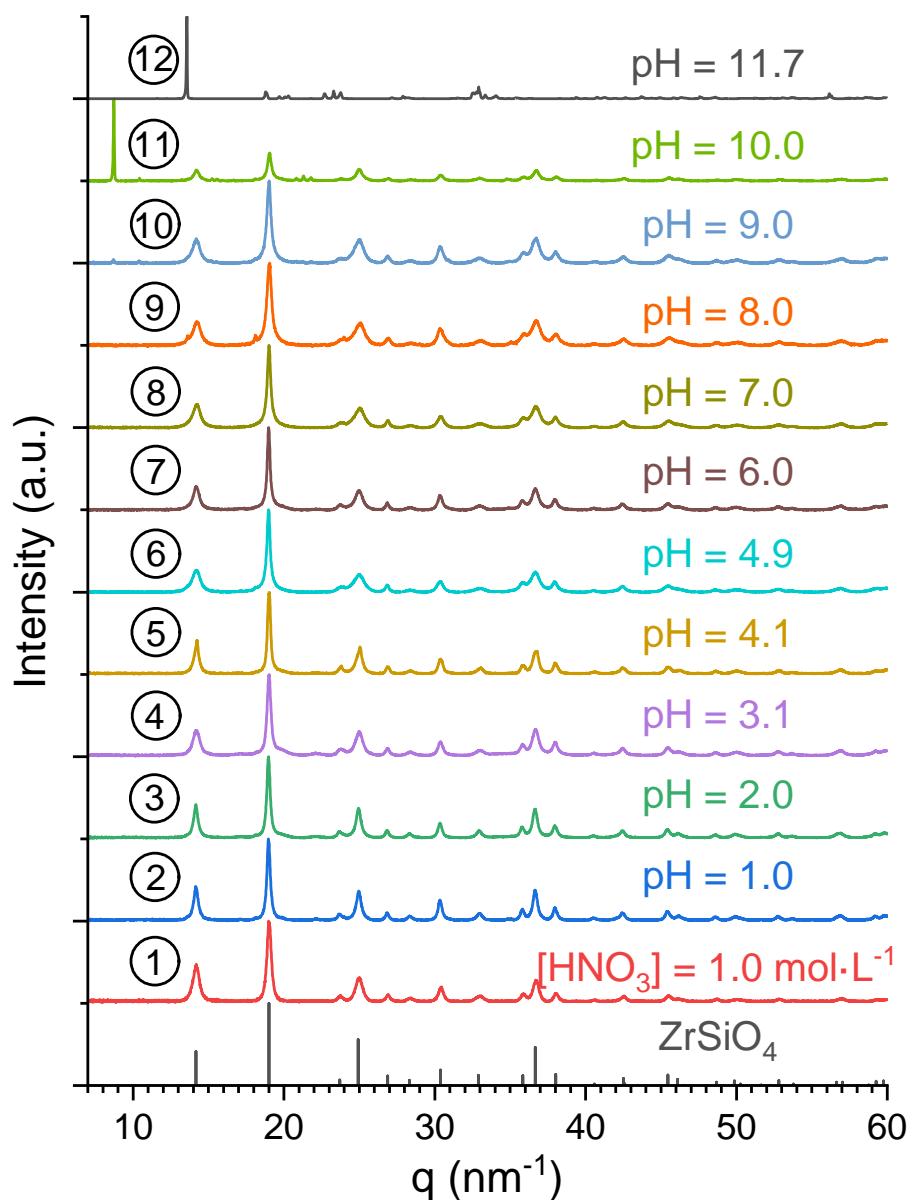


Figure SI 2. PXRD patterns obtained for ZrSiO₄ samples prepared under hydrothermal conditions (20 days, $T = 250^\circ\text{C}$) starting with zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ and for various initial pH values: $1.0 \text{ mol}\cdot\text{L}^{-1}$ HNO₃ (1), pH = 1.0 (2), pH = 2.0 (3), pH = 3.1 (4), pH = 4.1 (5), pH = 4.9 (6), pH = 6.0 (7), pH = 7.0 (8), pH = 8.0 (9), pH = 9.0 (10), pH = 10.0 (11) and pH = 11.7 (12), plotted in scattering vector.

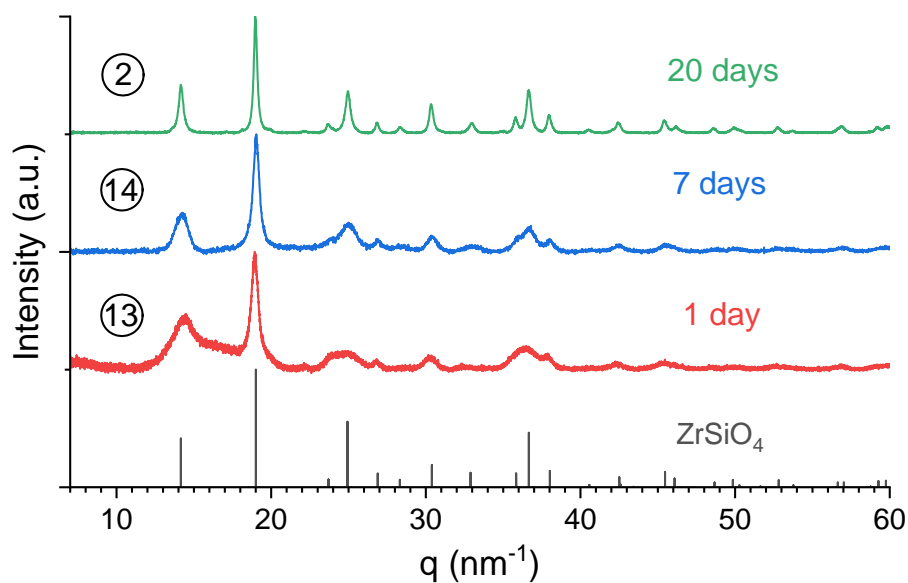


Figure SI 3. PXRD patterns obtained for ZrSiO_4 samples prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, $\text{pH} = 1$) starting with zirconium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ and for holding time of 1 day (13), 7 days (14) and 20 days (2), plotted in scattering vector.

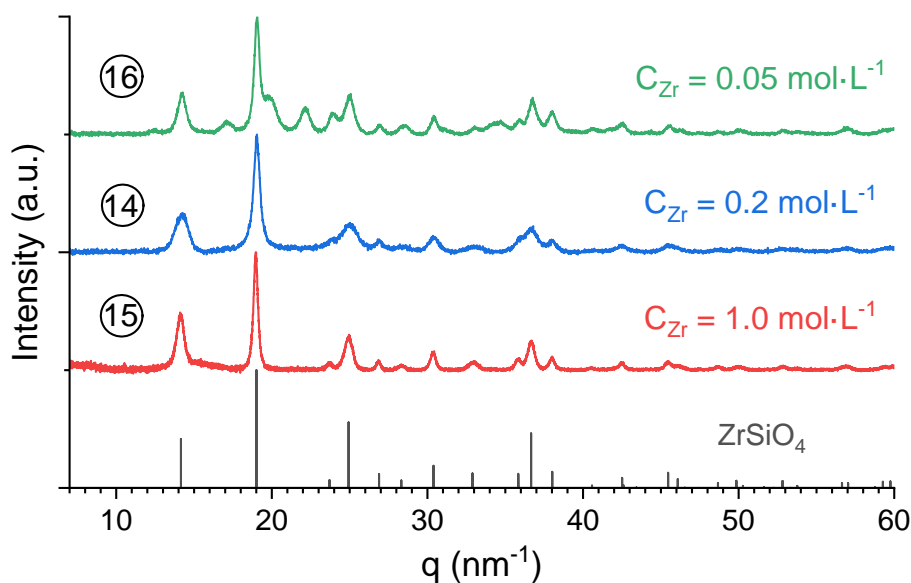


Figure SI 4. PXRD patterns obtained for ZrSiO_4 samples prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days) at $\text{pH} = 1.0$ starting with zirconium and silicate concentrations of $1.0 \text{ mol}\cdot\text{L}^{-1}$ (15), $0.2 \text{ mol}\cdot\text{L}^{-1}$ (14) and $0.05 \text{ mol}\cdot\text{L}^{-1}$ (16), plotted in scattering vector.

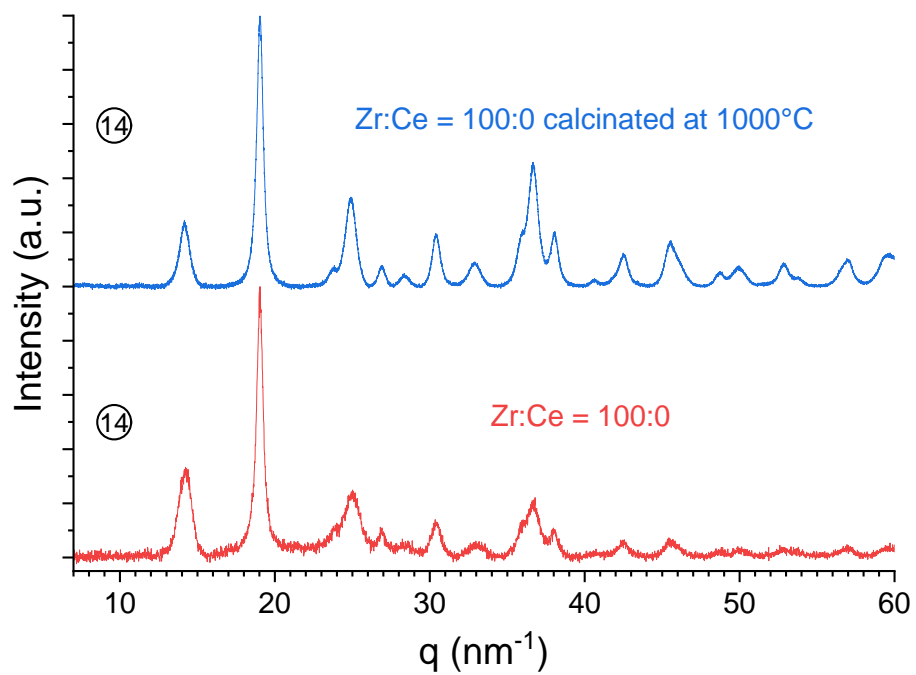


Figure SI 5. PXRD patterns recorded for a pristine and annealed (after TGA, 1000°C) ZrSiO₄ sample (14) prepared under hydrothermal conditions (T = 250°C, 7 days) at pH = 1.0 starting with zirconium and silicate concentrations of 0.2 mol·L⁻¹, plotted in scattering vector.

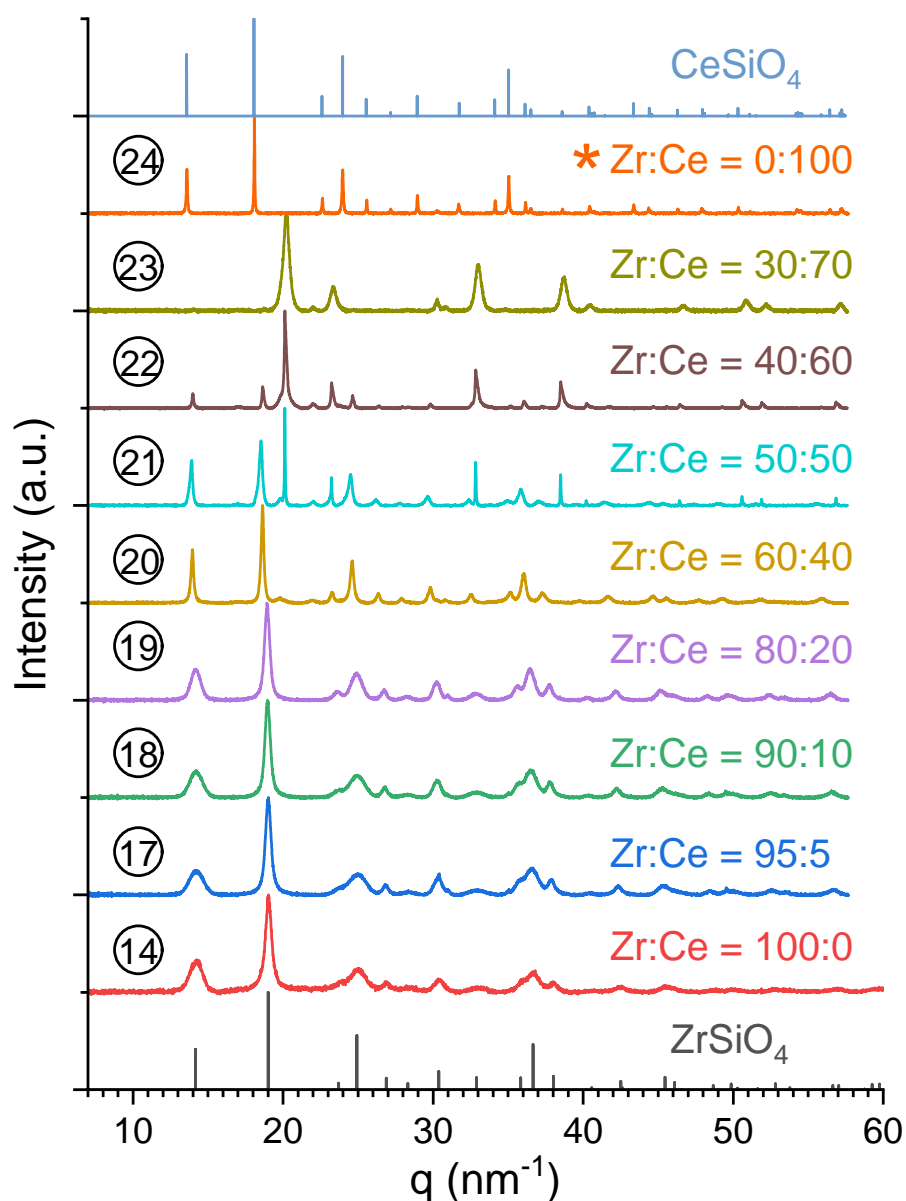


Figure SI 6. PXRD patterns recorded for pristine (Zr,Ce)SiO₄ solid solutions with various chemical compositions prepared under hydrothermal conditions ($T = 250^{\circ}\text{C}$, 7 days, $\text{pH} = 1.0$) starting with Zr + Ce and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ for Zr:Ce = 100:0 (14), 95:5 (17), 90:10 (18), 80:20 (19), 60:40 (20), 50:50 (21), 40:60 (22), 30:70 (23). * Reference CeSiO₄ sample ($T = 150^{\circ}\text{C}$, 7 days, $\text{pH} = 6.5$) starting with cerium and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$ (24), plotted in scattering vector.

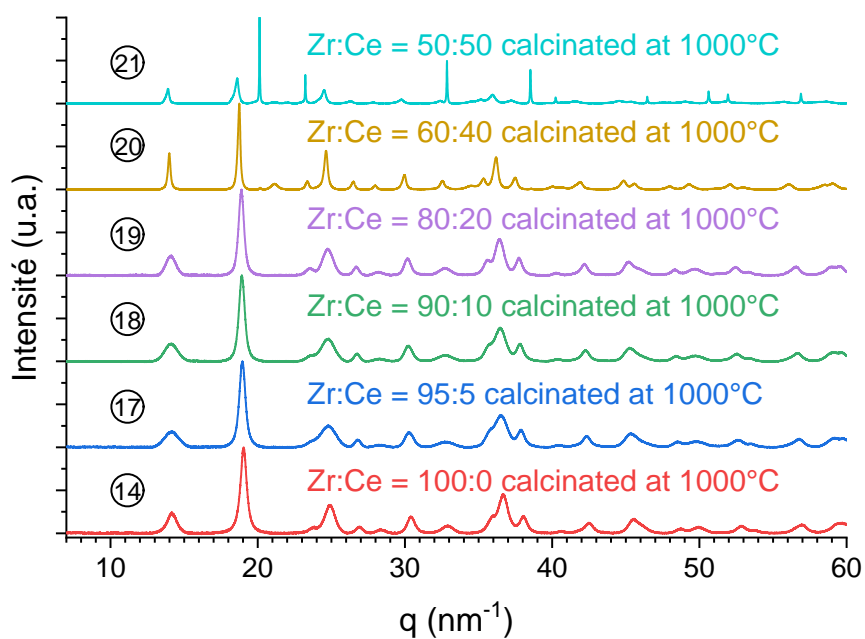


Figure SI 7. Synchrotron PXRD patterns recorded for $(\text{Zr,Ce})\text{SiO}_4$ solid solutions annealed at 1000°C with Zr:Ce = 100:0 (14), Zr:Ce = 95:5 (17), Zr:Ce = 90:10 (18), Zr:Ce = 80:20 (19), Zr:Ce = 60:40 (20) and Zr:Ce = 50:50 (21) prepared under hydrothermal conditions ($T = 250^\circ\text{C}$, 7 days, $\text{pH} = 1.0$) starting with Zr + Ce and silicate concentrations of $0.2 \text{ mol}\cdot\text{L}^{-1}$, plotted in scattering vector.

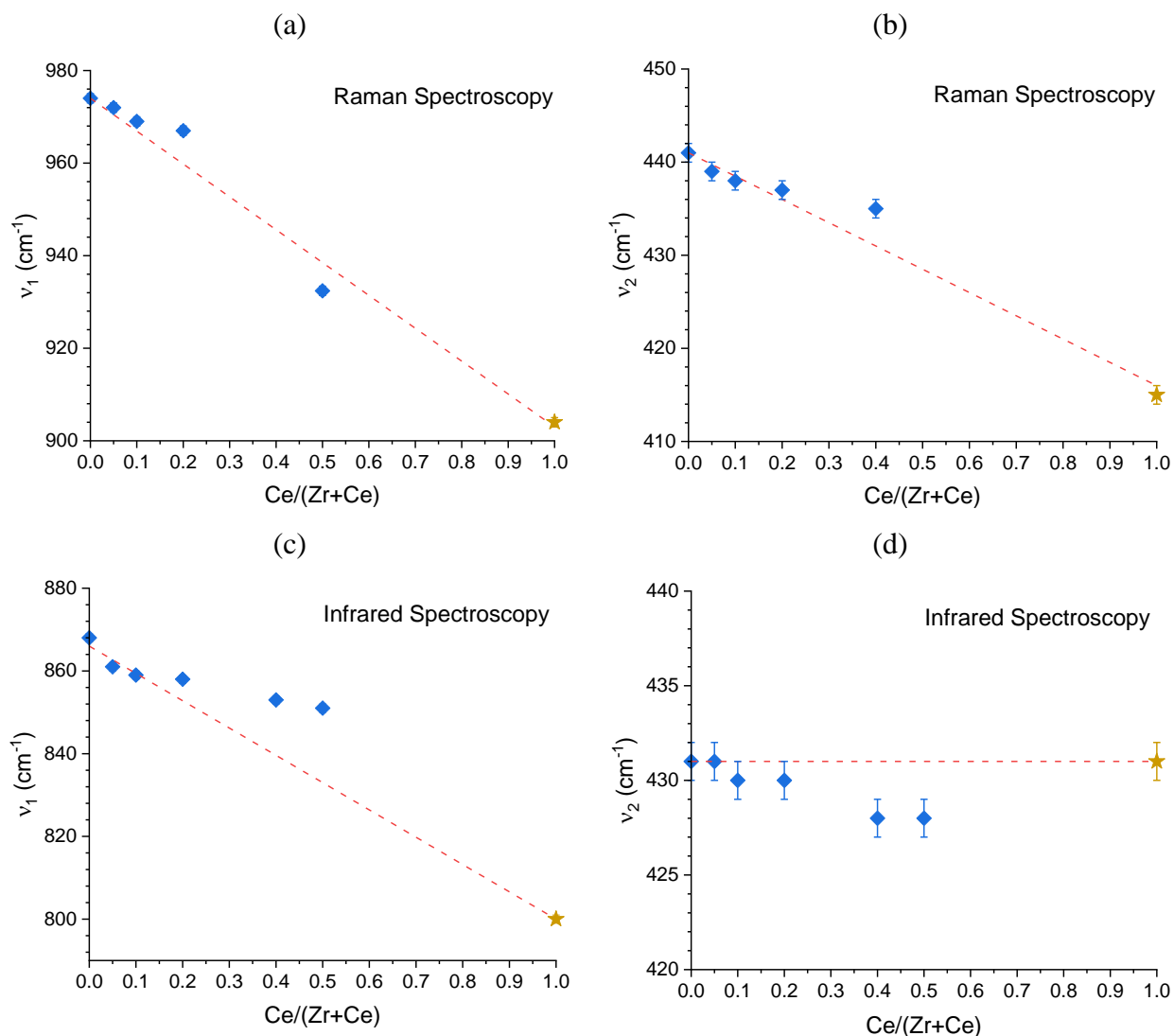
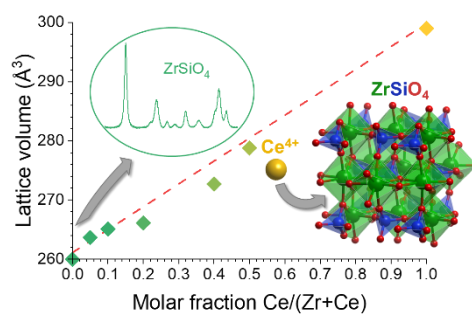


Figure SI 8. Variation of the ν_1 (a) and ν_2 (b) observed by Raman and ν_1 (c) and ν_2 (d) by IR spectroscopy for $(\text{Zr,Ce})\text{SiO}_4$ (annealed samples at 1000°C) versus the expected chemical composition. * Reference CeSiO_4 sample wasn't obtained in the same chemical conditions as the $(\text{Zr,Ce})\text{SiO}_4$ samples.

ν_4 is not observable for Raman spectroscopy, due to its low intensity. ν_3 band position was difficult to determine by IR spectroscopy was difficult to determine with a good accuracy due to its low intensity and nearness with SiO_2 characteristic band. Moreover, ν_1 band position for $\text{Zr}:\text{Ce} = 60:40$ by Raman spectroscopy was masked by ν_3 band and ν_2 band position for $\text{Zr}:\text{Ce} = 50:50$ by Raman spectroscopy masked by CeO_2 T_{2g} band.

FOR TABLE OF CONTENTS ONLY



Through the determination of the conditions adapted to the hydrothermal synthesis of ZrSiO₄, conditions allowing the formation of (Zr,Ce)SiO₄ solid solution with Ce content up to 40% have been identified.