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## Fabrication of (U,Am)O<sub>2</sub> pellet with controlled porosity from oxide microspheres

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### **Abstract**

U<sub>1-x</sub>Am<sub>x</sub>O<sub>2±δ</sub> mixed-oxides are considered as promising compounds for americium heterogeneous transmutation in Sodium Fast Neutron Reactor [1–3]. Porous microstructure is envisaged in order to facilitate helium and fission gas release and to reduce pellet swelling during irradiation and under self-irradiation. In this study, the porosity is created by reducing (U,Am)<sub>3</sub>O<sub>8</sub> microspheres into (U,Am)O<sub>2</sub> during the sintering. This reduction is accompanied by a decrease of the lattice volume that leads to the creation of open porosity. Finally, an (U<sub>0.90</sub>Am<sub>0.10</sub>)O<sub>2</sub> porous ceramic pellet (D~89% of the theoretical density TD) with controlled porosity (≥8% open porosity) was obtained from mixed-oxide microspheres obtained by the Weak Acid Resin (WAR) process [4].

## 1. Introduction

Even though neptunium, americium and curium, known as the minor actinides (MA) represent less than 0.1 wt.% of the spent nuclear fuel, they are the main responsible for long-term radioactivity and heat load of the ultimate waste after plutonium recycling [5,6]. A possible option to reduce the long-term radiotoxicity of spent fuel nuclear wastes is to deploy MA transmutation into lighter short-lived elements, for instance in Fast Neutron Reactors (FNRs) developed in the framework of GEN-IV International Forum [1,7]. The program specially focuses on the transmutation of americium since this element is the most abundant and the most active among MA [2,8]. In France, among the different MA transmutation modes [2,9–11], heterogeneous transmutation is one of the most considered options for the future nuclear fuel cycle [1]. It consists in irradiating uranium–americium mixed-oxide compounds ( $U_{1-x}Am_xO_{2\pm\delta}$ ) located at the core periphery. Such irradiation targets are called AmBB (Americium Bearing Blankets) [3]. The content of americium [ $Am/(Am+U)$ ] in such materials can reach 15 at.%. In this context, several processes are being developed to fabricate such mixed-oxide pellets which are to meet the specifications required to perform analytical irradiations such as Am content, O/M ratio (oxygen to metal with  $M = U + Am$ ), chemical and microstructural homogeneity, pellet dimensions and density, open porosity ratio, etc [2,8,12,13].

The development of a dustless innovative route, adapted from the Weak Acid Resin process (WAR) is currently studied. It consists in impregnating resin microspheres with the desired element(s) in their cationic form and then in submitting the metal loaded resin to a calcination step to obtain the single or mixed-oxide microsphere precursors. Those microspheres are then shaped into pellets which are sintered to produce single or mixed-oxide ceramic. The over-all process is the so called Calcined Resin Microsphere Pelletizing (CRMP) process [12].

The fabrication of dense pellets by CRMP process was demonstrated for compounds such as  $UO_2$  [14] and  $(U,Am)O_2$  [15]. Tailored-porous samples from oxide spheres were fabricated only for  $UO_2$  [16]. As the main objective of MARIOS and DIAMINO irradiations was to assess the influence of the fuel microstructure (dense and porous), it could be important to also demonstrate the feasibility of  $(U,Am)O_2$  pellets fabrication with controlled density from CRMP process for an eventual future irradiation. Indeed, the porous microstructure is considered to facilitate the release of helium produced during irradiation but also to limit pellet swelling under self-irradiation.

In this study, we focus on demonstrating the feasibility of the fabrication of porous ceramic pellets with controlled porosity (8-10% open porosity) from mixed-oxide microspheres obtained by the Weak Acid Resin (WAR) process [4]. An  $(U_{0.90}Am_{0.10})O_2$  pellet with controlled porosity (89%TD) and with an open porosity of 9% was fabricated.

## 2. Experimental

### 2.1. Oxide microsphere synthesis

### 2.1.1. Resin preparation

The ion exchange resin used for the fixation was an IMAC HP333 carboxylic resin supplied by Dow Chemicals (Chauny, France). The resin was manually sieved under deionised water and the 630-800  $\mu\text{m}$  diameter size range was collected for the fixation. The resin was introduced in a column to be washed by successive percolations of 1 M nitric acid ( $\text{HNO}_3$  - Fisher Chemical, Certified ACS Plus), demineralized water, 1 M ammonia solution (Merck, Pro Analysis) and demineralized water again. Eventually, the washed resin was prepared in its protonated form by a final percolation of 1 M  $\text{HNO}_3$ , followed by a demineralized water rinse cycle.

### 2.1.2. Preparation of the loading solution (U/Am)

The mixed loading solution was prepared in several steps. An americium nitrate solution ( $\text{Am}(\text{NO}_3)_3$ ) was first obtained by the dissolution of 3 g of  $\text{AmO}_2$  powder in a  $0.85 \text{ mol}\cdot\text{L}^{-1}$  nitric acid solution. The solution corresponds to a volume of 125 mL and to a  $0.84 \text{ mol}\cdot\text{L}^{-1}$  Am nitrate solution. The second step is the adding of 17 g of trioxide solid ( $\text{UO}_3$ ) in the previous solution to increase both pH and uranium concentration by reacting with acid excess and by forming hydrolysed uranyl species. The final solution corresponded to an Acid Deficient Nitrate Uranyl solution (ADUN) [17] with the following uranium formula  $\text{UO}_2(\text{NO}_3)_{1.25}(\text{OH})_{0.75}$  and with uranium and americium concentrations of  $0.111$  and  $0.012 \text{ mol}\cdot\text{L}^{-1}$ , respectively, an Am over metal  $[\text{Am}/(\text{Am}+\text{U})]= 10.1 \pm 0.2 \text{ at.}\%$  and a volume of 860 mL.

### 2.1.3. Impregnation

The resin microspheres (about 7 g) were first introduced and rehydrated into a plexiglass column column (1.5 cm internal diameter and 16 cm height). Fixation of both  $\text{UO}_2^{2+}$  and  $\text{Am}^{3+}$  cations was performed by percolating the ADUN solution by recirculation through the column during 5 h. pH of the loaded solution was recorded during impregnation. The cations exchange reaction was considered as achieved when the pH did not change significantly. The loaded resin was rinsed with deionised water before being drained under low vacuum. The microspheres were then transferred in quartz crucible, shown in Figure 1, to be calcined in air at  $700^\circ\text{C}$  for 4 h. A part of these oxide microspheres (65 wt.%) was secondly reduced in  $\text{Ar}/4\text{vol.}\%\text{H}_2$  at  $700^\circ\text{C}$  for 5 h.

A content of  $11.1 \pm 0.2 \text{ at.}\%$  Am measured by TIMS analysis by dissolving 20 mg of reduced oxide microspheres  $(\text{U,Am})\text{O}_2$  in nitric acid (5 M), which can demonstrate the quasi-congruence of the Am fixation ( $\sim 10 \text{ at.}\%$  in the initial solution).



Figure 1 : Pictures of U-Am oxide microspheres synthesis in hot cell in Atalante facility. From left to right: loaded resin in plexiglass column, loaded resin in the quartz crucible before calcination in airflow and oxidized  $(U,Am)_3O_8$  microspheres.

#### 2.1.4. Calcination - Thermal treatment

Two thermal treatments were then applied to the loaded resin in an instrumented tubular furnace.

The resin was first calcined at  $700^\circ\text{C}$  for 4 h under air flow (15 NL/h) with a low heating rate of  $2\text{K}\cdot\text{min}^{-1}$  preventing any microsphere degradation and guarantying a proper homogeneity of the calcination temperature (mass loss of 53.5%). A part of the latter was then submitted to a second heat treatment under reductive atmosphere  $\text{Ar}/4\text{vol.}\%\text{H}_2$  (12 NL/h) at  $700^\circ\text{C}$  for 5 h with a heating rate of  $10\text{K}\cdot\text{min}^{-1}$ .

Crystalline structure analyses on the resulting oxide were performed by X-ray diffraction (XRD) using a Bruker D8 Advance diffractometer (Madison, WI) with a Cu X-ray source ( $K\alpha_1/\alpha_2$ ,  $\lambda = 1.54059/1.54439\text{ \AA}$ ) and equipped with a special sample holder for radioactive material measurements. The oxide microspheres were first milled manually and then mixed in grease with Au powder (Sigma Aldrich, >99.9%). Gold is used as standard to control any potential instrumental deviation.

A nuclearized ZEISS Supra 55 VP Field Emission Gun (FEG) scanning electron microscopy (SEM) (Oberkochen, Germany) was used for the morphology and the microstructure observations of the oxide microspheres.

#### 2.2. Fabrication of sintered $(U,Am)O_2$ pellet with controlled porosity

In this study the porosity was not created by under densification but by the introduction of  $(U,Am)_3O_8$  as a inorganic pore-former. A reductive sintering of a mix of  $(U,Am)_3O_8$  and  $(U,Am)O_2$  microspheres led to the creation of open porosity thanks to the transformation of  $(U,Am)_3O_8$  into  $(U,Am)O_2$ .

In order to obtain porous samples with a density in the 86-90% TD range, the specific ratio of 65 wt.% of reduced microspheres  $(U,Am)O_2$  were mixed with 35 wt.% oxidized microspheres  $(U,Am)_3O_8$ . After the pelletization step at 700 MPa in a three-part matrix (5.0 mm diameter), the green pellet (about 600 mg) was sintered under reductive atmosphere at 1700°C-4h in a high temperature furnace. The flowchart of this porous pellet fabrication process is shown in Figure 2.

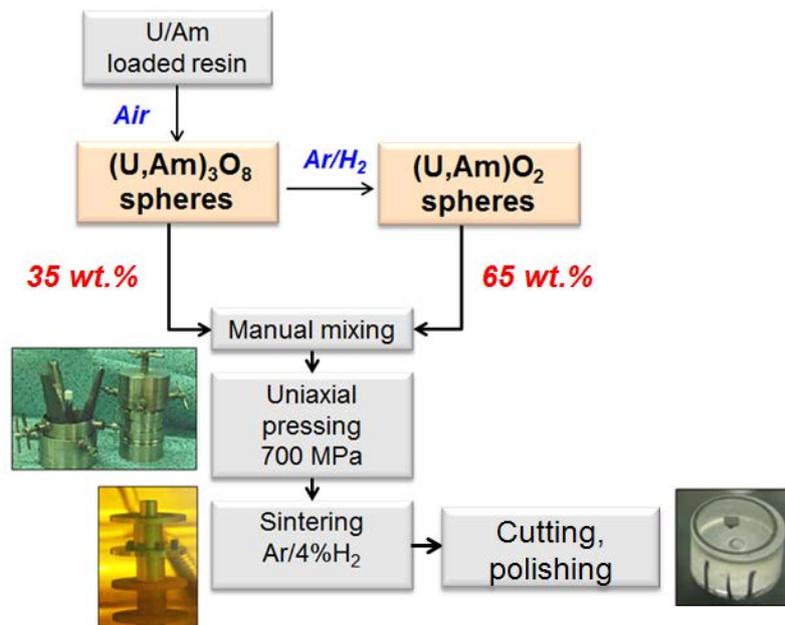


Figure 2 : Porous pellets fabrication flowchart using oxide microsphere precursors.

Accurate diameter measurement was performed for sintered pellet using a laser micrometer coupled to a step motor to allow measurements in the whole height every 20  $\mu\text{m}$ .

Optical microscopy was performed on polished sample using an Olympus BX30M.

XRD analyses were performed on powdered sintered pellet using the same method as described in paragraph 2.1.4.

### **3. Results and discussion**

#### **3.1. Characterization of U-Am mixed-oxide microspheres**

The XRD patterns obtained from the powdered oxide microsphere after calcination under air and followed by a reduction step under Ar/4vol.%H<sub>2</sub> are presented in Figure 3 in red (bottom) and blue (top), respectively.

After the calcination under air, the XRD pattern is characteristic of a single hexagonal phase (P-62m) with the following refined parameters  $a=6.838$  (1) Å and  $c=4.165$  (1) Å.

After the reduction under Ar/4vol.%H<sub>2</sub>, only peaks derived of a single fluorite-type phase (Fm-3m) are visible which proves the quantitative reduction of (U,Am)<sub>3</sub>O<sub>8</sub> to (U,Am)O<sub>2</sub> during the second thermal treatment. The refined lattice parameter was found equal to 5.467 (1) Å. This parameter is higher than the one previously obtained by E. Remy et al. [15] (5.4511 Å) with the same CRMP process, which indicates a higher reduction for our sample.

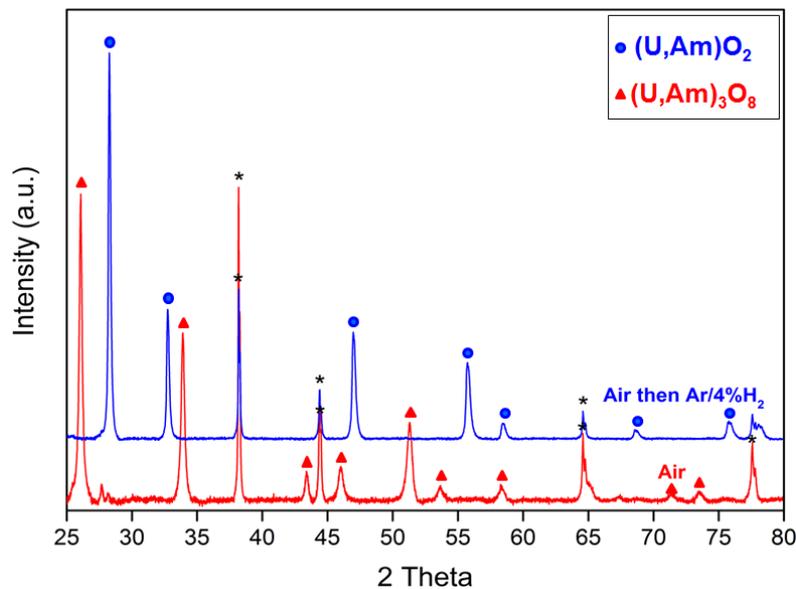
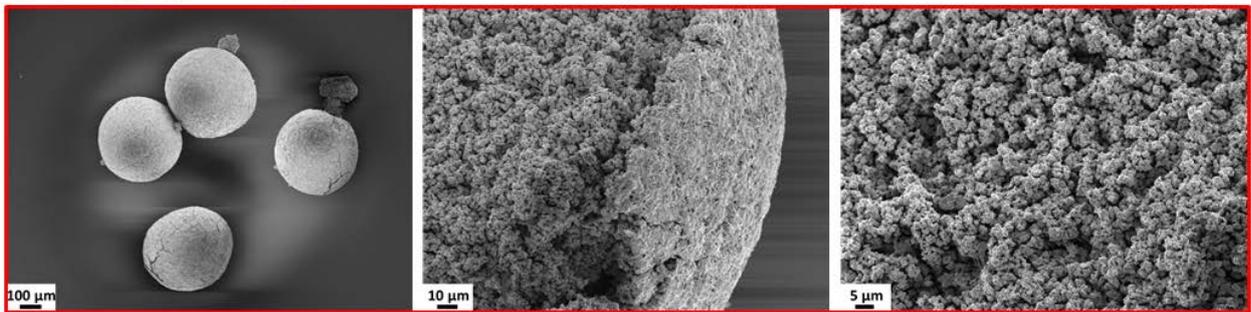


Figure 3 : XRD patterns for U-Am oxide microspheres after the first calcination (air, bottom), and after second heat treatment (Ar/4vol.%H<sub>2</sub>, top) (\*gold standard).

SEM observation (Figure 4) revealed that the spheres have a porous microstructure after calcination under air as well as after the second heat treatment under Ar/4vol.%H<sub>2</sub>.

**(U,Am)<sub>3</sub>O<sub>8</sub> - hexagonal (calcination under air)**



**(U,Am)O<sub>2</sub> - cubic (heat treatment under reductive atmosphere)**

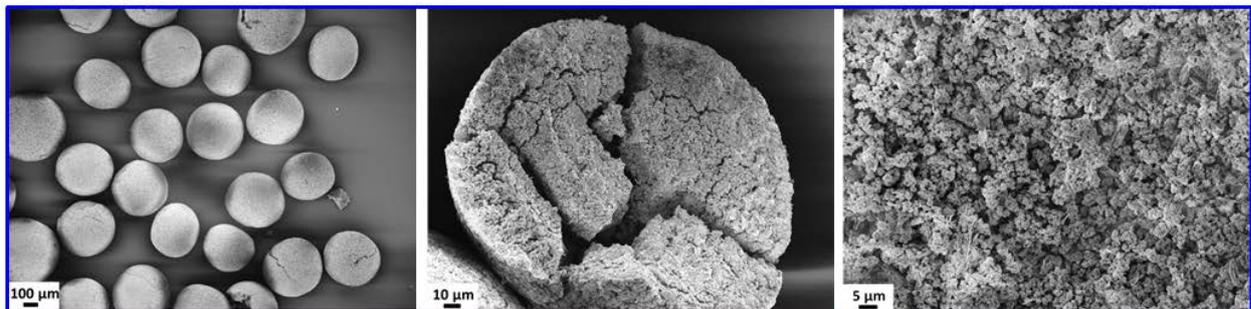


Figure 4 : FEG-SEM micrographs in secondary electron mode showing oxidized and reduced microspheres.

*3.2. Characterization of sintered (U,Am)O<sub>2</sub> pellet*

*After sintering, the final pellet ( $m = 560$  mg,  $h = 4.4$  mm, diameter = 4.1 mm) is rectilinear and neither deformations nor particular defects were observed.*

A percolating porosity network was created during the reductive sintering by volume shrinkage of (U,Am)<sub>3</sub>O<sub>8</sub> phase in (U,Am)O<sub>2</sub>. This network is homogeneously spread into the whole sample.

*3.2.1. Density and microstructure*

The final density was obtained via the Archimedes method in bromobenzene. It was found a density of 89% TD with 9% of open porosity and 2% of closed porosity. These characteristics strongly fit with the required porous pellet specifications for previous irradiation programs (86-89%TD, with open porosity  $\geq 8\%$ ) such as MARIOS [18,19].

Figure 5 presents a longitudinal observation of the sintered (U,Am)O<sub>2</sub> pellet by optical microscopy. The porosity is homogeneously distributed through the pellet and the open porosity is interconnected on a long path. Memory form of ex-(U,Am)<sub>3</sub>O<sub>8</sub> microspheres is rather hemispherical

and could be compared with the memory form of ex-U<sub>3</sub>O<sub>8</sub>-AmO<sub>2</sub> agglomerates observed in MARIOS samples which had no specific form.



Figure 5 : Optical microscopic observation of the internal microstructure of sintered (U,Am)O<sub>2</sub> porous pellet.

### 3.2.2. XRD analysis

XRD pattern obtained on the sintered pellet is characteristic of the fluorite-type phase, with a refined parameter of 5.470(1) Å. The latter is very close to usual lattice parameter found for U<sub>0.90</sub>Am<sub>0.10</sub>O<sub>2±δ</sub> sintered pellets fabricated by metallurgical process [18,20]. The sintering step provided the total reduction of (U,Am)<sub>3</sub>O<sub>8</sub> to (U,Am)O<sub>2</sub>, with a volume lattice decrease for an hexagonal to a fluorite phase creating the visible open porosity network (Figure 5).

## 4. Conclusion and perspectives

This work focused on the fabrication of mixed-oxide pellets with controlled porosity by implementing the CRMP process. The synthesis of mixed-oxide microspheres by impregnation and calcination of loaded resin beads really limits the production of fine particles. Moreover, the CRMP process is a simplified one, with only two steps after the oxide microsphere synthesis: pelletizing and sintering.

The feasibility of porous AmBB fabrication was demonstrated for a potential heterogeneous transmutation of americium in FNRs by the fabrication of an 89% DT porous pellet with 9% of open porosity. The developed process allows a homogeneous distribution of the open porosity through the entire pellet created by the complete reduction of hexagonal  $(U,Am)_3O_8$  into fluorite-type phase  $(U,Am)O_2$ .

This porosity network could be tuned by modifying the initial resin bead diameter or by doing a specific mix between several microsphere diameters.

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