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1 **A robust and parsimonious model for caesium sorption on clay minerals and natural clay**
2 **materials**

3 Mohamed A. Cherif ¹, Arnaud Martin-Garin ¹, Frédéric Gérard ² and Olivier Bildstein ^{3,*}

4 ¹IRSN, PRP-ENV/SERIS, Laboratory of Biogeochemistry, Bioavailability and Transfers of
5 radionuclides, Bldg. 183 - BP3, 13115 Saint-Paul-Lez-Durance Cedex, France

6 ²INRA, UMR Eco&Sols - Functional Ecology and Biogeochemistry of Soils & Agroecosystems,
7 Place Pierre Viala, 34060 Montpellier, France

8 ³CEA, DEN, Laboratory for Transfer Modelling in the Environment, Bldg. 727, 13108 Saint-Paul-
9 lez-Durance cedex, France

10 * author to whom correspondence should be addressed

11

12 **ABSTRACT**

13 Caesium (Cs) is one of the most studied radionuclides in the fields of nuclear waste disposal
14 and environmental sciences. The overall objective of this work is to improve the tools designed to
15 describe and predict migration, retention, and bioaccumulation processes in the geosphere and the
16 biosphere, particularly in the soil / solution soil / plant roots systems. Cs sorption on clay minerals
17 has been extensively measured and modelled because these minerals control Cs mobility and
18 (bio)availability in the environment.

19 A critical analysis of published experimental data on Cs sorption by clay minerals and
20 natural clay materials along with the different models was performed in an attempt to elaborate and
21 evaluate a generic model for Cs sorption. This work enabled us to propose a robust and
22 parsimonious model for Cs sorption, which combines the surface complexation and cation
23 exchange approaches invoking only two types of surface sites: frayed edge and exchange sites. Our
24 model, referred to as the “1-pK DL/IE model”, takes into account the competition between Cs and

25 other cations as well as the influence of the ionic strength and pH of the solution.

26 This model was successfully calibrated for Cs sorption on three reference clay minerals
27 (illite, montmorillonite and kaolinite), in a wide range of Cs concentrations and physicochemical
28 conditions. Using the same parameters, we tested our model on several natural clayey materials
29 containing a single to several clay minerals. The goodness-of-fit obtained with natural materials
30 containing a single clay mineral demonstrates the robustness of the model. The results obtained
31 with natural mixed clay materials confirm the predictive capability of the model and also allowed
32 us to test the sensitivity to the mineral composition of these materials (uncertainties). We found
33 that illite is usually the most reactive clay mineral with respect to Cs sorption and that component
34 additivity is applicable when the contribution of other clay minerals becomes non negligible. The
35 whole set of model-measurement comparisons performed in this study provides a high level of
36 confidence in the capabilities of the 1-pK DL/IE model as an interesting predictive tool.

37

38 **1 Introduction**

39

40 Since decades the pollution of natural ecosystems by radionuclides has become a major
41 concern for society. In particular, caesium (Cs) is widely studied by environmental scientists using
42 different analytic and experimental techniques (Benedicto et al., 2014; Bostick et al., 2002;
43 Chorover et al., 2003; Missana et al., 2014a; Savoye et al., 2012; Wendling et al., 2005). One of
44 the Cs radio-isotopes, ^{137}Cs , is an important fission product from the irradiation of uranium-based
45 fuels, has a relatively long life ($t_{1/2} = 30$ years), and constitutes a significant radioecological hazard
46 due to its hard gamma emission. After release of radioactivity into the environment, this
47 radioelement is considered as the main source of contamination of soils (Avery, 1996; Strebl et al.,
48 1999) and the principal source of radioactivity of nuclear waste in the timeframe of the first one
49 hundred years. Moreover, radiocaesium always exists as the monovalent cation Cs^+ , with chemical

50 properties similar to potassium (K^+) (Kamei-Ishikawa et al., 2011; Roca and Vallejo, 1995),
51 having very high solubility (Benedicto et al., 2014; Fuller et al., 2014; Missana et al., 2014b; NDA,
52 2010). Cs mobility in soils environment is also known to be significantly influenced by retention
53 capacity for Cs^+ (Avery, 1996; Strebl et al., 1999).

54 Understanding the processes that control the bioavailability and mobility of Cs in the
55 biosphere and geosphere constitutes a major challenge that has to be tackled in order to provide a
56 good estimate of its health and ecological hazard. Geochemical reactions, particularly those
57 occurring at mineral/water interfaces, closely control the fate and behavior of the major elements
58 or trace in these systems (Koretsky, 2000). The adsorption/desorption process (i.e. sorption)
59 generally dominates the interactions between Cs and soil as this element hardly forms solid species
60 in natural environments (Dzene et al., 2015).

61 Clay minerals are often invoked as the most important minerals that control the
62 bioavailability and mobility (migration-retention) of Cs in subsurface (e.g. soils and sediments)
63 and groundwater environments (Bostick et al., 2002; Chorover et al., 2003; Missana et al., 2014a;
64 Savoye et al., 2012; Shenber and Eriksson, 1993; Wendling et al., 2005). In soils, the most
65 efficient carrier phases of Cs^+ are clay minerals, especially 2:1 layer clays, with a permanent
66 charge arising from isomorphic substitution (Bostick et al., 2002; Comans et al., 1991; Comans
67 and Hockley, 1992; Cornell, 1993; Maes and Cremers, 1986; Nakano et al., 2003; Sawhney, 1972;
68 Watanabe et al., 2012); firstly because of their ubiquity and secondly because they have large
69 specific surface areas with a high density of (negatively and positively) charged surface sites
70 (Kraepiel et al., 1999; Langmuir, 1997).

71 Among the different properties of Cs/clay mineral interactions, the non-linear Cs sorption
72 isotherms on phyllosilicate clays, such as illite (Bradbury and Baeyens, 2000; Cornell, 1993;
73 Eliason, 1966; Missana et al., 2004; Poinssot et al., 1999; Staunton and Roubaud, 1997; Wahlberg
74 and Fishman, 1962), is considered to be the result of the heterogeneity of the surface adsorption
75 sites. (i) High affinity sites, which adsorb strongly and specifically Cs, are located on the edges of

76 clay minerals and are therefore usually termed as « Frayed Edge Sites » (FES) (Brouwer et al.,
77 1983; Eberl, 1980; Francis and Brinkley, 1976; Jackson, 1963; Maes and Cremers, 1986; Poinssot
78 et al., 1999; Rich and Black, 1964; Sawhney, 1972; Zachara et al., 2002). (ii.) The remaining sites,
79 which have a lower, non-specific, affinity for Cs, are present in large quantities and located on
80 planar surface and mainly constitute the cation exchange capacity (CEC). On these sites, the
81 adsorption of Cs strongly depends on the composition of the cationic exchangeable population,
82 since the adsorption reaction partly proceeds as cation exchange (Cornell, 1993; Rigol et al., 2002;
83 Staunton and Roubaud, 1997; Zachara et al., 2002). Cs can also be selectively fixed on interlayer
84 sites of phyllosilicate clays (Rigol et al., 2002; Salles et al., 2013). Adsorption therefore depends
85 on the relative affinities of Cs, the nature of the competitive exchanging cations, pH and the ionic
86 strength.

87 Caesium adsorption mechanisms on illite has received a lot of attention for several decades
88 and different studies were produced that inferred the prevailing mechanism (Sawhney, 1972) and,
89 more recently, gave spectroscopic evidence that Cs initially sorbs on illite FES and can eventually
90 enter into the interlayer and be incorporated into the mineral structure (Fuller et al., 2015; Lee et
91 al., 2017). Cesium adsorption on vermiculite and montmorillonite was described by Bostick et al.
92 (2002) as a function of surface coverage using EXAFS. These authors also found that inner-sphere
93 surface complexes may form within the interlayer or at frayed edge sites and were less accessible
94 than outer-sphere complexes. None of these studies exclude the possibility of having Cs adsorbed
95 onto amphoteric sites, at least in the early stage of sorption which would account for the pH-
96 dependency of Cs sorption as opposed to ion exchange in the interlayer. This aspect has received
97 much less attention and experimental data is still lacking, especially for illite. Furthermore, the
98 partial reversibility of Cs is difficult to reconcile with the “collapsed” interlayer resulting from
99 Cs incorporation into the structure of illite (Comans et al., 1991; Comans et al., 1992; De Koning
100 and Comans, 2004; Durant et al., 2018).

101 Several models have been developed in the last two decades in order to interpret and model

102 non-linear Cs sorption onto mineral phases under specific conditions (homoionized mineral phases,
103 many fixed experimental parameters such as pH and ionic strength) (Baeyens and Bradbury, 1997;
104 Benedicto et al., 2014; Bradbury and Baeyens, 2000; Brouwer et al., 1983; Chen et al., 2014;
105 Gutierrez and Fuentes, 1996; Liu et al., 2004; Marques Fernandes et al., 2015; Missana et al.,
106 2014a; Missana et al., 2014b; Montavon et al., 2006; Poinssot et al., 1999; Savoye et al., 2012;
107 Silva, 1979; Zachara et al., 2002). However, no compilation or comparison study of these models
108 has ever been proposed. Moreover, these models, whose parameters are obtained in a semi-
109 empirical manner, do not take into account the effect of pH on the mineral surface charge: this
110 implies that their transferability to heterogeneous environments is limited (Koretsky, 2000;
111 Missana et al., 2008). Another important point is that these models were calibrated for specific
112 physicochemical conditions proper to each study without considering the transferability to other
113 conditions. This constitutes the major limitation of cation exchange models found in literature,
114 which appear to be unsatisfactory to predict Cs sorption a wide range of conditions in natural
115 systems without any modification of the thermodynamic parameters.

116 The objective of this work was (i) to establish a compilation of the current models describing
117 Cs sorption on natural clayey materials and (ii) to perform a cross-comparison of all these models
118 with all the data available in the literature for pure clay minerals and for clay materials. (iii) An
119 alternative model was finally developed in order to improve the performance, the robustness and
120 the predictability of the existing models, based on the combining approach (i.e. cation exchange
121 and surface complexation models) described above, and requiring fewer parameters (i.e.
122 parsimonious) than most of the others modelling approaches. It is designed to be capable of
123 modelling Cs adsorption without any parameter changes over a larger range of chemical conditions
124 (i.e. robustness). We calibrated this model with all the data available in the literature for pure clay
125 minerals (illite, montmorillonite, and kaolinite) and we assessed its validity for clay materials of
126 increasing complexity, ranging for relatively simple materials containing only a single clay mineral
127 to materials such as bentonite and claystones containing several types of clay minerals.

128

129 **2 Material and methods**

130

131 *2.1 Available data on clay minerals and clay materials*

132

133 A compilation of existing experimental and modeled data found in the literature was
134 performed in order to evaluate the existing models for Cs sorption on clay minerals and the one
135 developed here. These data are mostly constituted by Cs sorption isotherms on the most common
136 clay minerals, i.e. illite, montmorillonite, and kaolinite (243 observations), and on a variety of clay
137 materials (378 observations) (**Table 2 and 4**). The adsorption data usually presented in the form of
138 a distribution coefficient (K_d) for the adsorbed Cs as a function of Cs concentration in solution at
139 equilibrium or as a function of pH. The distribution coefficient between the solid and the liquid
140 phase, K_d ($L\ kg^{-1}$), is calculated using the following relationship:

$$141 \quad K_d = \frac{[Cs_{tot}] - [Cs_{eq}]}{[Cs_{eq}]} \frac{V}{m} \quad (1)$$

142 where $[Cs]_{tot}$ is the total Cs concentration of the suspension ($mol\ L^{-1}$), $[Cs]_{eq}$ is the Cs
143 concentration in solution at equilibrium ($mol\ L^{-1}$), m/V is the solid: liquid ratio with m the mass of
144 the clay (kg), and V is the volume of the liquid (L).

145 **Table 1** shows the different clay minerals considered in the present investigation, together
146 with their mineralogical properties and the experimental conditions used in the studies. The most
147 abundant set of experimental data concerns illite with a total 191 observations. Fewer experimental
148 data were found for montmorillonite (51 observations) and kaolinite (67 observations). All the clay
149 minerals had been purified and homoionized in Na, K, Ca or NH_4 before performing Cs sorption.
150 The different types of illite considered in this investigation have different origins (Le Puy, Rojo
151 Carbonero and Morris), but exhibit consistent values for the cationic exchange capacities (CEC)

152 and Specific Surface Area (SSA) (**Table 1**). Note that the illite from Le Puy-en-Velay is usually
153 considered as the “reference” illite and its parameters are preferentially used in several studies
154 (Bradbury and Baeyens, 2000; Chen et al., 2014; Maes et al., 2008; Marques Fernandes et al.,
155 2015).

156 The experimental conditions (i.e. ionic strength and major cation concentration) used in Cs
157 sorption experiments performed on simple clay materials (Rojo Carbonero clay, MX-80 bentonite,
158 Boda claystone and Hanford sediments) are listed in **Tables 2** and **3**, along with their mineralogical
159 properties, the type and content of clay mineral. The experimental data obtained with clay
160 materials containing different clay minerals (FEBEX clay, San Juan clay, Callovo–Oxfordian
161 samples, Opalinus clay and Boom clay) and the corresponding experimental conditions are
162 summarized in **Tables 4** and **5**.

163

164 2.2 *Sorption models*

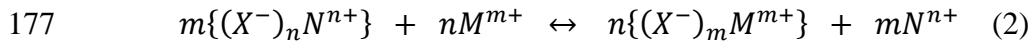
165

166 2.2.1 Ion Exchange

167

168 The theory of ion exchange is a macroscopic approach developed by Bolt (1982). This
169 exchange involves the structural negative charge of mineral clays, which is compensated by
170 cations adsorption within the interlayer space. In general, several exchange sites can be used to
171 calibrate the exchange data. For the model developed here, we chose to start with a model
172 containing only a single site with a capacity equal to the cationic exchange capacity (CEC).

173 The exchange of ions at the solid-water interface is usually described by reactions in which
174 an equivalent amount of counter-ion charge is conserved. For a single negatively charged site $\equiv X^-$
175 (e.g. cationic exchange sites of the clays), the ionic exchange reaction involving ions N^{n+} et M^{m+} is
176 expressed as follows:



178 Being reversible, the thermodynamic constant of cation exchange reactions is expressed by:

179
$$\frac{M}{N}K = \frac{[(X^-)_m M^{m+}]^n \cdot [N^{n+}]^m \frac{(f_{(X^-)_m M^{m+}}^n)(\gamma_{M^{m+}}^n)}{(f_{(X^-)_n N^{n+}}^m)(\gamma_{N^{n+}}^m)}}{[(X^-)_n N^{n+}]^m \cdot [M^{m+}]^n} \quad (3)$$

180 where [] is the concentration of species in solution (mol L⁻¹) or adsorbed to the solid surface (mol
181 kg⁻¹), and f and γ are the activities of the adsorbed and aqueous species, respectively.

182 In solutions of constant total molality, frequently, the ratio of the activity coefficients of the
183 adsorbed species is nearly constant (the ratio is set to 1) (Gaines and Thomas, 1953), and the
184 selectivity coefficient K_c is defined as follows (Jacquier et al., 2004; Savoye et al., 2012) :

185
$$\frac{M}{N}K_c = \frac{[(X^-)_m M^{m+}]^n \cdot [N^{n+}]^m \gamma_{M^{m+}}^n}{[(X^-)_n N^{n+}]^m \cdot [M^{m+}]^n \gamma_{N^{n+}}^m} \quad (4)$$

186

187 2.2.2 Surface Complexation

188

189 Analogous to ion complexation in solution, ‘Surface Complexation Models’ (SCM) (Avena
190 and De Pauli, 1998; Bradbury and Baeyens, 2002; Davis et al., 1978; Davis and Kent, 1990;
191 Ikhsan et al., 2005; Sposito, 1984) provide a molecular description of ion adsorption using an
192 equilibrium approach that defines surface species, chemicals reactions, equilibrium constants, mass
193 balances, and charge balances (Goldberg, 2013). This molecular approach was developed by
194 Schindler et al. (1976) who have demonstrated that the adsorption of a cation on a positively
195 charged surface is possible. These conceptual models take into account both the intrinsic affinity of
196 surface sites for solutes and the coulombic interaction between the charged surface and the
197 dissolved ions (Davis et al., 1978; Hayes and Leckie, 1987; Schindler et al., 1976).

198 Several chemical SCM have been developed during the last three decades to describe
199 potentiometric titration and metal adsorption data at the oxide-metal solution interface and have

200 been very successful in describing adsorption processes (Goldberg and Criscenti, 2007). These
201 models are distinguished by differences in their respective molecular hypotheses: each model
202 assumes a particular interfacial structure, resulting in the consideration of various kinds of surface
203 reactions and electrostatic correction factors to mass law equations (Davis and Kent, 1990). The
204 description of the most common models is supplied in the Appendix.

205

206 **3 Modelling approaches**

207 *3.1 Overview of existing models for Cs sorption on clay minerals*

208

209 A wide range of multi-site ion exchange models were proposed in the literature to simulate
210 experimental data and understand the non-linear Cs sorption behavior on clay minerals. For illite,
211 cation exchange (CE) models with three exchange sites are usually considered (Benedicto et al.,
212 2014; Bradbury and Baeyens, 2000; Brouwer et al., 1983; Cornell, 1993; Fuller et al., 2014;
213 Marques Fernandes et al., 2015; Missana et al., 2014b; Steefel et al., 2003). Such models have also
214 been reported for Cs sorption onto smectites and kaolinite (Missana et al., 2014b). Furthermore,
215 one-site CE models (Bradbury and Baeyens, 2010; Chen et al., 2014) and two-site models can also
216 be found for smectites (Liu et al., 2004; Missana et al., 2014b; Poinssot et al., 1999; Zachara et al.,
217 2002).

218 Poinssot et al. (1999) proposed a two-site CE model for illite, but they only considered the
219 exchange reactions between Na^+ , or K^+ and H^+ . Savoye et al. (2012) have developed a 5-sites CE
220 model for which selectivity coefficient values were calculated from Poinssot et al. (1999).
221 Selectivity coefficients of CE models are usually obtained from batch-type experiments, measured
222 with pure mineral phases in dispersed state at specific physicochemical conditions (nature of
223 exchange cations, ionic strength, pH, contact time, concentration of some ion). The major
224 differences between these models are the number of Cs sorption sites and the capacities of each

225 site. These semi-empirical parameters calculated for each study fall in a relatively close range,
226 considering that the uncertainty is up to $\pm 0.2 \text{ L kg}^{-1}$ (log unit) due to measurement uncertainty.
227 Missana et al. (2014a) also note that selectivity coefficients vary with ionic strength for different
228 clays.

229 The analysis of these models shows that only the “generalized caesium sorption” (GCS)
230 model of Bradbury and Baeyens (2000) has been applied to different experimental data (Brouwer
231 et al. (1983), Comans et al. (1991), Staunton and Roubaud (1997), and Poinssot et al. (1999)). No
232 other data compilation and/or model comparison on experimental sorption data has been
233 performed, leaving a wide range of choices for cation exchange modelling of Cs sorption onto illite
234 without clear selection criteria. Note that the model proposed by Poinssot et al. (1999) corresponds
235 to the GCS model without the planar site.

236 A large number of studies suggest that Cs adsorption depends on pH for various clay
237 minerals (Akiba et al., 1989; Cornell, 1993; Torstenfelt et al., 1982). For instance, Poinssot et al.
238 (1999) observed a small decrease in Cs sorption on illite, at pH lower than 4 and at low Cs
239 concentration ($<10^{-8} \text{ M}$), where sorption onto the FES dominates (Fuller et al., 2014). This pH-
240 dependent adsorption of trace elements, such as Cs^+ can easily be explained by analogy with the
241 sorption properties of oxides (Kraepiel et al., 1999). The FES result from broken bonds at the
242 edges of clay crystals and hydroxyl groups, and contribute to the development of negative and
243 positive charge (Cornell, 1993). They can potentially react with ions in solution to yield surface
244 complexes (Koretsky, 2000) like pure oxides phases (Angove et al., 1997; Ikhsan et al., 1999;
245 Kraepiel et al., 1999; Lackovic et al., 2003; Zachara and McKinley, 1993). These surface reactions
246 may be described with a SCM, in which the FES are commonly called ‘variable charge sites’ and
247 noted $\equiv\text{SOH}$ (Davis and Kent, 1990; Davis and Leckie, 1978; Sposito, 1984).

248 Metal and radionuclide (Ni^{2+} , Zn^{2+} , Eu^{3+} , Al^{3+} , Cu^{2+} , Cd^{2+} , Pb^{2+} , Sr^{2+} , Co^{2+} , Cs^+ , Ca^{2+} , Na^+ ,
249 Se^{2-} , F^- , Br^- and I^-) adsorption on clay minerals was successfully modeled in many studies using an
250 approach combining cation exchange and surface complexation (Baeyens and Bradbury, 1997;

251 Bruggeman et al., 2010; Charlet et al., 1993; Du et al., 1997; Gu and Evans, 2007, 2008; Gutierrez
252 and Fuentes, 1996; Lund et al., 2008; Mahoney and Langmuir, 1991; Missana et al., 2009; Missana
253 and Garcia-Gutiérrez, 2007; Stadler and Schindler, 1993; Tertre et al., 2006; Weerasooriya et al.,
254 1998; Zachara and McKinley, 1993). This approach usually involves two distinct types of surface
255 sites at the interface between the solid and the solution: (i) $\equiv X^-$ groups bearing a permanent
256 negative charge which account for the cation exchange part of sorption. Cations (e.g., Na^+ , K^+
257 Ca^{2+}) binding at this type of surface sites occurs through electrostatic interactions ; (ii) amphoteric
258 $\equiv SOH$ groups on FES with high affinity for metals and radionuclides but low site density which
259 control protonation/deprotonation (i.e., pH-dependent) which account explicitly for the adsorption
260 of background electrolyte.

261 Several authors have proposed sorption on amphoteric surface hydroxyl as a suitable
262 mechanism for pH-dependent sorption, as opposed to mineral dissolution or H^+ exchange
263 (Baeyens and Bradbury, 1997). Studies therefore considered the surface complexation mechanism
264 for the description of trace Cs sorption on amphoteric hydroxyl groups of montmorillonite (on
265 $\equiv SOH$ edge sites). For instance, Gutierrez and Fuentes (1996) used the triple layer model (TLM) to
266 simulate the adsorption of Cs^+ onto Ca-montmorillonite. Two types of adsorption sites were
267 considered to be responsible for Cs adsorption: interlayer and (frayed) edge sites. The dominant
268 mechanism of adsorption was identified as specific adsorption in the edge sites which, although
269 composing only 5% of the total adsorption sites, accounted for as much as 94% of Cs adsorption.
270 Silva (1979) proposed a combined approach for montmorillonite: (i.) a two-site ($\equiv SOH$ and $\equiv TOH$
271 groups) with a 1-pK Double Layer Model (DLM) and (ii.) a one-site cation exchange model. The
272 authors have considered the complexation of Cs^+ and the Na^+ on $\equiv SOH$ and $\equiv TOH$ edges groups.
273 This approach requires the fitting of 12 parameters which represents a major drawback. Hurel et al.
274 (2002) have used another surface complexation model for the adsorption of cations (Cs^+ , Na^+ , K^+ ,
275 Ca^{2+} and Mg^{2+}) on bentonite. Cs sorption was modeled on the silanol edge sites as well as cation
276 exchange. Wang et al. (2005), also found that at high pH Cs^+ sorption onto bentonite was

277 dominated by surface complexation.

278 To model Cs sorption on MX-80 bentonite, Montavon et al. (2006) used the simplified two-
279 pK non-electrostatic model (NEM) from Bradbury and Baeyens (1997) to describe the pH-
280 dependent interactions with clay edge surfaces in combination with a cation exchange in the
281 interlayer and on the basal plane surfaces. This model is also very “expensive” in terms of number
282 of adjustable parameters (four SCM and two cation exchange sites). The authors noted that the
283 parameters were only applicable for the conditions of their study. Bradbury and Baeyens (2010)
284 proposed a unique selectivity coefficient ($K_c=15$) for all the Cs concentration range, which does
285 not allow to account for the non-linearity of Cs sorption isotherm on MX-80 bentonite.

286 In summary, existing models are usually successful in reproducing a specific set of
287 experiments but are often complex in terms of the number of adjustable parameters and
288 implementation (2-pK TLM, 2-pK DLM...). Moreover, the applicability of these models is
289 restricted to the data on which they were calibrated and their ability to predict Cs sorption in
290 heterogeneous environments or natural systems appears therefore limited.

291

292 3.2 *Building a new, robust and parsimonious model for caesium sorption*

293

294 To model Cs sorption onto illite, montmorillonite and kaolinite, we tried to minimize the
295 number of adjustable parameters to fit the largest number of sorption data sets by combining:

- 296 • a surface complexation model which is used to describe the adsorption of cations (protons,
297 Cs, and cations from the background electrolyte) on a single FES site ($\equiv\text{SO}^{-0.5}$) with a high
298 affinity for Cs but with very low site density,
- 299 • and an Ion Exchange model (IE), which is used to simulate the adsorption of cations on
300 permanent negatively charged sites of planar surfaces of clay minerals, including outer-

301 basal and interlayers sites ($\equiv X^-$), which have a low affinity for Cs but a high site density
302 (i.e. large contribution to the global CEC). These sites will be called in this study
303 “exchange sites”.

304 Our modelling methodology is similar to that introduced by Tournassat et al. (2013) as a
305 “minimalist” modelling approach for sorption on high energy sites of montmorillonite edge
306 surfaces, which attempts to minimize the number of adjustable parameters for modeling the largest
307 number of sorption data and by using a 1-pK approach. This is an advantage over multi-site
308 models available in the literature for Cs sorption on illite and montmorillonite, which have more
309 than two sorption sites, with the exception of the model used by Poinssot et al. (1999).

310 As in many other studies (Baeyens and Bradbury, 1997; Bruggeman et al., 2010; Charlet et
311 al., 1993; Gu and Evans, 2007; Gutierrez and Fuentes, 1996; Lund et al., 2008; Mahoney and
312 Langmuir, 1991; Missana and Garcia-Gutiérrez, 2007; Stadler and Schindler, 1993; Tertre et al.,
313 2006; Zachara and McKinley, 1993), we used a SCM to model the adsorption of cations on surface
314 hydroxyl groups ($\equiv SOH$) localized in FES for two reasons : (i) these models allow a mechanistic
315 description based on the chemical nature of the adsorption process, and (ii) they take into account
316 both the intrinsic affinity of surface sites for solutes and the coulombic interaction between the
317 surface charge and the dissolved cations (Davis et al., 1978; Hayes and Leckie, 1987; Schindler et
318 al., 1976). We have chosen the Double Layer Model (DLM) (Dzombak and Morel, 1990; Huang
319 and Stumm, 1973; Stumm et al., 1970), as several previous authors (Bradbury and Baeyens, 1997;
320 Hoch and Weerasooriya, 2005; Silva, 1979; Sverjensky and Sahai, 1996; Tertre et al., 2006;
321 Tombácz and Szekeres, 2004; Wanner et al., 1996), because it is one of the simplest model
322 available with fewer fitting parameters than other surface complexation models. Moreover, the
323 values of reaction constants depend only on the nature of the solid and the adsorbing solute (not on
324 pH, adsorbate concentration, ionic strength and solution composition; (Koretsky, 2000). For more
325 details, see the DLM description in the Appendix.

326 We selected the 1-pK approach (Van Riemsdijk (1979) ; Bolt and Van Riemsdijk (1982))

327 combined with the DLM, using only one charge reaction, as opposed to the classical 2-pK-DLM
328 approach (adsorbing protons in two consecutive steps, each having its own affinity and charge
329 properties). Hiemstra et al. (1989) and Koopal (1993) suggested that the 1-pK model may be used
330 as a physically most realistic simplified model to describe the protonation/deprotonation reactions
331 of (hydr)oxides, when the structure of a surface is not well known (Avena and De Pauli, 1996).

332 Surface roughness was not explicitly treated in this paper. We consider that roughness is
333 embedded into the surface area value that we take from the literature.

334 3.3 *Geochemical software and statistical analysis*

335

336 The modelling of the adsorption of Cs and co-ions using our 1-pK DL/IE model was
337 performed with PHREEQC v2.18 (Parkhurst and Appelo, 1999). The goodness of fit of the
338 simulations was optimized with the adjusted determination coefficient (R^2_{adj} , Appendix, A.2) and
339 by minimizing the difference between predicted and measured values, using the Root Mean Square
340 Deviation (RMSD). Further details are supplied in the Appendix.

341

342 **4 Results and discussion**

343 4.1 *Calibration of the 1-pK DL/IE model for Cs sorption*

344

345 The mineralogical properties and parameters values used in the 1-pK DL/IE model to
346 simulate sorption data for Cs under the conditions defined in **Table 1** onto the three pure clay
347 minerals (illite, montmorillonite and kaolinite) are shown in **Table 6**.

348 The $\log K_H$ value for illite was calculated using Eqn. A.7 with the $\log K_{a1}$ and $\log K_{a2}$ ($\log K_{SC}$
349 reported by Liu et al. (1999). Note that the calculated value is comparable to the pH of zero point
350 charge (pH_{ZPC}) of other types of illite (Du et al., 1997; Liu et al., 1999). Similarly, the $\log K_H$ value

351 used for kaolinite was derived from the 2-pK DLM developed by Hoch and Weerasooriya (2005)
352 for the KGa kaolinite. Again, the calculated value is close to the pH_{ZPC} (4.9) determined by the
353 same authors. In the case of montmorillonite, the $\log K_{\text{H}}$ value was assumed to be equal to pH_{ZPC}
354 reported by Ijagbemi et al. (2009).

355 The equilibrium constants used to simulate cation adsorption on FES ($\log K_{\text{SOM}}$) are shown in
356 **Table 6**. Some values are taken directly from the literature while others were determined by fitting
357 the experimental data listed in **Table 1**.

358 Firstly, the $\log K_{\text{SONa}}$ values for illite and montmorillonite were assumed to be equal to the
359 values reported by Mahoney and Langmuir (1991) (2.2 and -2, respectively for $\equiv\text{SOH}$ site).
360 Secondly, the $\log K_{\text{SOCs}}$ values for both clay minerals were determined by concurrently fitting the
361 1-pK DL/IE model to the experimental data (5.2 and 4.3, respectively), and the $\log K_{\text{SONa}}$ values
362 were revised to -1.8 and -1.3, respectively. Note that revised $\log K_{\text{SONa}}$ values are close to the
363 values proposed by Duputel et al. (2014) for illite (-1.6) and Lumsdon (2012) for both minerals (-
364 1). Similarly, the $\log K_{\text{SOK}}$ value for both clay minerals, and $\log K_{\text{SONH}_4}$ value for illite were
365 determined using the same method.

366 The values of Ca, Mg and Sr surface complexation constants for illite, montmorillonite and
367 kaolinite have been taken directly from literature: Bradbury and Baeyens (2005) for $\equiv\text{S}^{\text{w}2}\text{OH}$ sites,
368 Mahoney and Langmuir (1991) for $\equiv\text{SOH}$ sites and Riese (1982) for $\equiv\text{SiOH}$ sites, respectively.
369 The $\log K_{\text{SONa}}$ and $\log K_{\text{SOK}}$ values for kaolinite, were taken directly from Mahoney and Langmuir
370 (1991) for $\equiv\text{SOH}$ sites and Jung et al. (1998) for $\equiv\text{SiOH}$ sites. The $\log K_{\text{SONH}_4}$ value for kaolinite
371 was assumed to be equal to $\log K_{\text{SOK}}$ value.

372 The selectivity coefficients (K_c) values for the cation exchange model were generated by
373 starting from literature values and fitting the 1-pK DL/IE model to published Cs sorption data from
374 **Table 1**. Note that the values obtained are close to selectivity coefficients of existing cation
375 exchange models.

376 The values for the site capacity of the permanent negatively charged sites ($\equiv X^-$), i.e. mainly
377 the CEC, and the specific surface area (SSA) for the three studied clay minerals are within the
378 literature range. The $\equiv SOH^{0.5}$ site density value used for illite was $2.7 \cdot 10^{-3}$ sites $nm^{-2} \approx 0.43$ meq
379 kg^{-1} , assumed to be equal to the mean FES site density values reported by Missana et al. (2014b)
380 and Benedicto et al. (2014) (0.58 and 0.29 meq kg^{-1} , respectively). For montmorillonite and
381 kaolinite, the values of $\equiv SOH^{0.5}$ site density were obtained by fitting the experimental data.
382 Furthermore, $\equiv SOH^{0.5}$ site density value for montmorillonite ($3.6 \cdot 10^{-5}$ sites $nm^{-2} \approx 4 \cdot 10^{-2}$ meq kg^{-1}
383 with $SSA = 800$ $m^2 g^{-1}$) was in the same order of magnitude than the value proposed by Montavon
384 et al. (2006) for exchangeable interlayer cations site in MX-80 montmorillonite ($2 \cdot 10^{-2}$ meq kg^{-1}).
385 For kaolinite, the value of $\equiv SOH^{0.5}$ site density ($1.5 \cdot 10^{-4}$ sites $nm^{-2} \approx 2.4 \cdot 10^{-3}$ meq kg^{-1}) is close to
386 the value suggested by Missana and al. (2014b) for the FES ($5 \cdot 10^{-3}$ meq kg^{-1}).

387

388 4.1.1 Cs sorption on illite

389 Results obtained with the 1-pK DL/IE model (continuous red lines) for illite are shown in
390 **Figure 1**. The dotted lines represent results calculated by the different sorption models proposed in
391 the literature. It can be seen that the models globally reproduced well the decrease of K_d with
392 increasing dissolved Cs concentration, as well as the difference between Na and Ca-illite (**Figure**
393 **1a, b, g, h**) versus K and NH_4 -illite (**Figure 1c-f and i**) and ionic strength of the solution. These
394 well-known differences reflect the selectivity of illite (especially on FES) for cations according to
395 the following order of affinity: $Cs^+ > NH_4^+ > K^+ > Na^+ > Sr^{2+} > Ca^{2+}$ and Mg^{2+} , (Appelo and
396 Postma, 1993; Brouwer et al., 1983; Dyer et al., 2000; Staunton and Roubaud, 1997). The model
397 results show, in particular, that K^+ competes most effectively with Cs^+ adsorption (Eberl, 1980;
398 Sawhney, 1972). Concerning the effect of ionic strength (I), the modelling also confirms that the
399 K_d increases as the ionic strength of the solution decreases, which is more marked for Na and K-
400 illite than for Ca-illite.

401 The contribution of each type of sorption site is illustrated for the 1pK DL/IE model in
402 **Figure A.4** for Na-illite at pH= 7 and I = 0.1 mol L⁻¹. At low total Cs concentrations ([Cs]_{tot} = 10⁻⁹
403 - 10⁻⁷ mol L⁻¹), sorption is dominated by FES (≡SOCs^{0.5}) which have a high affinity for Cs (logK_d
404 values are higher), a very small concentration (0.43 meq kg⁻¹), and are therefore quickly saturated.
405 At medium and high total Cs concentrations ([Cs]_{tot} = 10⁻⁷ - 10⁻⁴ mol L⁻¹), sorption occurs on the
406 remaining sites on planar surfaces (exchange sites) which have a higher capacity (190-225 meq kg⁻¹)
407 and a lower selectivity (i.e. lower logK_d values).

408 Results also show that the goodness of fit of the 1pK DL/IE model is globally better than the
409 other models (dotted lines). These performances were obtained using of a unique set of parameters
410 to simulate all the experimental data. For example, in the case of Na-illite at I = 0.01 mol L⁻¹
411 (**Figure 1a**, data from Benedicto et al. (2014)), the 1pK DL/IE model and the 3-sites cation
412 exchange model developed by Benedicto et al. (2014), which was specifically calibrated for these
413 experimental data, adequately reproduce the measured logK_d (see **Table 7**). However, their model
414 is out-performed by the 1pK DL/IE model when simulating experimental data from other authors.
415 The same behavior is observed in the case of K-illite (**Figure 1c and 1e**, data from Brouwer et al.
416 (1983)): the calculated logK_d values from the 1-pK DL/IE model and the 3-sites cation exchange
417 model proposed by Brouwer et al. (1983) both fall within the experimental errors in the whole
418 range of Cs concentrations, and have the best goodness of fit (RMSD values of 0.21 and 0.11 L kg⁻¹
419 ¹, and R²_{adj} of 1.07 and 1.11, respectively). Globally, the cation exchange models provide a
420 satisfactory fit only for a limited Cs concentration range: some underestimate and other
421 overestimate strongly Cs sorption as suggested by RMSD values ≥ 0.65 L kg⁻¹.

422 Further all the models underestimate Cs sorption on K-illite for high ionic strength and high
423 Cs concentrations ([Cs]_{eq} > 10⁻⁴ mol L⁻¹) (**Figure 1.f**). This may be explained by the expansion of
424 the illite layers and the increase of the CEC (de-collapse) which facilitate Cs uptake, as postulated
425 by Benedicto et al. (2014) (CEC_{decollapsed} = 900 meq kg⁻¹), or by the existence of a low amount of
426 another clay mineral, which becomes relevant at high Cs concentrations when saturation of sorbed

427 Cs is reached for illite.

428 The global RMSD and R^2_{adj} index, i.e. the average values of RMSD and R^2_{adj} corresponding to
429 all the Cs sorption isotherms simulated for illite ($n = 95$ observations) are shown in **Table 7**. The
430 model proposed in this study performs well for all the conditions (smallest global RMSD index
431 value of 0.25 L kg^{-1} and $R^2_{\text{adj}} = 0.83$), which confirms its accuracy and robustness. The model
432 proposed by Benedicto et al. (2014) also reproduces reasonably well the experimental sorption
433 isotherms (RMSD = 0.27 L kg^{-1} and $R^2_{\text{adj}} = 0.8$) but does not account for the effect of NH_4^+
434 concentrations and pH, and requires small adjustments of the parameters in order to improve the fit
435 in each condition. The global performance of the other models is lower.

436 The simulations of Cs sorption on illite as function of pH using the 1-pK DL/IE model and
437 the model from Poinssot et al. (1999) are shown in **Figure 2**. Our model reproduced the
438 experimental data of Poinssot et al. (1999) with a better global accuracy (global RMSD index of
439 0.13 L kg^{-1} vs. 0.25 L kg^{-1} for Poinssot et al., 1999) for the range of ionic strength from 1 to 0.01
440 mol L^{-1} (and particularly at the lower value). At low total Cs concentrations ($[\text{Cs}]_{\text{tot}} < 10^{-8} \text{ mol L}^{-1}$),
441 the simulations indicate that sorption on illite weakly depends on pH (sorption decreases only at
442 pH lower than 4) and is dominated by the $\equiv\text{SO}^{-0.5}$ site, as it is also suggested by Poinssot et al.
443 (1999).

444

445 4.1.2 Cs sorption on montmorillonite

446

447 Only a few Cs sorption experimental data and models have been published for
448 montmorillonite. Moreover, these models have a large number of adjustable parameters. Note that
449 the data from experiments with MX-80 bentonite have not been used for the calibration of the
450 model but are considered hereafter for its evaluation (cf. section.4.2).

451 The CEC of purified SWy-montmorillonite (838 meq kg^{-1}) measured by Gorgeon (1994) was

452 similar to that reported by Staunton and Roubaud (1997) (940 meq kg⁻¹). In our model we assumed
453 that the CEC for montmorillonite is equal to value proposed by Baeyens and Bradbury (1997) for
454 conditioned Na-montmorillonite, i.e., 870 meq kg⁻¹.

455 **Figure 3a** represents the sorption isotherms of montmorillonite initially saturated with Na, K
456 and Ca, reported by (Staunton and Roubaud, 1997) and **Figure 3b** shows the effect of pH on Cs
457 sorption for montmorillonite. **Figure 3** shows that Cs sorption behavior for montmorillonite is
458 similar to that of illite. The sorption isotherms are non-linear especially for Ca-montmorillonite:
459 little changes in logK_d values for Na- and K-montmorillonite were observed, whereas the value of
460 logK_d for Ca-montmorillonite decreases progressively with increasing Cs concentration, which
461 suggests some heterogeneity of the montmorillonite exchange sites and confirms the order of
462 affinity for cations according to the sequence Cs⁺ > K⁺ > Na⁺ > Ca²⁺ (Appelo and Postma, 1993;
463 Brouwer et al., 1983; Dyer et al., 2000; Staunton and Roubaud, 1997).

464 The analysis of the sorption distribution shows that Cs is adsorbed predominantly on
465 exchange sites for Na- and K-kaolinite. In the case of Ca-illite the distribution of adsorbed Cs
466 follows the same behaviors of illite (**Figure A.5**). Measured logK_d were appropriately reproduced
467 using the 1-pK DL/IE model (**Table 6**) with the global RMSD index of 0.09 L kg⁻¹ for the sorption
468 isotherms and 0.06 L kg⁻¹ for the sorption on FES.

469

470 4.1.3 Cs sorption on kaolinite

471

472 Only a few experimental data have been found in literature for the kaolinite, and only the two-site
473 cation exchange model of Missana et al. (2014b) was compared to the model developed here.

474 Gorgeon (1994) used a synthetic kaolinite (Sigma) which has similar properties than KGa-1-
475 b. The CEC values of KGa-1-b and Sigma kaolinite are equal to 20 and 28 meq kg⁻¹ respectively
476 and their total specific surface area (SSA) are equal to 10 and 13.2 m² g⁻¹, respectively.

477 **Figure 4** shows that Cs sorption on kaolinite also follows the same pattern as illite, and was
478 appropriately simulated using the 1-pK DL/IE model with a global RMSD index of 0.09 for
479 sorption isotherms (**Figure 4a**) and 0.15 L kg^{-1} for the sorption edge (**Figure 4b**). Cs sorption was
480 slightly underestimated at pH 2-3 in the case of Na-kaolinite at high ionic strength $I = 1 \text{ mol L}^{-1}$
481 and $[\text{Cs}]_{\text{tot}} = 10^{-7} \text{ mol L}^{-1}$. Our model (**Figure 4a**, solid lines) and the cation exchange model
482 proposed by Missana et al. (2014b) (**Figure 4a**, dotted lines) show a similar goodness of fit for
483 sorption isotherms.

484 The sorption isotherms are sensitive to $[\text{Cs}]$ except for kaolinite initially saturated with K and
485 NH_4 , for which sorption was constant in a very wide range of Cs concentrations. A non-linear
486 change in $\log K_d$ values was observed on a small range (0.5 log unit) for Na and Ca-kaolinite. Note
487 that both models failed to reproduce the high K_d value at high Cs concentration ($\text{Cs} = 10^{-3} \text{ mol/L}$).

488 Simulations predict that sorption is dominated by exchange site ($\equiv\text{X}^-$) for all Cs
489 concentrations range, especially for K-kaolinite (95%). The model could therefore be reduced to a
490 single sorption site, as proposed by Shahwan and Erten (2002). However Missana et al. (2014b)
491 suggested that the existence of high affinity sites (FES) on kaolinite and smectite is most probably
492 due to the existence of traces of micaceous or interstratified minerals.

493 **Figure A.6** shows (a) the distribution of sorbed Cs on sorption sites ($\equiv\text{SO}^{-0.5}$ and $\equiv\text{X}^-$) and
494 (b) the evolution of species distribution on $\equiv\text{SO}^{-0.5}$ site (%) as function of pH as calculated by the
495 1-pK DL/IE model, for Na-kaolinite at $I = 1 \text{ mol L}^{-1}$ and $[\text{Cs}]_{\text{tot}} = 10^{-7} \text{ mol L}^{-1}$. At lower pH, the Cs
496 is totally adsorbed by surface planar sites (exchange sites) because $\equiv\text{SO}^{-0.5}$ sites are initially
497 saturated by H^+ ($\equiv\text{SOH}^{0.5}$). These sites progressively deprotonated in favor of the adsorption of Na
498 and Cs when pH increases from 2 to 5, resulting in an increase of the contribution of FES sites to
499 total Cs sorption. Above pH 5 ($\log K_H = 5$), the $\equiv\text{SOH}^{0.5}$ species is negligible and the occupancy of
500 $\equiv\text{SO}^{-0.5}$, $\equiv\text{SOCs}^{0.5}$, $\equiv\text{SONa}^{0.5}$ species represent 3%, 25% and 71%, respectively.

501 4.1.4 Effect of electrostatic correction factor

502

503 The effect of the electrostatic component was also estimated by applying the model without
504 the electrostatic correction factor. The contribution of electrostatic term in Cs sorption on the
505 studied clay minerals was negligible, and the mean difference between K_d values calculated with
506 and without electrostatic term is in order of 0.01%. The maximum of this difference represented
507 only 5% of the mean error of K_d values ($\pm 0.2 \text{ L kg}^{-1}$) in very isolated cases, which means that the
508 electrostatic correction always remains within the uncertainty range. This ascertainment is in
509 agreement with the observation of Bradbury and Baeyens (1997) who showed that the contribution
510 of the electrostatic term was lower than the range of error. In this study Cs surface complexation at
511 edge sites is treated with very larger stability constants, frequently higher than those for proton
512 adsorption. So as a consequence the chemical contribution is in most cases so strong that it
513 overcomes competition with protons and repulsive electrostatics. So the double layer model can be
514 to reduce to a non-electrostatic model. This would also be numerically more easily included into a
515 transport code. Tournassat et al. (2013) have also demonstrated, by testing surface complexation
516 models available in the literature, that a non-electrostatic model is the most efficient model, in
517 terms of simplicity and accuracy, to represent correctly the physical nature of the metal/clay
518 surface interactions.

519

520 4.1.5 Performance of the model calibration

521

522 **Figure 5** represents the values of $\log K_d$ simulated by 1-pK DL/IE model as function of
523 $\log K_d$ measured for illite, montmorillonite and kaolinite. This figure confirms that the model
524 proposed in this study (1-pK DL/IE model) successfully reproduces the experimental data
525 available in the literature for Cs sorption on these three clay minerals in a wide range of physico-
526 chemical conditions (ionic strength, pH, and solid/liquid ratio). Moreover, these results were

527 obtained with a unique set of parameters for each mineral whereas the selectivity coefficients are
528 often adjusted, for each curve, within a range of $\log K_c \pm 0.2 \text{ L kg}^{-1}$ by authors who want to
529 optimize the fit on their own data (McBride, 1979; Missana et al., 2014a; Staunton and Roubaud,
530 1997; Wahlberg and Fishman, 1962). The few points further from the 1:1 line correspond mostly
531 to the K_d values obtained for the highest Cs concentrations where all the models have difficulties to
532 simulate the experimental data.

533

534 4.2 *Model evaluation in natural material made of a single type of clay mineral*

535

536 The 1-pK DL/IE model was first tested in simple cases with other experimental data obtained
537 on clay materials (Rojo Carbonero clay, MX-80 bentonite, Boda claystone and Hanford sediment)
538 containing only one type clay mineral (illite or montmorillonite). The mineralogical properties of
539 these clay materials, the content and type of clay mineral are listed in **Table 2**. The experimental
540 conditions (i.e. ionic strength and major cationic environments) used in these experiments are
541 listed in **Tables 2** and **3**. The 1-pK DL/IE model was applied directly without changing the
542 parameters adjusted previously (**Table 6**).

543 For the Rojo Carbonero (RC) clay, the illite content considered was 57%, which result in a
544 capacity of exchange site $\equiv X^-$ of the 1-pK DL/E model equal to $127.9 \text{ meq Kg}^{-1}$. This hypothesis
545 slightly overestimates the CEC measured by Missana et al. (2014b) (110 meq Kg^{-1}). In the case of
546 MX-80 bentonite, the content of montmorillonite was assumed to be 81.4% (Karnland, 2010), and
547 the CEC equal to 870 meq kg^{-1} . The resulting capacity of the exchange site $\equiv X^-$ of the 1-pK DL/E
548 model was equal to 705 meq kg^{-1} in agreement with the CEC suggested by Bradbury and Baeyens
549 (2011) ($787 \pm 48 \text{ meq kg}^{-1}$) and the exchangeable interlayer cations site ($\equiv X$) for the model
550 proposed by Montavon et al. (2006) (640 meq kg^{-1}). For the content and type clay mineral in the
551 Hanford sediment (precise data not available), we assumed that only illite was present with a

552 capacity of exchange site $\equiv X^-$ equal to the measured CEC as proposed by Zachara et al. (2002).
553 These authors performed sorption experiments with different ranges of particle size (63 μm and
554 125–250 μm) to demonstrate that the fraction of high affinity sites was independent of size fraction
555 and mineralogical composition. They noted that the concentration of high-affinity sites was very
556 low, making it difficult to identify their location. The edge site density ($\equiv\text{SOH}$) used in modelling
557 was calculated from the fine fraction of sediment ($<63 \mu\text{m}$), which was 9.2% (i.e. $2.5 \cdot 10^{-4}$ sites nm^{-2}).
558 2). Our modelling results for Hanford sediments confirmed the dependence of the caesium sorption
559 (K_d) on Cs and competing cation (Na^+ , K^+ , and Ca^{2+}) concentration.

560 Concerning the Boda clay sample (Breitner et al., 2014), (Marques Fernandes et al., 2015)
561 measured a CEC of 113 meq kg^{-1} . This value is consistent with the measured illite fraction of 50%
562 (considering the reference CEC value of 225 meq kg^{-1} for illite).

563 The experimental data were successfully predicted with the 1-pK DL/IE model according to
564 the global $\text{RMSD} = 0.21 \text{ L kg}^{-1}$ and $R^2_{\text{adj}} = 0.90$. The $\log K_d$ values simulated are highly correlated
565 (95%) to the $\log K_d$ values measured for the studied simple clay materials (**Figure 6**).

566 **Table 8** summarizes the RMSD and R^2_{adj} index calculated for each material. These results
567 show the good performance of the 1-pK DL/IE model to predict Cs sorption on natural simple
568 clayey materials when complex natural solutions are used (i.e. presence of several competitive
569 cations). Cs sorption isotherms for Rojo Carbonero clay, MX-80 bentonite, Boda claystone and
570 Hanford sediment are supplied in the Appendix (**Figures A.7 to A.10**).

571 For the RC clay, the results obtained with the 1-pK DL/IE model are a little better than those
572 obtained with the three-site cation exchange model proposed by Missana et al. (2014b) ($\text{RMSD} =$
573 0.18 L kg^{-1} and $R^2_{\text{adj}} = 0.70$). Also, for the Hanford sediment a similar goodness of fit is obtained for
574 the two-cation exchange model proposed by Zachara et al. (2002) (data not shown, $\text{RMSD} = 0.22 \text{ L}$
575 kg^{-1} and $R^2_{\text{adj}} = 0.44$). In the case of Boda clay, the modelling results show that the 1-pK DL/IE
576 model yields a better estimation for Cs sorption than the GCS model (Bradbury and Baeyens,

577 2000), in terms of accuracy (RMSD= 0.64 L kg⁻¹).

578

579 4.3 Cs sorption model evaluation for mixed-clays natural materials

580

581 Cs sorption was simulated for five clay materials in “prediction mode” by considering only
582 illite as the main reactive clay or using a component additivity approach (CA) to represent the
583 different clay minerals (illite, montmorillonite and kaolinite). Again, the 1-pK DL/IE model was
584 used with the parameters obtained previously for each clay mineral (**Table 6**). In all simulations,
585 only the clay mineral fractions and pore water compositions were used as input parameters (**Table**
586 **4** and **5**). The content and the type of clay mineral for each mixed-clay material are shown in
587 **Table 9**.

588

589 4.3.1 FEBEX clay

590

591 **Figure 7** shows that Cs sorption isotherms for Na, Ca and K-FEBEX clay were satisfactorily
592 simulated with the 1-pK DL/IE model using the CA approach considering 93% of montmorillonite
593 – illite mixed layers with 15% of interstratified illite and 78% of montmorillonite (global RMSD of
594 0.18 L kg⁻¹ and R²_{adj} index of 0.89, calculated from 12 sorption curves and 201 measured points).
595 Insignificant differences were calculated between the global RMSD for the 3-site IE model
596 proposed by Missana et al. (2014a) (≈ 10%).

597 In the case of Na-FEBEX clay at I = 0.001 mol L⁻¹, to obtain a good agreement with
598 experimental data, we followed Missana et al. (2014a) and assumed that the concentration of
599 exchange site ($\equiv X^-$) was reduced by 30% considering that illite affected only the FES (RMSD =
600 0.18 L kg⁻¹). For the modelling of Cs sorption on Na, Ca and K-FEBEX as function of pH (**Figure**

601 **A.11**), a good fit was also obtained ($\text{RMSD} = 0.19 \text{ L kg}^{-1}$).

602 The K_d values predicted by the model are slightly overestimated for Na-, Ca- and K-FEBEX
603 clay at 0.01 mol L^{-1} , 0.3 mol L^{-1} (only for lower Cs equilibrium concentrations) and 0.1 mol L^{-1}
604 (only for Cs equilibrium concentrations between 10^{-7} and $10^{-5} \text{ mol L}^{-1}$) (**Figure 7**). Similar results
605 were observed in pH-edge simulations (**Figure A.11**).

606 For the Na-FEBEX clay, with $I = 0.01 \text{ mol L}^{-1}$, at low Cs concentration, sorption mostly
607 occurs onto illite, especially on FES (82%, **Figure 8**, dotted red line). At higher Cs loading,
608 sorption on montmorillonite exchange sites becomes dominant (solid blue line). At intermediate Cs
609 concentration, sorption is nearly equivalent between illite and montmorillonite. The contribution of
610 montmorillonite FES is always very low and can be neglected. Although the illite fraction in
611 FEBEX clay does not exceed 15%, it has a very important contribution to Cs sorption, particularly
612 at low Cs concentration where sorption on FES dominates (Brouwer et al., 1983; Eberl, 1980;
613 Francis and Brinkley, 1976; Jackson, 1963; Maes and Cremers, 1986; Poinssot et al., 1999; Rich
614 and Black, 1964; Sawhney, 1972; Zachara et al., 2002). It also has a critical role in the non-linear
615 character of the Cs sorption (Missana et al., 2014a; Missana et al., 2014b).

616 The model reproduces the main features of the experimental data: (i) Cs sorption increases
617 by decreasing the ionic strength of the electrolyte from 0.1 to 0.001 mol L^{-1} , (ii) Cs sorption is non-
618 linear resulting from the existence of two different types of sorption site, as observed by different
619 authors (Eliason, 1966; Missana et al., 2004; Staunton and Roubaud, 1997; Wahlberg and
620 Fishman, 1962). Missana et al. (2014a) explain this behavior by the existence of interstratified
621 illite-montmorillonite mixed-layers (with 10-15% of illite layers).

622

623 4.3.2 San Juan clay

624

625 Interestingly, the natural San Juan clay material contains a mixture of the three “reference”

626 clay minerals: illite (30 - 50%), kaolinite (10 - 15%) and montmorillonite (5 -10%) (Missana et al.,
627 2014b). Missana et al. (2014b) have simulated Cs sorption using the CA approach with multi-sites
628 cations exchange models assuming contributions from illite, kaolinite and montmorillonite of 40%,
629 12% and 10%, respectively.

630 Cs sorption isotherms were obtained in different conditions: (case A) several simple
631 electrolyte solutions (0.5 mol L⁻¹ NaCl, with addition of 0.02 mol L⁻¹ KCl and 0.02 mol L⁻¹ NH₄Cl)
632 and (case B) the Natural Saline Water (NSW, **Table 5**). The curves obtained using the CA
633 approach with the 1-pK DL/IE model for illite, montmorillonite and kaolinite (**Table 6**) are
634 superimposed to the experimental points for all conditions (**Figure 9**). A RMSD value of 0.10 L
635 kg⁻¹ was obtained (calculated on all the modeled data including those not shown here), vs. 0.19 L
636 kg⁻¹ for the model of Missana et al. (2014b).

637 The relative contribution of the different sorption sites to Cs sorption, for the experiment
638 with NSW is presented in **Figure A.12a** showing that Cs sorption is dominated by illite, over the
639 whole range of Cs loading. The contribution of kaolinite and montmorillonite FES can be
640 neglected (confirming the hypothesis of Missana et al., 2014b), and montmorillonite (exchange
641 sites) only contributes significantly at high Cs concentration.

642 The underestimation of the K_d value (1.53 L kg⁻¹) at high Cs equilibrium concentration (≈
643 [Cs]_{eq} = 2·10⁻² mol L⁻¹) after the saturation of exchange sites of illite and montmorillonite, can be
644 due to the existence of a low amount of another clay mineral which was not taken into account. In
645 **Figure A.12b**, the sorption curves are plotted for three different clay mineral fractions in SJ clay: a
646 minimum (30% illite, 5% montmorillonite and 10% kaolinite), medium (40% illite, 10%
647 montmorillonite and 12% kaolinite) and maximum value (50% illite, 20% montmorillonite and
648 15% kaolinite). These simulations show that the uncertainty on the clay mineral fraction results in
649 envelope curves bracketing the experimental error. The average clay minerals content proposed by
650 Missana et al. (2014b) results in the best fit (solid curve).

651

652 4.3.3 Callovo-Oxfordian samples (COx)

653

654 Cs sorption data for five samples of Callovo-Oxfordian claystone from Chen et al. (2014)
655 and Savoye et al. (2012) are considered here (**Table 6**). These samples contain pure and
656 interstratified illite and montmorillonite. The sorption isotherms were obtained in synthetic pore
657 waters (SPW) and in a reference pore water (RPW) (**Table 5**) containing the following cations:
658 Na^+ , Ca^{2+} , Mg^{2+} , and K^+ (and also Sr^{2+} in SPW-1).

659 The mineral fractions used in the calculations correspond to the lower and upper limit, and to
660 the mean values proposed by Chen et al. (2014) (**Table 9**). We considered also that the
661 contribution of the $\equiv\text{SO}^{-0.5}$ site exclusively from the pure illite (as confirmed by the previous
662 results). The experiment results are successfully simulated with 1-pK DL/IE model with the
663 different mineralogical compositions (**Figure A.13**): Cs sorption is dominated by illite over the
664 whole Cs concentration range and the contribution of montmorillonite (exchange sites) is
665 significant only at high Cs loading. Results obtained with sample EST27861-B show that when the
666 FES from the I/S illite fraction are considered (solid lines), Cs sorption increases by about 10%
667 (**Figure A.13B**). The EST27862 sample was modeled alternatively with interstratified illite and
668 montmorillonite only (dotted red line strongly underestimating measured data), and by adding
669 1.5% of illite (solid line) which greatly improved the fit (**Figure A.13D**). Note that the model of
670 Chen et al. (2014) underestimates Cs sorption at medium to high Cs concentration (**Figure A.13A**
671 **& D**).

672 Cs sorption on another COx sample (EST27337) was also modelled by Savoye et al. (2012).
673 The total clay fraction is 32–65% of clay minerals (Gaucher (1998) with 62% of illite/smectite
674 mixed-layer mineral (with 65% of illite, 25% of montmorillonite, and 10% of vermiculite),
675 smectite (9%), illite (26%), kaolinite (2%), and chlorite (1%) (Claret et al. (2004). We assumed

676 that this sample is composed by 65% of clay minerals and we used the mineralogy proposed
677 by Claret et al. (2004) (**Table 9**). **Figure A.14** shows Cs sorption isotherm onto EST27337 sample
678 in RPW (**Table 5**), with red lines representing the results the 1-pK DL/IE model using CA
679 approach considering clay minerals fraction given in **Tables 6**, and the green line those of the
680 model proposed by Savoye et al. (2012). These authors modeled Cs sorption using multi-site
681 cation exchange model considering four sorption sites for illite and three sites for montmorillonite
682 (for a total of 7 sites). The two models show the same goodness fit of the experimental curves with
683 $\text{RMSD} = 0.4 \text{ L kg}^{-1}$, but the 1-pK DL/IE model required much less sorption sites and fitting
684 parameters than the one used by Savoye et al. (2012).

685

686 4.3.4 Opalinus clay

687

688 The Opalinus clay (OPA) sample used in the calculations, referenced as SLA-938, contains
689 30% of illite/smectite mixed layer, 17% of pure illite and 21% of kaolinite, with a CEC value of
690 183 meq kg^{-1} (Marques Fernandes et al., 2015).

691 The modelling results are obtained considering that I/S mixed layers (30%) are exclusively
692 constituted by illite (for a total fraction of illite of 47%, as proposed by Marques Fernandes et al.
693 (2015) using the 1-pK DL/IE and the GCS model (green and red lines, respectively) (**Figure**
694 **A.15**). In addition, our model also integrates the kaolinite fraction. The Cs sorption data are very
695 well reproduced by both models ($\text{RMSD} = 0.23 \text{ L kg}^{-1}$). The results obtained with our model show
696 that the contribution of kaolinite is negligible.

697

698 4.3.5 Boom clay

699

700 Mineralogical studies of the Boom clay Formation show a variable clay mineral composition
701 with a total fraction in the interval between 30 and 60% and the main components as follows: illite
702 (10-45%), illite/smectite mixed layer (10-30%), kaolinite (5-20%) (Aertsens et al., 2004; De-Craen
703 et al., 2004).

704 The Cs adsorption isotherm on Boom clay in equilibrium with the Reference Boom Clay
705 Water (RBCW; **Table 5**) was first modelled using the 1-pK DL/IE and the GCS model considering
706 only illite (20, 30 and 45%), as tested by Maes et al. (2008). The best agreement with experimental
707 data is obtained using 30% illite with both models (RMSD = 0.32 L kg⁻¹ for the 1-pK DL/IE and
708 0.51 L kg⁻¹ for the GCS model) (**Figure 10**). With this simplified mineral composition, the GCS
709 model and the 1-pK DL/IE model both underestimates measured K_d values for Cs equilibrium
710 concentrations higher than 10⁻⁴ mol L⁻¹.

711 To obtain a better fit, we have added a montmorillonite and kaolinite fraction: 20% illite,
712 20% montmorillonite and 10% kaolinite (**Figure 10**, dotted red line). This change significantly
713 improved the sorption at high Cs concentrations and the RMSD value (0.11 L kg⁻¹) (Cs sorption
714 decomposition is shown in **Figure A.16**). Note that the contribution of kaolinite is less than 1% of
715 Cs total sorption.

716

717 **5 Conclusion**

718

719 We propose a new model for Cs sorption that combines a one-site 1-pK double layer model
720 and a one-site ion exchange model to account for varying levels of Cs concentrations (10⁻⁹-10⁻⁴
721 mol L⁻¹) and physicochemical conditions (pH, ionic strength, and competing ions). A single set of
722 parameters was adjusted for each clay mineral: illite, montmorillonite, and kaolinite. Since only 2
723 types of sorption site are considered and the same formalism is used for the three clay minerals,
724 this approach reduces the number of parameters compared to existing multi-site model for these

725 clay minerals. The application of this parsimonious model to simple clayey material, even with
726 complex natural pore-water compositions, shows the robustness of our model. The use of the
727 component additivity approach for complex mixed-clay materials, considering the clay mineral
728 fractions, was also successful without any adjustment of parameters.

729 The simulations performed with the 1-pK DL/IE model on claystones show that the
730 uncertainty concerning the clay mineral fraction can be taken into account and results in envelope
731 curves bracketing the experimental measurement error. To this regard, our model outperforms the
732 “generalized Cs sorption” (GCS) model (Bradbury and Baeyens, 2000) which is largely used in the
733 context of radioactive waste management. The GCS model is quite reliable with clay materials that
734 are comparatively rich in illite (e.g. Opalinus clay and Palfris marl) but underestimates Cs sorption
735 when the illite-montmorillonite mixed layer content is comparable or higher than illite (e.g. Boom
736 clay) (as also note by Chen et al. (2014). Our model, as well as the other models from the
737 literature, tends to underestimate the K_d values for very high Cs concentrations [10^{-4} to 10^{-3}
738 mol/L].

739 Another interesting result is that the effect of the electrostatic component in Cs sorption on
740 the clay minerals and naturel clayey materials can be neglected. In a next step, the 1-pK DL/IE
741 model will be used to simulate the behavior of Cs in natural soils.

742

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747

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