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The OpenCalphad thermodynamic software interface

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Applications using the OpenCalphad thermodynamic software interface including parallelization.

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Abstract

Thermodynamic data are needed for all kind of simulations of materials processes. For kinetic models, for example, thermodynamic quantities, such as chemical potentials, driving forces for precipitation, thermodynamic factors for converting mobilities to diffusion coefficients, solubilities in different phases etc., are needed. It is very important that these data are consistent and accurate, also for predicting metastable states during phase transformations. Equally important is that the computation time must be reasonably short.

Frequently thermodynamic data are pre-evaluated into “lookup tables” to speed up calculations. This creates additional uncertainties in the data as they must be interpolated or extrapolated and conditions may differ from those originally assumed for the creating the lookup table. It would be much better if a full local equilibrium calculation could be performed when

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5 data are needed, e.g. for each grid-point during a finite element simulation.
6
7 Efficiency requires that the thermodynamic software is fully parallelized. The
8
9 OpenCalphad (OC) software is the first such software to do so.

10
11 *Keywords:* Computational Thermodynamics, Phase Transformations,
12 Simulations, Parallel Computing,
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14 15 **1. Introduction**

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18 The background for the development of a free thermodynamic software
19 for multicomponent calculations using the CALPHAD method (CALculation
20 of PHase Diagrams) is described by Sundman et al. [1]. A very limited first
21 version of the OpenCalphad (OC) software was released in 2013. In February
22 2015 a second version capable also of multicomponent single equilibrium and
23 some phase diagram calculations was released. During 2015 extensive new
24 facilities were added and the software has become much more stable and both
25 version 2 and 3 are available on [<http://www.opencalphad.org>]. Version 3 in-
26 cludes parallel equilibrium calculations using the OpenMP [<http://www.openmp.org>]
27 library, a rudimentary assessment module and an Application Software Inter-
28 face, OCASI, callable from C++ as well as Fortran and other programming
29 languages. OC is written in the new Fortran standard.
30
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34 There is extensive literature describing the use of thermodynamic data
35 for materials and process development. Although somewhat outdated, an
36 excellent review can be found in the book by Miodownik and Saunders [3].
37 Especially the development of phase field techniques [4, 5, 6] and other tech-
38 niques, such as fluid dynamics [7] and finite element methods [8] for simulat-
39 ing microstructure evolution has increased the demand for fast and reliable
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5 calculations of chemical potentials, driving forces, mobilities etc. Within the
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7 framework of Integrated Computational Materials Engineering (ICME) [9]
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9 the necessary tools are coupled to accomplish the desired simulations and it
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11 is the intention that the OC software becomes one of these tools.

12
13 The OC software has been extensively tested on Windows and Linux sys-
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15 tems and although there are limitations and occasional problems the overall
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17 performance is comparable with commercial thermodynamics software for
18
19 materials science, such as FactSage [10], Pandat [11] or Thermo-Calc [12].
20
21 In addition, the OC software can perform many equilibrium calculations
22
23 in parallel, including calls to OCASI. The OCASI software is still under
24
25 development but has been used in some applications to demonstrate how
26
27 it can be integrated in simulations to utilize accurate and consistent ther-
28
29 modynamic data. Examples and full documentation of the source code as
30
31 well as macro files showing the interactive use of the software and the use
32
33 of OCASI are available on the OC web site [<http://www.opencalphad.org>].
34
35 A development version with some new features and bugfixes is available
36
37 at [<http://www.github.com>].

38 39 **2. Computational Thermodynamics**

40
41 Computational thermodynamics consists of software and databases for
42
43 calculating phase diagrams and other kinds of thermodynamic data. The
44
45 databases are a collection of parameterized model descriptions of the phases
46
47 as function of constitution and state variables. The model descriptions are
48
49 obtained from the thermodynamic assessment (or optimization) of individual
50
51 systems. The basic techniques are described in the book by Lukas et al. [15].
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5 An important advantage from using an assessed thermodynamic database is
6 that all data calculated with such a database are consistent, including phase
7 solubilities, melting temperatures, heat capacities, chemical potentials etc.
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9

10 11 *2.1. State variables*

12
13 A thermodynamic system is described by a set of state variables like
14 T, P, V for temperature, pressure and volume. The amounts of the compo-
15 nents can be given in moles, N , mole fractions, x , or mass fractions, w and
16 the equilibrium state is at a minimum of the appropriate state function like
17 internal energy, U , Helmholtz energy, A or the Gibbs energy, G .
18
19

20
21 Many properties can be derived from the state functions and some can
22 also be used to control the system, such as chemical potentials, μ , enthalpies,
23 H etc. All state variables and properties available in OC are listed in Ap-
24 pendix 2.
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30 31 *2.2. Thermodynamic models*

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33 In CALPHAD-type thermodynamic databases the Gibbs energy of each
34 phase is described by a model reflecting its structure. Several different kinds
35 of models are described by Lukas et al. [15]. The simplest model is that for
36 an ideal gas:
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39
40

$$41
42 G_M^{\text{gas}} = \sum_i y_i \text{}^\circ G_i + RT y_i \ln(y_i) + RT \ln(P) \quad (1)
43
44
45$$

46 where $\text{}^\circ G_i$ is the Gibbs energy of formation of the gas molecule i from the
47 reference state of the elements, y_i is the constituent fraction of the molecule in
48 the gas and P the pressure. R is the gas constant, T the absolute temperature
49 and $y_i \ln(y_i)$ is the random configurational entropy.
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5 The Gibbs energy model is per mole formula unit of the gas as the number
6 of atoms will vary with the species. The mole fractions of the elements, x_j ,
7 are given by
8
9

$$10 \quad x_j = \frac{\sum_i b_{ji} y_i}{\sum_k \sum_i b_{ki} y_j} \quad (2)$$

11
12
13
14 where j and k denote elements and i the molecules. b_{ji} is the stoichiometric
15 factor of element j in molecule i .
16
17

18 More complex phase models with sublattices for interstitials, chemical
19 ordering or defects, including ionic constituents are implemented in OC. The
20 interactions between the constituents on a sublattices can be described by
21 excess parameters and a model to describe the contribution of magnetic or-
22 dering depending on the Curie or Néel temperature and the Bohr magneton
23 number to the heat capacity is available. All these models are special cases of
24 Compound Energy Formalism (CEF) [16] and are implemented in OC as well
25 as the partially ionic 2-sublattice liquid model [17] for short range ordering
26 (SRO) in the liquid phase. More models will be implemented with time and
27 since the source code of OC is available skilled users can implement their
28 own models.
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40 *2.3. Equilibrium calculations*

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42 The equilibrium algorithm used in OC was proposed by Hillert [18] and
43 implemented by Jansson [19] in the POLY module of Thermo-Calc software
44 and in the PMLFKT software by Lukas [20]. The implementation of this
45 algorithm in OC is described in [21]. In this context it is sufficient to mention
46 that the data structure, described in [21], is designed to allow calculation of
47 several equilibria in parallel and it has an additional grid minimizer to ensure
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5 that the calculated equilibrium is a global and not a local one. It can also
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7 detect miscibility gaps.

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9 A special feature of Hillerts algorithm [18] is the possibility to prescribe
10 that a phase should be stable. This feature is essential for the ability to
11 calculate phase diagrams and the direct calculation of a transition, such as
12 the melting temperature of an alloy.
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17 *2.4. Parallel calculations*

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19 Calculation of multiple equilibria in parallel is not only useful for simula-
20 tions but also for assessments of model parameters. Such assessments usually
21 include a large number of experimental data, each representing a value deter-
22 mined at equilibrium. The assessment procedure calculates these equilibria
23 many times varying the selected model parameters in order to find the best
24 fit. Traditionally the equilibria are calculated sequentially but with the OC
25 software it is possible to calculate them in parallel.
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33 When tested on a PC with 4 CPU, each with 2 threads, the gain in speed
34 for a calculation with 400 experimental equilibria in parallel was almost a
35 factor of 4, i.e. no significant overhead.
36
37

38 The occurrence of a miscibility gap in a phase requires special treatment
39 for parallel processing. A miscibility gap means that the phase is stable with
40 two different compositions which requires that an extra composition set must
41 be created in OC, see section 3.7. This must be done at the same time in
42 all equilibria otherwise the data structure would become inconsistent. Thus
43 the threads must be stopped while the extra composition set is created.
44 Normally this is not a problem during an assessments as miscibility gaps
45 should be known from the experimental data and the user can manually
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5 create a sufficient number of composition sets before starting the parallel
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7 session. However, miscibility gaps or order/disorder transformations can
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9 create problems when encountered during simulations.

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11 *2.5. Additional data depending on T, P and phase composition*

12
13 Thermodynamic data are important but not sufficient to simulate a phase
14
15 transformation. For example, to consider the diffusion of elements their
16
17 mobilities together with the thermodynamic factor need to be described to
18
19 obtain the diffusion coefficient. In addition, there are also interface energies,
20
21 interface mobilities, elastic constants and other phase based properties that
22
23 may be of interest [22].

24
25 In the OC software such property descriptions can be stored together with
26
27 the thermodynamic data for each phase as functions of T, P and constitution
28
29 of the phase. The values of these properties are calculated together with the
30
31 thermodynamic properties during the equilibrium calculation and can be
32
33 accessed by the application software.

34
35 For example, the natural logarithm of the mobility of Fe in the face
36
37 centered cubic (FCC) phase of the binary Fe-Ni system, in OC denoted
38
39 $MQ\&Fe(FCC)$, can be described by a linear function:

40
41
$$MQ\&Fe(FCC) = x_{Fe}^{FCC} MQ\&Fe(FCC, Fe) + x_{Ni}^{FCC} MQ\&Fe(FCC, Ni) \quad (3)$$

42
43 where x_{Fe}^{FCC} and x_{Ni}^{FCC} are the mole fractions of Fe and Ni in FCC, re-
44
45 spectively, $MQ\&Fe(FCC, Fe)$ is the mobility of Fe in pure FCC Fe and
46
47 $MQ\&Fe(FCC, Ni)$ is the mobility of a single Fe atom in pure FCC Ni, both
48
49 of which may depend on T as:

50
51
$$MQ\&Fe(FCC, Ni) = \frac{Q_{Fe}}{RT} \ln(\nu_{Fe}) \quad (4)$$

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5 where Q_{Fe} is the activation energy and ν_{Fe} the pre-exponential factor. As
6 mobilities can vary several orders of magnitude it is a better approximation
7 to describe the composition dependence by a linear relation of its logarithm.
8 If sufficient experimental or theoretical data are available, non-linear terms
9 can be added to eq. 3.
10

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12 Together with the thermodynamic factor the mobility can be transformed
13 to a diffusion coefficient needed for the simulation of microstructure evolution
14 in diffusion controlled processes.
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21 **3. The OCASI application software interface**

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24 Software interfaces for coupling thermodynamic calculations with appli-
25 cation software following a proposed standard [23] implemented in slightly
26 different ways in ChemApp [24] and Thermo-Calc TQ [25]. The idea behind
27 this proposed standard is to overcome the differences between the different
28 thermodynamic software packages which can make the implementation into
29 application software difficult. The methodology with many separate sub-
30 routines and functions performing specific tasks makes the implementation
31 straightforward and used also for OCASI.
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39 A full documentation of OCASI is under preparation but a preliminary
40 version is available at the websites mentioned above. Again, as the source
41 code is available any user can directly access the source code and also data
42 which are not primarily available. However, any such access should be done
43 with caution as the data structure and source code may change in a later
44 update whereas the OCASI interface should remain stable. A summary of
45 the available subroutines is given in Appendix 1.
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3.1. Fortran and C++ data structure compatibility

Using the ISO C standard it is possible to access data structures defined in the OC software written in Fortran from C++ and other languages. For example, the data for a composition set of a phase in OC are described by:

```
TYPE gtp_phase_varres
! Data here are for a composition set of a phase in an equilibrium
! There can be several equilibria for assessments or parallel calculations
! phlink: is index of phase record for this phase_varres record
! status2: has phase status bits like
! phstate: indicate state: fix/stable/entered/unknown/dormant/suspended
! phtupx: phase tuple index
    integer nextfree,phlink,status2,phstate,phtupx
! abnorm(1): amount moles of atoms for a formula unit of the composition set
! abnorm(2): mass/formula unit (both set by call to set_constitution)
! prefix and suffix are added to the name for composition sets 2 and higher
    double precision, dimension(2) :: abnorm
    character*4 prefix,suffix
! constat: array with status word for each constituent
! yfr: the constituent fraction array
! mmyfr: min/max fractions, negative is a minimum
! sites: site ratios (which can vary for ionic liquids)
    integer, dimension(:), allocatable :: constat
    double precision, dimension(:), allocatable :: yfr
    real, dimension(:), allocatable :: mmyfr
    double precision, dimension(:), allocatable :: sites
! for ionic liquid derivatives of sites wrt fractions (it is the charge),
    double precision, dimension(:), allocatable :: dpqdy
    double precision, dimension(:), allocatable :: d2pqdvay
! disfra: a structure describing the disordered fraction set (if any)
    type(gtp_fraction_set) :: disfra
! arrays for storing calculated results for each phase (composition set)
! amfu: is amount formula units of the composition set (calculated result)
! netcharge: is net charge of phase (must be zero for stable phases)
! dgm: driving force
    double precision amfu,netcharge,dgm
! nprop: the number of different properties (set in allocate)
! listprop(1): is number of calculated properties
! listprop(2:listprop(1)): identifies the property stored in gval(1,ipy) etc
! 1=G, 2=TC, 3=EMAG. Properties defined in the gtp_propid record
    integer nprop
    integer, dimension(:), allocatable :: listprop
! gval etc are for all composition dependent properties, gval(*,1) for G
! gval(*,1): is G, G.T, G.P, G.T.T, G.T.P and G.P.P
! dgval(1,j,1): is first derivatives of G wrt fractions j
! dgval(2,j,1): is second derivatives of G wrt fractions j and T
! dgval(3,j,1): is second derivatives of G wrt fractions j and P
! d2gval(ixsym(i,j),1): is second derivatives of G wrt fractions i and j
```

```

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5     double precision, dimension(:,:), allocatable :: gval
6     double precision, dimension(:,:,:), allocatable :: dgval
7     double precision, dimension(:,:), allocatable :: d2gval
8 ! added for strain/stress, current values of lattice parameters
9     double precision, dimension(3,3) :: curlat
10    END TYPE gtp_phase_varres
11 ! an array of these records is created inside each equilibrium record
12

```

This can be summarized for C++ software by the ISO C interface as:

```

15 typedef struct {
16     int nextfree, phlink, status2, phstate;
17     double abnorm[2];
18     char prefix[4], suffix[4];
19     int *constat;
20     double *yfr;
21     double *mmyfr;
22     double *sites;
23     double *dpqdy;
24     double *d2pqdvay;
25     //struct gtp_fraction_set disfra;
26     double amfu, netcharge, dgm, amcom, damount;
27     int nprop, ncc;
28     int *listprop;
29     double **gval;
30     double ***dgval;
31     double **d2gval;
32     double curlat[3][3];
33 } gtp_phase_varres;

```

3.2. Interfacing with Java

Based on Fortran and C(++) compatibility, it is also possible to interface OC functionalities in the Java language thanks to the Java Native Interface (JNI) [26]. For the PROCOR software [27], such an interfacing has been developed for C. Cardon Ph.D. thesis regarding the development of a Cahn-Hilliard based multicomponent multiphase diffusion model [28]. This application is briefly described in section 4.4. In addition to the Cahn-Hilliard model requirements, equilibrium calculation functionalities have been interfaced for the coupling of other physical models to evaluate some physical parameters using the phase compositions obtained from OC equilibrium calculations.

3.3. Reading the database

The thermodynamic data for a system will normally be read from a database. The OC software supports the TDB format [29] proposed originally by SGTE [30] and is used by several other software, most notably by Thermo-Calc and DICTRA [12]. The application software allows to select the elements for which data that should be extracted from the database and then to suspend phases that are not of interest for the application.

The application software must initiate the OC memory structure by a special subroutine call. This returns a pointer to an equilibrium record in the dynamic memory structure which contains information about the constitution and calculated results for all the phases as described in section 3.1. Only after this initialization other subroutines, such as reading the database, can be called. The database has data like the models of phases and their thermodynamic model parameters and these data are stored in the static part of the OC memory structure since they are independent of the external conditions. The external conditions, the constitution of the phases and the calculated results are in the dynamic data memory. A schematic representation of both static and dynamic memory is shown in Fig. 1.

3.4. Setting conditions

The conditions are set individually by a call to a subroutine. In the simplest case these represent the temperature, T , pressure, P , and amount of the components, N_i . The number of conditions needed to calculate the equilibrium is given by Gibbs phase rule

$$f = n + 2 - p \quad (5)$$

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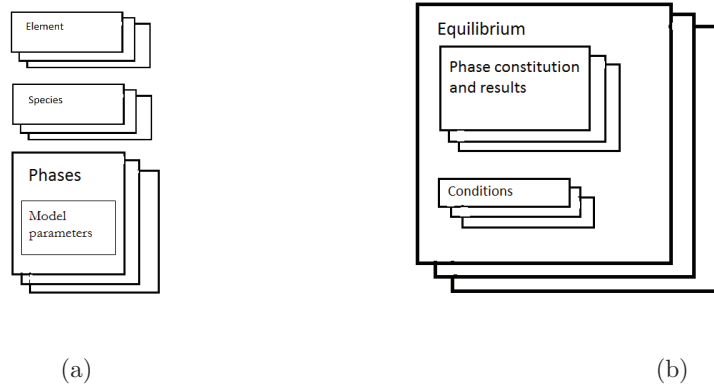


Figure 1: The static (a) and dynamic (b) parts of the OC memory. Each equilibrium record in (b) is independent and can be calculated in parallel.

where f is the degrees of freedom, n the number of components and p is the number of stable phases. Normally the number of stable phases is the result of the equilibrium calculation but OC allows that one or more phases are prescribed to be stable. If there are no prescribed stable phases, a system with a single component, $n = 1$, must have 3 conditions, for example T, P and N where N is the size of the system. For a binary system 4 conditions must be set and so on.

The amount of a component i can be given in moles, N_i , or mass and if there is a major component they are often given in mole or mass fractions or per cent and the total size of the system is given as N . Chemical potentials or activities of the components can also be used instead of amounts.. It is also possible to specify the volume or the enthalpy of the system, or as already mentioned, specify that a phase has to be stable. It is important that all conditions are independent of one another. However, this does not mean that it is not possible to set irrational conditions, i.e. those that do

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5 not represent a possible equilibrium state. A full list of state variables in OC
6 is given in Appendix 2.
7
8

9 10 *3.5. Creating additional equilibria*

11 Additional equilibrium records can be created to represent local equilibria
12 in a simulation or assessment. These may be calculated sequentially or in
13 parallel as they have independent sets of conditions and related data. Each
14 equilibrium record has a pointer to its own dynamic data structure and this
15 pointer must be supplied in the calls to most of the OCASI routines to
16 indicate which equilibrium should be used.
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24 *3.6. Calculations*

25 When the degree of freedom according to eq. 5 is zero the equilibrium
26 can be calculated by calling a subroutine. The calculation procedure in OC
27 is described in detail in [21]. In the call it is possible to specify whether the
28 global minimizer should be used or the calculation should start from the last
29 calculated set of stable phases and constitutions.
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37 *3.7. Composition sets and phase tuples*

38 The number of phases in the system when reading from the database
39 is equal to the number of phases with unique names and structures in the
40 database. However, as phases can have miscibility gaps or several types of
41 ordering transformations some of the phases may become stable with two
42 or more unique compositions. In some software the user must specify this
43 in advance but in OC the global grid-minimizer will detect miscibility gaps
44 automatically. A phase may thus have several independent composition sets
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5 and each of these is called a phase tuple. The composition set index of a
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7 phase is specified after the phase name by the character “#” followed by a
8
9 digit, for example LIQUID#2.

10
11 In some cases, for example phases with order/disorder transitions, it is
12
13 important to know the state of ordering of a particular phase tuple. For
14
15 such cases composition sets can be created explicitly and they can be given
16
17 a unique name and default composition that can be used to identify the
18
19 ordered state.

20
21 The first composition set created when the phase is read from the database
22
23 has no explicit set indication. A composition set created automatically by
24
25 the grid-minimizer will have a suffix “_AUTO” added to its name (before the
26
27 # sign).

28
29 In OCASI there is an array of phase tuples where each index represents
30
31 the combination of a phase and composition set. Thus the number of phase
32
33 tuples may change whenever the grid-minimizer is called.

34 35 *3.8. Extracting results*

36
37 The normal way to obtain calculated results is to call a subroutine using
38
39 state variable symbols from the list in Appendix 2. The values of properties
40
41 that are modeled separately, like the Curie temperature, mobilities etc. and
42
43 that have been read from the database can be obtained in the same way.

44
45 Using the pointer to the equilibrium data structure provided when read-
46
47 ing the database it is possible to extract individual calculated results of an
48
49 equilibrium. For example, if the equilibrium record pointer is called “ceq”
50
51 and as the array of records for each phase is called “phase_varres”, the value
52
53 of the Gibbs energy per mole formula unit of the phase after an equilibrium
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5 calculation can be obtained as
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7 `gfu = ceq%phase_varres(lokcs)%gval(1,1)`
8
9

10 where the % sign is used in the new Fortran standard to indicate items
11 in a record. The lokcs variable is the index of the composition set of the
12 phase and this is stored in the phase tuple record and thus provides an easy
13 link between the phase tuple index and the phase data. For example, if the
14 calculated mobility of component 1 of the same phase is stored in the location
15 “ipy”, see section 3.1, its value is obtained by
16
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22 `mob(1) = ceq%phase_varres(lokcs)%gval(1,ipy)`
23
24

25 Great care must be taken that the relevant data is extracted. The application
26 program must not change the OC data structure.
27
28

29 30 **4. Examples of application software** 31 32

33 Several simple examples of how to use the OC software interface are
34 provided with the source code. Some of the examples presented here may
35 require the use of proprietary software components.
36
37
38
39

40 *4.1. The beginnings of a reactor module* 41

42 A process can sometimes be simulated by dividing it into sections, each
43 of which represent a local equilibrium, and a transport of matter between
44 these by diffusion or other means. The equilibrium calculation can be done
45 by OC in a function “localequil” with the following arguments:
46
47
48
49

```
50 integer function localequil(todo,nse1,sele1,tp,nmz,outspec,outval)  
51 !  
52 ! This routine calculates an equilibrium and return values  
53
```



```

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4
5      ! todo character variable          with "commands" to be executed
6      ! nsel integer                    number of components
7      ! selel character*2 array         element names
8      ! tp double precision array(2)   with values of T and P
9      ! nmz double precision array(nsel) with molar content for elements
10     ! outspec character variable      with information which data to extract
11     ! outval double precision array(*) with extracted values
12     ! if there is no error the function value returned is zero

```

13 The `todo` argument passes a number of instructions to `localequil`, `nsel`
14 gives the number of elements specified in argument `selel`, `tp` provides the
15 values of T and P and `nmz` the element amounts. The argument `outspec`
16 specifies which state variable values will be returned in `outval`. A call of this
17 function to calculate a local equilibrium for the system Si-O-C at 1800 K and
18 1 bar with a specified amount of moles of the elements can be:

```

25     nsel=3
26     selel(1)='C '; selel(2)='O '; selel(3)='Si'
27     tp(1)=1.80D2; tp(2)=1.0D5;
28     nmz(si)=1.0D0; nmz(2)=2.0D0; nmz(3)=1.8D0
29     todo='verbose open cosi.tdb '
30     outspec='n(gas,*) mu(*) h h(gas) '
31     ierr=localequil(todo,nsel,selel,tp,nmz,outspec,outval)

```

32 Here the `todo` argument specifies `VERBOSE` to generate extra output,
33 then to open the database file `cosi.tdb` and read the data for the system
34 with the elements in the argument `selel` and calculate the equilibrium at
35 the specified values of T , P and amounts of the elements. After a successful
36 calculation the amounts of all elements in the gas, $n(\text{gas},*)$, their chemical
37 potentials, $\mu(*)$, the total enthalpy, h , and the enthalpy of the gas, $h(\text{gas})$,
38 will be provided in the `outval` array argument. The state variables that can
39 be used are listed in Appendix 2. If there is an error the function will return
40 a non-zero number.

41 The `localequil` function will remember whether it has already read a
42 database and if no database is specified in a subsequent call the routine will

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5 use the static data that it has already stored. In the application software the
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7 amounts of elements in the gas and the other phases may be redistributed
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9 between different local equilibria to simulate a process. The range of options
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11 in the localequil function can easily be extended. One of the intentions of
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13 this function is to provide an application programmer a simple access of the
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15 OCASI interface.
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17 *4.2. Simulation of homogenization of an alloy AA7449*

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19 The OpenCalphad Software has been recently coupled to the Constellium
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21 thermodynamic package, ProPhase, using OCASI in order to use the faster
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23 equilibrium routine included in OC as well as the new parallelization recently
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25 implemented. ProPhase includes a long range diffusion subroutine which is
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27 used when simulating a solidification path with back diffusion or a homog-
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29 enizing treatment of an as cast structure. The long-range diffusion within
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31 the FCC aluminum solid solution is simulated using an explicit finite differ-
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33 ence scheme, assuming local equilibrium in each volume element, calculating
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35 the solute concentration in solid solution, and using Fick's diffusion equation
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37 in the solid solution to compute solute exchange between adjacent volume
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39 elements. For each time step, this routine performs many independent equi-
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41 librium calculations, one for each volume element, and is very well suited
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43 for parallelization. The simulation of a laboratory heat treatment at 472 °C
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45 of a multicomponent aerospace 7449 alloy (an Al-Mg-Zn-Cu-Fe-Si alloy, see
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47 Table 1) after casting has been used to measure the gain in speed provided
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49 by OpenCalphad in comparison with the previous software. The cast was
50
51 performed in the laboratory using the Aluminum Association TP-1 grain
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53 refiner test crucible. The corresponding cast structure, as shown in Fig. 2
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was simulated by a modified Scheil calculation with back diffusion. For this particular case, a two dimensional cylindrical geometry was chosen with a 61 μm secondary dendrite arm spacing. Care must be taken to mesh the dendritic cylinder with sufficiently small enough radius increments especially towards the end of solidification where the concentration gradients may be very steep.

Table 1: Composition range of AA7449 according to the Aluminum Association

	Si	Fe	Cu	Mn	Mg	Zn	Ti+Zr	rest
min	0	0	1.4	0	1.8	7.5	0	Al
max	0.12	0.15	2.1	0.2	2.7	8.7	0.25	Al

The evolution of composition gradients has been studied for a heat treatment consisting in a linear heating of 40 $^{\circ}\text{C}/\text{h}$ up to 472 $^{\circ}\text{C}$ followed by holding at 472 $^{\circ}\text{C}$ for 10 h. The corresponding predicted and measured incipient melting temperatures and integrated enthalpies for the DSC first melting peak are compared in Fig. 3. Overall, good agreement is obtained indicating that the proposed physical description of the phenomena occurring during solidification and homogenizing is acceptable. With our previous software such a calculation took 3.5 days on one processor. By using the parallel OpenCalphad routine the total computation time was reduced to 3.5 hours on 12 processors.

4.3. Phase-field models and OCASI

Many modern thermodynamic consistent phase-field models [31, 32, 33, 34] rely on thermodynamic and kinetic properties that can be obtained or ap-

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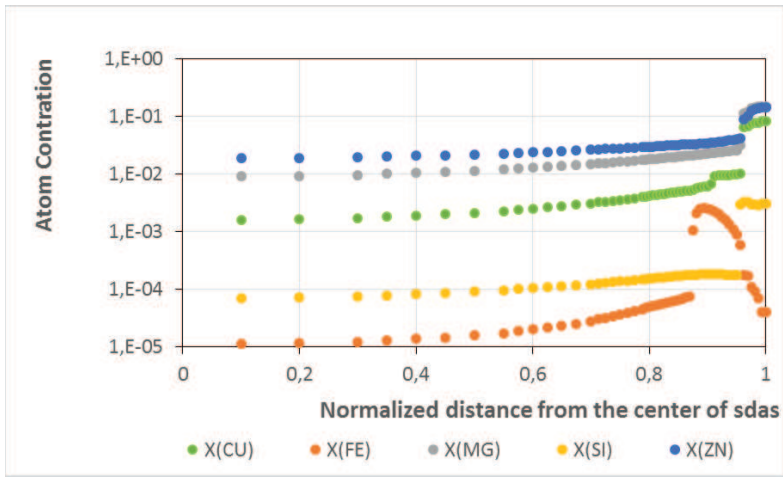


Figure 2: Simulated concentration gradients after solidification of aerospace alloy AA7449

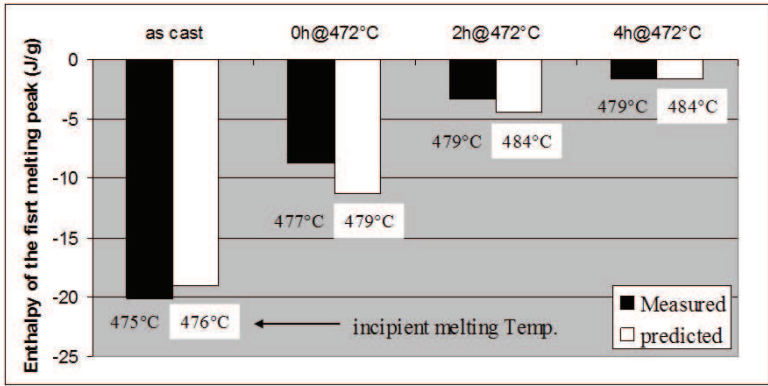


Figure 3: Simulated and measured incipient melting temperatures and enthalpies of the first melting peak during calorimetric heat up at a rate of 20 °C/min.

proximated from CALPHAD calculations. OpenCalphad [<http://www.opencalphad.org>] as an open source tool is perfectly suited for this purpose. The speed of the OCASI software with additional parallelization allows for an efficient way to compute separate phase-properties in the diffuse interfaces with equilibrium calculations or with separate phase-compositions as in-

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5 put parameters. Coupled to the phase-field software framework of Open-
6 Phase [<http://www.openphase.de>], OCASI has been proven to be an efficient
7 tool for meso-scale microstructure simulations. As extensive phase data can
8 be shared via this interface too, a user friendly simulation setup can be ar-
9 ranged.
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14 To be able to compute the equations governing phase-transformation and
15 diffusion in the phase-field model with finite interface dissipation [34], Gibbs
16 energies of the phases, G_m^α , their derivatives $\partial G_m^\alpha / \partial y_i^{\alpha,(s)}$ and the atomic
17 mobilities M_i^α can be calculated with an input consisting of the temperature
18 T , pressure P and compositions $y_i^{\alpha,(s)}$ through the OCASI interface.
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24 This does not only speed up the simulations and increases the accuracy
25 of the results, it also allows for a quick simulation setup of multi-component
26 multi-phase simulations, as no analytical implementation of the Gibbs en-
27 ergies or parallelization of phase equilibria calculations, has to be carried
28 out.
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33 This is demonstrated here for a small microstructure simulation example
34 with two austenitic steel samples with different alloy composition in contact.
35 The two samples containing iron, manganese, silicon and carbon are chosen to
36 represent the experiments on uphill diffusion by Darken [36]. Apart from the
37 chosen thermodynamic and kinetic databases [37] under the Open Database
38 License, only the temperature, the time step and the grid spacing had to be
39 specified as in Tab. 2
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47 The results of the quaternary simulation are shown in Fig. 4 for an exper-
48 iment with a duration of 10 days at constant 1323 K and are compared with
49 the results of a DICTRA sharp-interface simulation as well as measurements
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Table 2: Simulation parameters

Box size	50x10x10
Grid spacing	0.001 m
Time step	1000 s
Temperature T	1323 K
Thermodynamic database	MatCalc mc_fe_v2.057.tdb [37]
Mobility database	MatCalc mc_fe_v2.008.ddb [37]

from Darke, [36]. The results obtained with only openly available tools and data show excellent agreement with the results from the commercial DIC-TRA software.

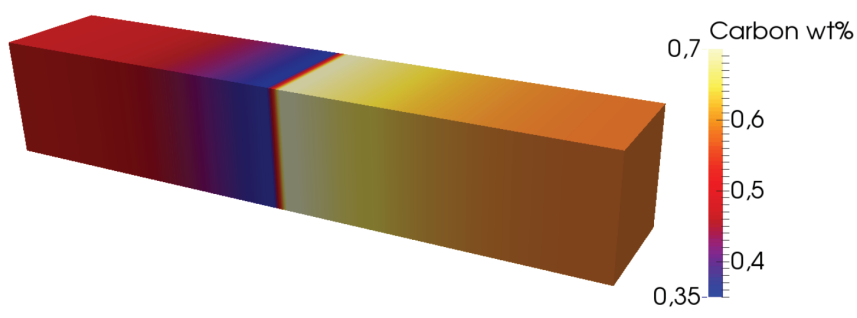
Beyond that, the coupling of OpenCalphad and phase-field software like OpenPhase has been proven to be an efficient, accurate and easy to use tool for meso-scopic microstructure simulations.

4.4. Mesoscopic modelling of a liquid miscibility gap

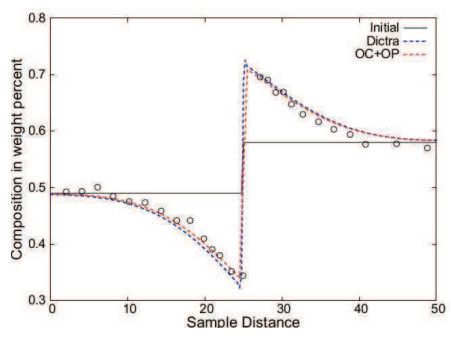
For the study of severe accidents in light water nuclear reactors (LWR), the OpenCalphad software has been used in the development of a Cahn-Hilliard based diffusion model for treating the liquid-liquid miscibility gap associated with the corium (oxidic and metallic liquid materials present during a reactor core meltdown) in a reactor vessel lower head.

In-vessel corium is a complex thermodynamic system because of the phases that are present. In addition to partially oxidized cladding and fuel materials (UO₂, ZrO₂ and Zr), the corium contains stainless steel elements

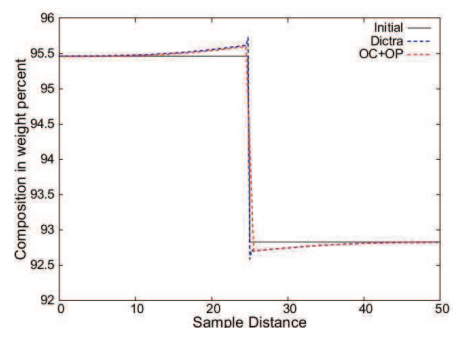
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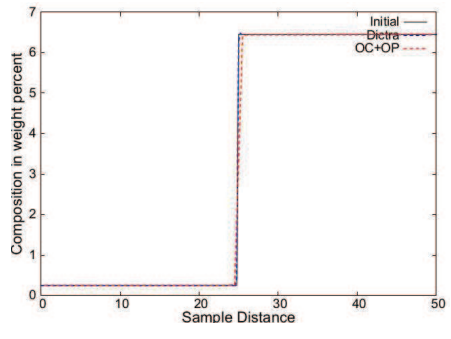
(a) Carbon composition over the whole simulation sample



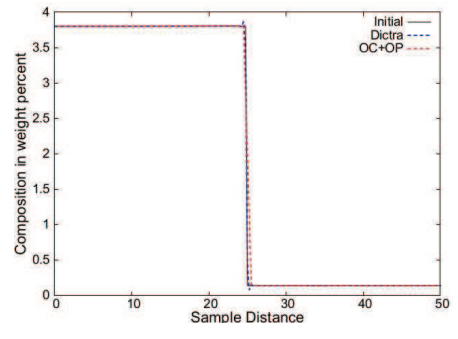
(b) Carbon composition



(c) Iron composition



(d) Manganese composition



(e) Silicon composition

Figure 4: Simulated composition profiles after 10 days of tempering at 1050 °C, with experimental measurements by Darken [36] as circles.

(Fe, Ni, Cr) from the reactor’s internal structures. In particular, the U-O-Zr-Fe system exhibits a liquid-liquid miscibility gap responsible for a liquid phases stratification of tremendous importance on the corium propagation [38].

This stratification phenomenon results from the combination of multi-component multiphase species diffusion and gravitational material movement due to buoyancy effects initiated by Rayleigh-Taylor instabilities. The natural convection in a corium pool can be simulated by Computational Fluid Dynamics (CFD) and a Cahn-Hilliard based mesoscopic modeling of the species diffusion appears to be a promising way to account for the stratification phenomenon in detailed simulations. This R&D effort regarding a detailed modeling of the thermohydraulic/thermochemical behavior of in-vessel corium has been undertaken at CEA Cadarache in support of integral models development in the context of the PROCOR software platform for sensitivity/uncertainty studies regarding corium propagation.

A Cahn-Hilliard based diffusion model was first developed for the U-O binary case [28] and is under development for U-O-Zr and U-O-Zr-Fe. Under the assumption of constant molar volume, the model for a n-component system is based on Cahn-Hilliard evolution equations for molar fractions of n-1 components with diffusion potentials, $\tilde{\mu}_i$, expressed as:

$$\tilde{\mu}_i = a \left(\frac{\partial \tilde{g}}{\partial x_i} \right)_{x_j \neq i} - \sum_j \kappa_{i,j} \nabla^2 x_j \quad (6)$$

where a and $\kappa_{i,j}$ are parameters of the model related to the interface thickness (a “numerical” parameter in this mesoscopic modeling) and interface tension while \tilde{g} is the free energy density (in Jm^{-3}) of the homogeneous system.

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5 In the binary U-O case, the $\left(\frac{d\bar{g}}{dx_O}\right)$ term was calculated applying the chain
6 rule to the liquid Gibbs energy G_m^{liq} and its derivatives $\partial G_m^{liq} / \partial y_i^{liq}$ obtained
7 from the OCASI interface [28].
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10 For the ternary and quaternary systems, the additional hypothesis that
11 local equilibrium of the redox chemical reactions is instantaneously reached is
12 made. Therefore, $\left(\frac{\partial \bar{g}}{\partial x_i}\right)_{x_j \neq i}$ can be related to the chemical potentials obtained
13 from “local” equilibrium calculations performed using the OCASI interface
14 where the phase separation associated with the miscibility gap is not taken
15 into account (*i.e.* the grid minimizer is turned off, *cf.* Section 2.3).
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22 In order to illustrate this modeling, a 1D calculation for the U-O binary
23 case is presented here; the domain length is $L = 10.0$ cm and the temperature
24 and pressure conditions are $T = 3200$ K and $p = 1$ bar. Fig. 5 schematically
25 depicts the initial configuration and the steady-state of the system. The
26 interface thickness has been set to 1 cm.
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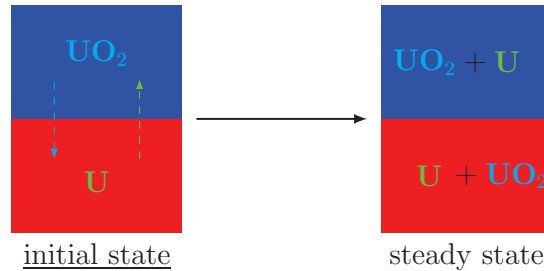
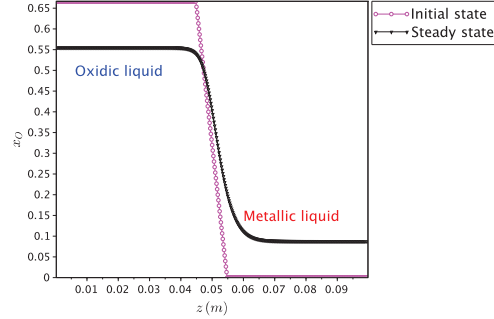


Figure 5: U-O system initial configuration and steady state

The initial configuration corresponds to an oxidic phase (uranium dioxide UO_2) above a metallic phase (in this case metallic uranium U). The diffusion process leads to a steady state configuration with two suboxidized phases.

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5 Fig. 6 shows the profile of oxygen molar fraction x_O in the initial state and
6 at steady state where it is verified that, far from the interface, the oxygen
7 molar fraction is equal to the values given by an equilibrium calculation.
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25 Figure 6: Profile of molar oxygen fraction x_O versus position z
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28 5. Summary

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31 Thermodynamic data needed in simulations, such as melting temper-
32 atures, solubilities, heat capacities etc., are often collected from different
33 sources. In most cases little or no efforts are made to verify that these data
34 are consistent. Data from calculations using thermodynamic databases have
35 the advantage of being consistent but the fact that the equilibrium calcu-
36 lations are usually quite slow poses a severe problem. The possibility to
37 calculate local equilibria at each point in space in parallel is a major im-
38 provement. Still, it may not be possible to completely avoid the use of inter-
39 polation methods in very large simulations with several thousand grid-points.
40 In this case the availability of parallel processing offers the opportunity to
41 update the results in real time during the simulation avoiding the need for
42 look-up tables created in advance thus allowing more accurate input for the
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5 simulation. The examples selected are from very different fields showing the
6 flexibility of the OC software.
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8 9 10 **6. Acknowledgement**

11
12 One author (BS) is grateful for a senior research grant from the Humboldt
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15 in Section 4.4 has been carried out within the framework of the PROCOR
16 platform development funded by CEA, EDF and AREVA. Commercial prod-
17 ucts are identified for reference purposes this does not imply an endorsement
18 by the National Institute of Standards and Technology.
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Appendix 1

Summary of current subroutines included in OCASI.

tqcceq	Create and copy equilibrium
tqce	Calculate equilibrium
tqcph1	Calculate phase tuple properties
tqdceq	Delete equilibrium
tqgcef	Get stoichiometry of component
tqgcom	Get the component names
tqgetv	Get state variable value
tqgnp	Get number of phase tuples
tqgnpc	Get number of constituents in phase
tqgpci	Get index of phase constituent using name
tqgpcn2	Get phase constituent index using name
tqgpcs	Get stoichiometry of constituent
tqgphc1	Get phase constitution
tqgpi	Get phase tuple index using phase name
tqgpn	Get phase name using phase tuple index
tqini	Initiate OC and the OCASI interface
tqphtupsts	Set phase tuple status
tqrpfil	Read data for selected elements from a database
tqselceq	Select equilibrium
tqsetc	Set condition
tqsphc1	Set constitution of phase tuple

Appendix 2

The available state variables and their notation in OC. All of them can be used to extract calculated values and many also to set conditions.

In this notation N is the total amount of moles, $N(O)$ is the total amount of oxygen in moles, $N(\text{GAS},O)$ is the amount of moles of O in the gas phase. $NM(O)$ or $X(O)$ is the overall mole fraction of O. The total number of moles of components in the gas phase is $NP(\text{GAS})$.

The heat capacity is not included in this list since it can be calculated as a “dot derivative, H.T” , as explained in [21].

Symbol	Index		Normalizing	Meaning
	1	2	suffix (z)	
Intensive properties				
T	-	-	-	Temperature
P	-	-	-	Pressure
MU	component	-/phase	-	Chemical potential
AC	component	-/phase	-	Activity
LNAC	component	-/phase	-	LN(activity)
Extensive and normalized properties				
U	-/phase#set	-	-	Internal energy for system
UM	-/phase#set	-	M	Internal energy per mole
UW	-/phase#set	-	W	Internal energy per mass
UV	-/phase#set	-	V	Internal energy per m ³
UF	phase#set	-	F	Internal energy/mole formula unit
Sz	-/phase#set	-	-/M/W/V/F	Entropy
Vz	-/phase#set	-	-/M/W/V/F	Volume
Hz	-/phase#set	-	-/M/W/V/F	Enthalpy
Az	-/phase#set	-	-/M/W/V/F	Helmholtz energy
Gz	-/phase#set	-	-/M/W/V/F	Gibbs energy
NPz	phase#set	-	-/M/W/V/F	Moles of phase
BPz	phase#set	-	-/M/W/V/F	Mass of phase
Qz	phase#set	-	-/M/W/V/F	Stability of phase
DGz	phase#set	-	-/M/W/V/F	Driving force of phase
Nz	-/phase#set/comp	-/comp	-/M/W/V/F	Moles of component
X	phase#set/comp	-/comp	-%	Mole fraction/per cent
Bz	-/phase#set/comp	-/comp	-/M/W/V/F	Mass of component
W	phase#set/comp	-/comp	-%	Mass fraction/per cent
Y	phase#set	const#subl	-	Constituent fraction
Some model parameter identifiers				
TC	phase#set	-	-	Curie temperature
BMAG	phase#set	-	-	Aver. Bohr magneton number
MQ&X	phase#set	constituent	-	Mobility of X
THET	phase#set	-	-	Debye temperature