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Guanine Quadruplexes

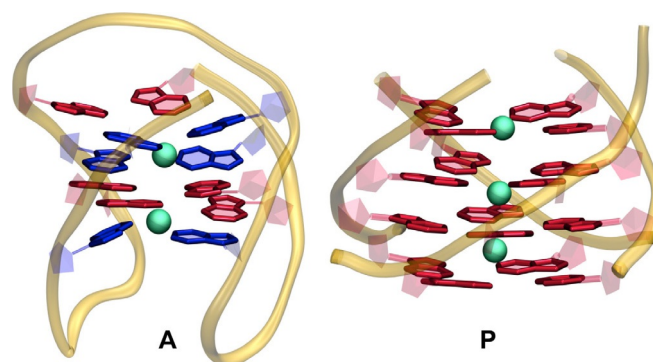
Topology Controls the Electronic Absorption and Delocalization of Electron Holes in Guanine Quadruplexes

Lara Martínez-Fernández,^[a] Akos Banyasz,^[a] Dimitra Markovitsi,^{*,[a]} and Roberto Improta^{*,[a, b]}

Abstract: Guanine quadruplexes (**G4**) are four-stranded DNA structures involved in biological processes and are promising candidates for potential nanotechnological applications. This study examines how the **G4** topology affects the electronic absorption and the delocalization of electron holes, which play a key role in charge transport and oxidative damage. Combining transient absorption spectroscopy with PCM/TD-DFT calculations both parallel (**P**) and antiparallel (**A**) **G4** are investigated, which are formed, respectively, by association of four TGGGGT strands and folding of the human telomeric sequence GGG(TTAGGG)₃. The experimental absorption spectra obtained upon photo-ionization of **A** and **P** are different. This is explained by the different topology of the two **G4**, as well as by hole delocalization between two stacked guanines, possible only in **P**. The spectral signature of delocalized hole in guanine-rich regions is provided and the chemical physical effects which rule the hole delocalization are discussed.

Electron holes in molecular systems play a key role in natural phenomena, such as photosynthesis or oxidative DNA damage, as well as in molecular electronics. The hole may be localized on a single building block of the system or delocalized over several of them. The extent of hole delocalization is a fundamental property which governs charge migration and charge transfer important to many biological and technological processes. Both aspects concern guanine quadruplexes (**G4**), which are four-stranded structures formed by guanine-rich DNA sequences. They are characterized by vertically stacked tetrads (**T**) composed of four guanines connected via Hoogsteen-type hydrogen bonds. **G4**, in which the occurrence of hole transport was evidenced,^[1] have been correlated, on the one hand, to the oxidative damage of telomeres,^[2] and, on the other, proposed as nano-wires for the development of pro-

grammable circuits.^[3] Here, we show that hole delocalization in **G4** may be affected by the relative orientation of guanines located in different tetrads (Scheme 1). We focus on two prototypical **G4** structures: parallel (**P**), resulting from association of four TGGGGT strands, and antiparallel (**A**), formed by folding of



Scheme 1. Antiparallel **A** (143d.pdb)^[6] and parallel **P** (244d.pdb)^[7] guanine quadruplexes formed by folding of the GGG(TTAGGG)₃ sequence and association of four TGGGGT strands, both stabilized by Na⁺ ions (green). “anti” guanines, in respect to the glycosidic bond, are represented in red while “syn” guanines in blue. For simplicity adenine and thymine bases are not shown.

the human telomeric sequence GGG(TTAGGG)₃, both stabilized by Na⁺ ions. Using time-resolved absorption spectroscopy we show that electron holes generated in **P** and **A** by photo-ionization^[4] exhibit different absorption spectra. This observation is interpreted by a thorough computational analysis, combining full quantum mechanical (QM) and mixed QM/molecular mechanics (QM/MM) calculations, which reveal that the degree of localization is different in **A** and **P**. This fundamental difference, together with the specific topology of each **G4**, accounts for the observed optical behavior. On this ground, we identify the spectral signature of hole-delocalization in **G4** as well as the main structural effects ruling this process.

Experimentally, the electron holes in **G4** were created via photo-ionization by low intensity 5 ns laser pulses at 266 nm as described in ref. [4b]. **P** and **A** were dissolved in phosphate buffer for which the pH was adjusted to 3 in order to prevent deprotonation of the guanine cation.^[4b, 5] The steady-state absorption spectra of **G4** at pH 3 were the same as at pH 7 (Figure SI-1) and show that the four-stranded structure is preserved under acidic conditions. Transient absorption signals were recorded on a 90 μs scale (Figure SI-2). The spectra at 50 μs are shown in Figure 1. Under our experimental condi-

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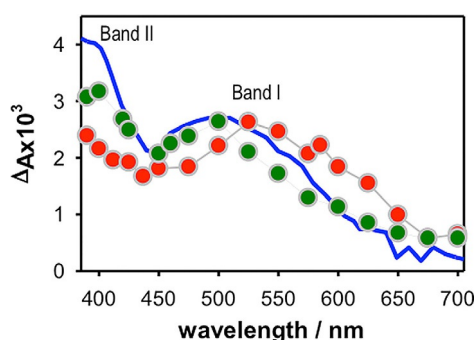


Figure 1. Transient absorption spectra determined at 50 μ s for P^+ (red) and A^+ (green) using the same excitation intensity (2.2 MW cm^{-2}) at 266 nm. The blue solid line corresponds to the spectrum of the dGMP cation^[5a] arbitrarily normalized. Estimated error: 10%.

tions only nucleobases undergo one- and two-photon ionization.^[4] Moreover, the concentration of ejected electrons is one order of magnitude lower than the **G4** concentration. Therefore, the spectra in Figure 1 correspond to **A** and **P** bearing only one electron hole, noted as A^+ and P^+ . The spectrum of A^+ is characterized by a band around 500 nm (I) and a second one peaking at 400 nm (II) with an intensity which is 17% higher than that of band I. It resembles the spectrum of the mononucleotide dGMP cation,^[5a] also shown in Figure 1, in which, however, the relative intensity of band II is 35% more intense than band I. Going from A^+ to P^+ , the spectrum exhibits remarkable changes: band I shifts clearly to longer wavelengths, while the intensity of band II further decreases.

The structural and electronic properties of the neutral and cationic species were then explored at the QM/MM (exploiting the ONIOM^[6] procedure as implemented in Gaussian09^[9]) and full QM levels using the crystal structure of **P** and **A** as reference.^[6,7] When included in the calculations, the phosphodeoxyribose backbone was always treated at the MM level with

the AMBER parm96.dat^[10] force field. In the case of **A**, for which the three tetrads are not topologically equivalent, we used only two computational models: **A3T**, sketched in Figure 2 (QM/MM calculations with three tetrads in the QM part) and **A3T_b** (full QM calculations including only the bases). A larger number of computational models were used for **P**, which is bigger than **A** but contains topologically equivalent tetrads. Thus, at the QM/MM level we studied several models differing in the number of tetrads included in the QM region (**P2T**, **P3T**, **P4T**), the location of the considered tetrads (inner or external) and the treatment of the metal cation in the central cavity (see below). In particular, model **P2T_b** (full QM calculations with two tetrads including only bases) was thoroughly explored. Our reference QM method was DFT with M052X^[5b,11] functional, combined with the 6-31G(d) basis set. Test calculations were also performed using B3LYP,^[12] CAM-B3LYP^[13] and ω B97XD^[14] functionals and more extended basis sets (6-31 + G(d,p)). Solvent effects were included by the PCM model in both QM and QM/MM calculations.^[15] A critical extensive discussion of our computational approach, successfully applied in the study of UV-induced processes in DNA (including radicals),^[4b,16] is reported in the Supporting Information.

Despite the small structural differences of the minima which result from all computational models used in the analysis, the electronic circular dichroism and absorption spectra computed for “neutral” **A** and **P** are in agreement with the experimental ones (see Supporting Information). The only significant difference is a uniform blue-shift of the computed spectra, which, besides possible inaccuracies of our electronic calculations, is due to the lack of vibrational and thermal effects.^[17] In order to account for this shift and to allow an easier comparison with the experimental spectra, we decided to shift the computed absorption energies by 0.50 eV (unshifted spectra are given in the Supporting Information). This value provides a good agreement between the computed and the experimental spectra of

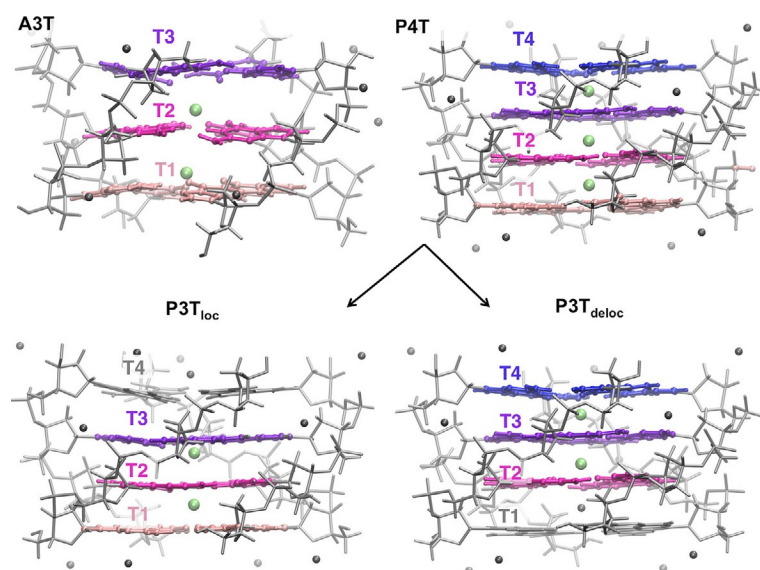


Figure 2. Systems studied at the QM/MM level for **A** and **P**. Balls and sticks represent the QM parts and tubes the MM parts. Inner QM- Na^+ are represented in green balls, whereas MM- Na^+ are shown in grey balls. Models studied at the QM level: **A3T_b** ($T1 + T2 + T3 + 2\text{Na}^+$) and **P2T_b** ($T1 + T2 + 2\text{Na}^+$).

guanine cation (see Supporting Information).^[18] For the cationic **G4**, we started our analysis with A^+ , which has an experimental spectrum similar to that of the free dGMP cation. Our calculations on $[A3T_b]^+$ and $[A3T]^+$ indicate that the hole is essentially localized over a single guanine, located in the external (**T3**) and internal (**T2**) tetrads, respectively (Figure 2). The computed absorption spectrum reproduces the main features of the experimental spectrum and, in particular, the relative intensities of bands I and II (Figure 3a).^[4b]

Hole localization is also predicted for $[P4T]^+$, in which a single base bears an excess of charge $\delta_q = 0.74$ a.u. The corresponding guanine adopts a geometry typical of the free guanine cation, whereas the other guanines conserve the geometry found for the free neutral guanine. Subsequently, we used smaller models, and, hence, more affordable from a computational point of view, in order to explore hole localization in **P** in a more complete way. Calculations on $[P2T]^+$ indicate also hole localization independently of whether inner or external tetrads are described at the QM level (see Supporting Information). For the $[P3T]^+$ model, however, both localized and delocalized minima were obtained depending on the starting geometry for the cation optimization (see below). A similar picture emerges from the study of $[P2T_b]^+$. When starting from the PCM/M052X optimized $[P2T_b]^+$ minimum, geometry optimization of $[P2T_b]^+$ predicts that the hole is almost perfectly delocalized over two guanines belonging to two different tetrads ($[P2T_b]^+_{deloc}$). In contrast, when starting from the stacking arrangement experimentally determined for **P**,^[7] geometry optimization of $[P2T_b]^+$ leads to a minimum with the hole localized on a single guanine, hereafter $[P2T_b]^+_{loc}$, which is more stable by ca. 0.3 eV than $[P2T_b]^+_{deloc}$. At this stage, we have

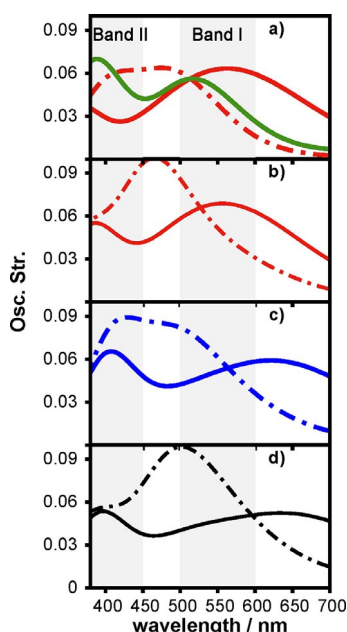


Figure 3. Computed absorption spectra for different models and different functionals. a) $[A3T]^+$ (green), $[P3T]^+_{loc}$ (red, dashed) and $[P3T]^+_{deloc}$ (red, solid) at the QM(PCM/TD-M052X/6-31G(d))/MM level of theory. $[P2T_b]^+_{loc}$ (dashed) and $[P2T_b]^+_{deloc}$ (solid) according to b) PCM/TD-M052X, c) PCM/TD-CAM-B3LYP and d) PCM/TD- ω B97XD calculations.

reached an important conclusion: for cationic parallel **G4** structures, minima with different degrees of hole delocalization are possible.

As a next step, we investigated the effects which influence the localization/delocalization equilibrium in **P**. For example, we found that symmetry effects, connected to the position of the inner cations, are key factors in modulating hole delocalization (Table SI-2). When the neutral species **PT3** is optimized starting from the crystal structure, where the inner Na^+ ions are not symmetric with respect to the three tetrads ($PT3_{loc}$, Figure 2), our calculations provide hole localization. On the contrary, when starting from structures where the two Na^+ ions are symmetrically “sandwiched” between the tetrads, hole delocalization is found ($PT3_{deloc}$, Figure 2). Furthermore, test calculations on $[P2T_b]^+_{loc}$ and $[P2T_b]^+_{deloc}$ (Table SI-1) show that closer stacking between bases favors hole delocalization. We also stress that this equilibrium is not significantly affected by the selected density functional and basis set or by the polarity of the embedding medium (see Supporting Information). Finally, conformational fluctuations—the inclusion of which at the full QM level is not yet feasible for the μs timescale—would likely favor hole localization.

All spectra computed for P^+ model systems (see Figure 3a and SI-6) are different from that of A^+ species. Both localized ($[P3T]^+_{loc}$) and delocalized ($[P3T]^+_{deloc}$) spectra show an inversion of the relative intensity of bands I and II compared to A^+ , the band I being more intense (Figure 3a). On the other hand, the position of band I strongly depends on hole localization in P^+ (Figure 3a). For $[P3T]^+_{loc}$ band I is blue-shifted compared to A^+ , and almost coalesces with band II. In contrast, for $[P3T]^+_{deloc}$ the band I is red-shifted with respect to that of A^+ . The spectra computed for the model $P2T_b$ present similar trends (Figure 3b and SI-6), which are full confirmed by test calculations adopting CAM-B3LYP and ω B97XD functionals. These spectra, which are shown in Figure 3b, provide a similar picture of the effect of hole delocalization in P^+ , which leads in particular to a significant red-shift of band I.

Comparison between Figures 1 and 3a shows that only the computed spectra corresponding to delocalized species exhibit all features of the P^+ experimental spectra, thus providing a second important conclusion for our study. Both larger hole delocalization (band I shift) and different stacking arrangements (relative intensity of bands I and II) are responsible for the different spectrum of P^+ and A^+ . Our conclusions are confirmed by test calculations using the B3LYP functional, which overestimates delocalization effects, and thus predicts delocalization for A^+ (Figure SI-6). Also for this antiparallel structure, hole delocalization is indeed accompanied by a relative increase of the intensity of band I (Figure SI-6).

Our calculations allow full assignment of the differences observed in the experimental spectra of P^+ and A^+ (Figure 2) and unveil consistent trends, which do not depend on the adopted computational approach; they were verified for many systems of different size (up to 10 minima have been considered). Independently of the degree of delocalization, the spectrum of cationic **G4** is affected by the presence of “neutral” guanines located close to the electron hole. On one hand,

transitions with predominant charge transfer (CT) character, involving the hole transfer from the guanine bearing the positive charge towards close lying bases, are present in both A^+ and P^+ . On the other hand, the different stacking arrangement in P and A is responsible for the spectral differences found in the case of localized holes in these $G4$. In particular, the stabilization of the bright state, $\pi\pi^*_2$ ($G^+ \rightarrow G^{+*}$, Figure S1-7 and S1-8), for parallel arrangements accounts for the larger intensity of band I in the P^+ spectrum compared to that of A^+ . In addition, if the hole is delocalized, this electronic transition acquires partial CT character (0.17 a.u.) between the guanines sharing the hole (Figure S1-8 and S1-9). The mixed $\pi\pi^*_2$ /CT character precisely explains the red-shift of the band I.

Our joint experimental and computational work has brought several new insights regarding hole transport in DNA.^[19] We provide evidence for the first time that electron holes may be delocalized in four-stranded structures; this phenomenon is strongly modulated by the $G4$ topology, being clearly favored by parallel arrangements. Hole delocalization in such structures is controlled by both conformational factors and the position of metal cations; its importance in charge migration has been highlighted.^[19e,20] More importantly, we identified the spectral fingerprint of delocalized holes, analyzing the associated electronic transitions. We showed that the effect of topology is not limited only on the spectra of delocalized holes but is also extended to those of localized ones. This impact is much more pronounced compared to that reported for the spectra neutral $G4$.^[21] Although electron holes in DNA have been treated quantum mechanically,^[22] no correlation has been established so far with their electronic absorption spectra. Yet, the knowledge of their spectral fingerprint could be very useful for distinguishing them in ultrafast transient experiments where other transient species potentially absorb.^[23] Our approach, based on photoionization by low energy UV light and QM calculations, provides precisely a tool that could be proved valuable for direct observation of electron holes in DNA.

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Conflict of interest

The authors declare no conflict of interest.

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- [1] a) S. P. Liu, S. H. Weisbrod, Z. Tang, A. Marx, E. Scheer, A. Erbe, *Angew. Chem. Int. Ed.* **2010**, *49*, 3313–3316; *Angew. Chem.* **2010**, *122*, 3385–3388; b) S. Delaney, J. K. Barton, *Biochemistry* **2003**, *42*, 14159–14165;

- c) J. Choi, A. Tanaka, D. W. Cho, M. Fujitsuka, T. Majima, *Angew. Chem. Int. Ed.* **2013**, *52*, 12937–12941; *Angew. Chem.* **2013**, *125*, 13175–13179; d) A. K. Thazhathveetil, M. A. Harris, R. M. Young, M. R. Wasielewski, F. D. Lewis, *J. Am. Chem. Soc.* **2017**, *139*, 1730–1733; e) J. Y. Wu, Z. Y. Meng, Y. P. Lu, F. W. Shao, *Chem. Eur. J.* **2017**, *23*, 13980–13985.
- [2] a) E. Fouquere, J. Lormand, A. Bose, H. T. Lee, G. S. Kim, J. F. Li, R. W. Sobol, B. D. Freudenthal, S. Myong, P. L. Opreko, *Nat. Struct. Mol. Biol.* **2016**, *23*, 1092–1100; b) A. M. Fleming, C. J. Burrows, *Chem. Res. Toxicol.* **2013**, *26*, 593–607.
- [3] G. I. Livshits, A. Stern, D. Rotem, N. Borovok, G. Eidelstein, A. Migliore, E. Penzo, S. J. Wind, R. Di Felice, S. S. Skourtis, J. Carlos Cuevas, L. Gurevich, A. B. Kotlyar, D. Porath, *Nat. Nanotechnol.* **2014**, *9*, 1040–1046.
- [4] a) A. Banyasz, T. Ketola, A. Muñoz-Losa, S. Rishi, A. Adhikary, M. D. Sevilla, L. Martínez-Fernández, R. Improta, D. Markovitsi, *J. Phys. Chem. Lett.* **2016**, *7*, 3949–3953; b) A. Banyasz, L. Martínez-Fernández, C. Balty, M. Perron, T. Douki, R. Improta, D. Markovitsi, *J. Am. Chem. Soc.* **2017**, *139*, 10561–10568; c) A. Banyasz, T. Ketola, L. Martínez-Fernández, R. Improta, D. Markovitsi, *Faraday Discuss.* **2018**, *207*, 181–197.
- [5] a) L. P. Candeias, S. Steenken, *J. Am. Chem. Soc.* **1989**, *111*, 1094–1099; b) L. Wu, K. Liu, J. Jie, D. Song, H. Su, *J. Am. Chem. Soc.* **2015**, *137*, 259–266.
- [6] Y. Wang, D. J. Patel, *Structure* **1993**, *1*, 263–282.
- [7] G. Laughlan, A. I. Murchie, D. G. Norman, M. H. Moore, P. C. Moody, D. M. Lilley, B. Luisi, *Science* **1994**, *265*, 520.
- [8] S. Dapprich, I. Komaromi, K. S. Byun, K. Morokuma, M. J. Frisch, *J. Mol. Struct.* **1999**, *461*, 1–21.
- [9] Gaussian 09, Revision A.1, M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski, D. J. Fox, Gaussian, Inc., Wallingford CT, **2009**.
- [10] W. D. Cornell, P. Cieplak, C. I. Bayly, I. R. Gould, K. M. Merz, D. M. Ferguson, D. C. Spellmeyer, T. Fox, J. W. Caldwell, P. A. Kollman, *J. Am. Chem. Soc.* **1995**, *117*, 5179–5197.
- [11] Y. Zhao, N. E. Schultz, D. G. Truhlar, *J. Chem. Theory Comput.* **2006**, *2*, 364–382.
- [12] A. D. Becke, *J. Chem. Phys.* **1993**, *98*, 5648.
- [13] T. Yanai, D. P. Tew, N. C. Handy, *Chem. Phys. Lett.* **2004**, *393*, 51–57.
- [14] J.-D. Chai, M. Head-Gordon, *Phys. Chem. Chem. Phys.* **2008**, *10*, 6615–6620.
- [15] a) S. Miertus, E. Scrocco, J. Tomasi, *Chem. Phys.* **1981**, *55*, 117–129; b) J. Tomasi, B. Mennucci, R. Cammi, *Chem. Rev.* **2005**, *105*, 2999–3093.
- [16] a) R. Improta, *Chem. Eur. J.* **2014**, *20*, 8106–8115; b) M. Deiana, B. Mettra, L. Martínez-Fernández, L. M. Mazur, K. Pawlik, C. Andraud, M. Samoc, R. Improta, C. Monneré, K. Matczyszyn, *J. Phys. Chem. Lett.* **2017**, *8*, 5915–5920.
- [17] F. J. Avila Ferrer, J. Cerezo, E. Stendardo, R. Improta, F. Santoro, *J. Chem. Theory Comput.* **2013**, *9*, 2072–2082.
- [18] Y. Rokhlenko, J. Cadet, N. E. Geacintov, V. Shafirovich, *J. Am. Chem. Soc.* **2014**, *136*, 5956–5962.
- [19] a) J. C. Genereux, J. K. Barton, *Chem. Rev.* **2010**, *110*, 1642–1662; b) K. Kawai, T. Majima, *Acc. Chem. Res.* **2013**, *46*, 2616–2625; c) A. Capobianco, T. Caruso, A. M. D'Ursi, S. Fusco, A. Masi, M. Scrima, C. Chatgililoglu, A. Peluso, *J. Phys. Chem. B* **2015**, *119*, 5462–5466; d) N. Renaud, M. A. Harris, A. P. N. Singh, Y. A. Berlin, M. A. Ratner, M. R. Wasielewski, F. D. Lewis, F. C. Grozema, *Nat. Chem.* **2016**, *8*, 1015–1021; e) S. Kanvah, J. Joseph, G. B. Schuster, R. N. Barnett, C. L. Cleveland, U. Landman, *Acc. Chem. Res.* **2010**, *43*, 280–287.
- [20] a) Y. Zhang, R. M. Young, A. K. Thazhathveetil, A. P. N. Singh, C. Liu, Y. A. Berlin, F. C. Grozema, F. D. Lewis, M. A. Ratner, N. Renaud, K. Siriwing, A. A. Voityuk, M. R. Wasielewski, D. N. Beratan, *J. Phys. Chem. Lett.* **2015**, *6*, 2434–2438; b) E. Beall, S. Ulku, C. Liu, E. Wierzbinski, Y. Zhang, Y. Bae,

- P. Zhang, C. Achim, D. N. Beratan, D. H. Waldeck, *J. Am. Chem. Soc.* **2017**, *139*, 6726–6735; c) M. Zarea, Y. Berlin, M. A. Ratner, *J. Chem. Phys.* **2017**, *146*, 114105.
- [21] A. I. Karsisiotis, N. M. Hessari, E. Novellino, G. P. Spada, A. Randazzo, M. W. da Silva, *Angew. Chem. Int. Ed.* **2011**, *50*, 10645–10648; *Angew. Chem.* **2011**, *123*, 10833–10836.
- [22] a) A. Capobianco, T. Caruso, A. Peluso, *Phys. Chem. Chem. Phys.* **2015**, *17*, 4750–4756; b) A. Kumar, M. D. Sevilla, *J. Phys. Chem. B* **2011**, *115*, 4990–5000; c) M. Pavanello, L. Adamowicz, M. Volobuyev, B. Mennucci, *J. Phys. Chem. B* **2010**, *114*, 4416–4423; d) T. Kubar, U. Kleinekathoefer, M. Elstner, *J. Phys. Chem. B* **2009**, *113*, 13107–13117; e) N. Renaud, Y. A. Berlin, M. A. Ratner, *Proc. Natl. Acad. Sci. USA* **2013**, *110*, 14867–14871;
- f) Y. A. Mantz, F. L. Gervasio, T. Laino, M. Parrinello, *Phys. Rev. Lett.* **2007**, *99*, 058104; g) C. J. Lech, P. Anh Tuan, M.-E. Michel-Beyerle, A. A. Voityuk, *J. Phys. Chem. B* **2013**, *117*, 9851–9856; h) P. B. Woiczikowski, T. Kubar, R. Gutierrez, G. Cuniberti, M. Elstner, *J. Chem. Phys.* **2010**, *133*, 035103.
- [23] M. A. Harris, A. K. Mishra, R. M. Young, K. E. Brown, M. R. Wasielewski, F. D. Lewis, *J. Am. Chem. Soc.* **2016**, *138*, 5491–5494.

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