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Boronic Acid-Functionalized Oxide-Free Silicon Surfaces for the Electrochemical Sensing of Dopamine

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ABSTRACT. Boronic acid monolayers covalently bound to hydrogen-terminated Si(111) surfaces have been prepared from the UV-directed hydrosilylation reaction of 4-vinylbenzeneboronic acid. X-ray photoelectron spectroscopy (XPS) analysis of the modified surface revealed characteristics peaks from the attached organic molecule with the expected molecular composition and without oxidation of underlying silicon. From XPS data, the surface

coverage was estimated to be ca. 0.34 ± 0.04 ethylbenzene boronic acid chain per surface silicon atom (i.e. $(4.4 \pm 0.5) \times 10^{-10}$ mol cm⁻²) which is consistent with a densely packed monolayer. The electrochemical impedance spectroscopy measurements performed at pH 7.4 in the presence of the Fe(CN)₆³⁻/Fe(CN)₆⁴⁻ reporter couple showed specific dopamine-induced changes, as a result of the binding of the guest molecule to the immobilized boronate species. The charge transfer resistance (R_{ct}) was found to decrease from 4.9 MΩ to 14 kΩ upon increasing the dopamine concentration in the range 10 μM – 1 mM. Furthermore, the presence of the interfering ascorbic acid until a concentration of 10 mM did not change significantly the electrochemical response of the functionalized surface. Comparative electrochemical data obtained at the reference ethylbenzene monolayer provided clear evidence that the immobilized boronic acid units were responsible for the observed changes.

1. Introduction

Owing to the high affinity of boronic acids with some diols (*e.g.*, sugars and catecholamines), their irreversible immobilization on electrode surfaces has appeared as a promising strategy toward the development of electrochemical (bio)sensing devices.^{1,2} Among the procedures used for attaching such functional units onto conducting surfaces, the incorporation in self-assembled monolayers^{3,4,5,6,7,8,9,10,11} and in polymer films^{12,13,14,15,16,17} has been the most explored and some of these boronic acid-modified electrodes have been successfully applied for the electrical detection of sugars^{1,4,12,13} and catecholamines.^{1,6,7,14,16,17} Compared with polymer films, the monolayer approach offers some advantages because it allows a molecular-level control of the functional layer structure (homogeneity, ordering and packing density) which makes, for example, more sensitive electrochemical detection. In that context, the functionalization of

technologically important semiconducting surfaces, such as oxide-free, hydrogen-terminated silicon (Si-H) with molecular receptor-incorporating monolayers could constitute an attractive approach for designing novel interfaces for chemical/biological sensing. Indeed, Si-H has been demonstrated to be a particularly appealing substrate for electrochemical applications owing to the ease and reproducibility of its preparation,^{18,19} its well-defined structure, its very low density of electrically active surface defects,²⁰ and its propensity to be chemically modified with organic monolayers linked through nonpolar and robust interfacial Si-C bonds.^{21,22,23,24,25,26,27} Moreover, unlike most of traditional electrodes (e.g. glassy carbon) used as working electrodes for the electrochemical sensing, silicon-based devices can be directly integrated within existing electronic circuitry^{28,29} and the numerous existing technological processes used for the micro- and nanopatterning of silicon are mature enough for producing highly miniaturized functional electronic components. For these reasons, such functional interfaces should show great promise for the monitoring of molecules of biological interest. Furthermore, it is noteworthy that the electrochemical detection of dopamine from a functionalized silicon surface is unprecedented.

Herein, we describe our results on the derivatization of Si(111)-H surfaces with boronic acid-terminated monolayers. These surfaces were produced by direct hydrosilylation of Si(111)-H with vinylbenzeneboronic acid. They were thoroughly characterized by various experimental techniques including contact angle goniometry, X-ray photoelectron spectroscopy (XPS) and atomic force microscopy (AFM). The electrochemical response of the boronic acid-modified surfaces was then used as the signal transduction pathway of the binding event between the immobilized boronic acid moieties and dopamine. To highlight the role of boronic acid group in the dopamine detection, a reference ethylbenzene monolayer on silicon was also prepared and electrochemically studied in the presence of the target guest.

It is worth emphasizing that the quantitative determination of dopamine in physiological fluids is of great interest in medicine. This molecule plays an important role in neurotransmission and other physiological processes.³⁰ Low levels in dopamine have been correlated to some neurodegenerative diseases, such as Parkinson's disease, or to mental disorders, such as schizophrenia and anorexia, whereas high levels may be associated with other health problems like pheochromocytoma.³¹ Different analytical techniques have been used to detect dopamine in biological samples, such as high-performance liquid chromatography (HPLC),³² capillary electrophoresis,³³ HPLC coupled mass spectroscopy,³⁴ electrochemistry,^{35,36} and field-effect transistors.³⁷ The electrochemical approach is very attractive because it is an easily scalable, cost-effective method and allows dopamine determination with fast response time. However, for modified electrode-based electrochemical detection, attention must be paid to both the morphology and the ordering of the sensitive molecular layer which may impact on the measurement reproducibility.

2. Experimental Section

2.1. Reagents. Acetone (MOS electronic grade, Erbatron from Carlo Erba), anhydrous ethanol (RSE electronic grade, Erbatron from Carlo Erba), toluene (HPLC grade, Fisher Chemical) and methylene chloride (RSE electronic grade, Carlo Erba) were used without further purification. The chemicals used for cleaning and etching of silicon wafer pieces (30% H₂O₂, 96-97% H₂SO₄, and 40% NH₄F solutions) were of semiconductor grade (Riedel-de-Haën). 4-vinylbenzeneboronic acid (98%, Sigma-Aldrich), styrene (>99%, stabilized with 4-*tert*-butylcatechol, Sigma-Aldrich) and dopamine hydrochloride (99%, Alfa-Aesar) were used as received.

2.2. Covalent Attachment of the Ethylbenzene and Ethylbenzeneboronic Acid Monolayers on Silicon. A single side polished silicon(111) shard ($1.5 \times 1.5 \text{ cm}^2$, 1-5 ohm cm resistivity, *p*-type, boron doped, thickness = $525 \pm 25 \mu\text{m}$, Siltronix) was sonicated for 10 min successively in acetone, ethanol, and ultrapure 18.2 M Ω cm water (Purelab Classic UV from Veolia Water-STI). It was then cleaned in 3:1 v/v concentrated H₂SO₄/30% H₂O₂ at 100 °C for 30 min, followed by copious rinsing with ultrapure water.

Caution: The concentrated H₂SO₄:H₂O₂ (aq) piranha solution is very dangerous, particularly in contact with organic materials, and should be handled extremely carefully.

The surface was etched with ppb grade 40% aqueous argon-degassed NH₄F for 15 min to obtain atomically flat, oxide-free hydrogen-terminated Si(111) surface (Si-H).³⁸ It was then dipped in argon-degassed ultrapure water for several seconds, dried under an argon stream, and transferred immediately into a Pyrex Schlenk tube containing 4-vinylbenzeneboronic acid (or styrene) at 10 mM in toluene, previously deoxygenated at 100 °C for 1 h at least and then allowed to cool down to ca. 30-40°C before introducing the Si-H substrate. After Ar was bubbled through the solution for 30 min, the Si-H surface was irradiated in a Rayonet photochemical reactor (300 nm) for 3 h. The ethylbenzeneboronic acid-modified silicon surface (**Si-EBBA**) [or ethylbenzene-modified silicon surface (**Si-EB**)] was rinsed copiously with toluene, ethanol and methylene chloride, then dried under an argon stream.

2.3. Surface Characterizations. XPS measurements were performed using a VSW HA100 system with MgK α photons at 54° incidence and a 45° takeoff angle. This angle is used to increase surface sensitivity, especially to check for Si-O bonds appearance. Spectra were fitted with a convolution of Lorentzian and Gaussian profiles using standard procedures. Binding energies are referenced to the main C 1s peak which is set at 285.0 eV.

AFM images were recorded with a PicoSPM II microscope from Molecular Imaging using conventional gold coated-silicon nitride (contact mode) tips from Scientec.

Water contact angles were measured with a homemade goniometer under ambient conditions using a horizontal light beam to illuminate the liquid droplet. Uncertainty $\pm 2^\circ$.

2.4. Electrochemical Characterizations. The impedance spectroscopy measurements were performed with an Autolab electrochemical analyzer (PGSTAT 20 potentiostat/galvanostat from Eco Chemie B.V., equipped with the GPES/FRA software) in a self-designed three-electrode Teflon cell. The amplitude of the ac signal was 10 mV and the frequency was varied from 100 kHz to 0.1 Hz with a logarithmic distribution (50 frequencies). The working electrode, modified Si(111), was pressed against an opening in the cell bottom using a Viton O-ring seal. An ohmic contact was made on the previously polished rear side of the sample by applying a drop of an In-Ga eutectic (Alfa-Aesar, 99.99%). The electrochemically active area of the Si(111) surface (namely, 1.1 cm^2) was estimated by measuring the charge under the voltammetric peak corresponding to the ferrocene oxidation on Si(111)-H and comparing this value to that obtained with a 1 cm^2 -Pt electrode under the same conditions. The counter electrode was a platinum foil and a potassium chloride saturated calomel electrode (SCE) was used as the reference electrode. The applied potential during the impedance measurements was fixed at the open circuit potential of the electrochemical cell in contact with the redox couple $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$, namely 0.15 V vs SCE. Potassium ferrocyanide (99%, Acros) and ferricyanide (99%, Sigma-Aldrich) were used at a concentration of 2 mM in buffered aqueous solution (0.1 M phosphate buffer, pH 7.4 or 0.1 M acetate buffer, pH 4).

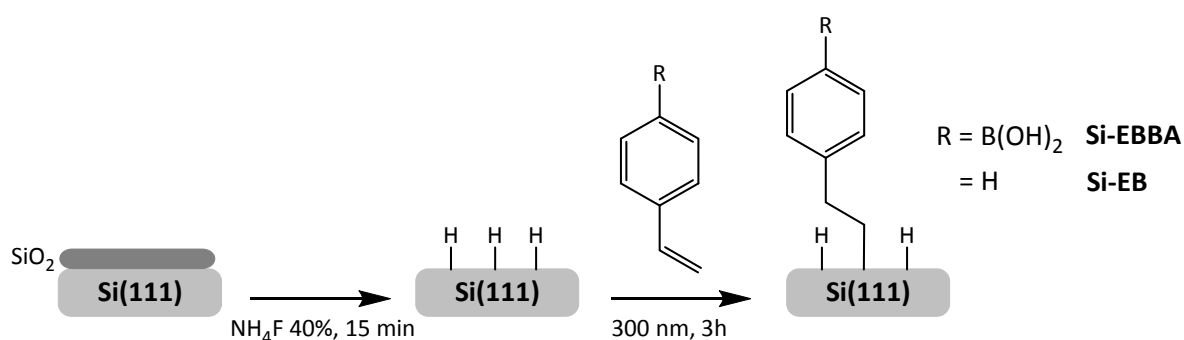
All electrochemical measurements were carried out inside a homemade Faraday cage at room temperature ($20 \pm 2^\circ\text{C}$) and under a constant flow of argon.

3. Results and Discussion

The **Si-EBBA** surface was prepared from the photochemical reaction of 4-vinylbenzeneboronic acid at 300 nm for 3 h in the presence of Si(111)-H (Scheme 1). This direct hydrosilylation route yielded a boronic acid-terminated monolayer covalently bound to the silicon surface through strong Si-C bonds.³⁹ For comparison, a Si(111) surface was similarly modified by an ethylbenzene monolayer from styrene (**Si-EB**).

Static contact angles measured with water reveal, as expected, the hydrophilic character of the **Si-EBBA** surface compared with **Si-EB**, $40 \pm 2^\circ$ vs $94 \pm 2^\circ$. These values are in close agreement with those reported for other boronic acid-terminated monolayers bound to gold ($\sim 32^\circ$)⁹ and ethylbenzene-modified silicon (90°)⁴⁰ surfaces. Such a result precludes thus the formation of a mixed monolayer with some alkene groups protruding to the surface resulting from the reaction of boronic acid groups with either Si-H or some hydroxyl species present on the silicon surface.⁴¹

SCHEME 1. Preparation of the Boronic Acid- and Ethylbenzene-Modified Silicon(111) Surfaces



XPS. The X-ray photoemission spectroscopy (XPS) analysis of the silicon surface derivatized with the **EBBA** monolayer reveals characteristic peaks from the silicon substrate itself and from

the C 1s, O 1s and B 1s core levels of the attached organic molecule (Figure S1, Supporting Information). The experimental atomic ratios calculated from the peak areas of O 1s, C 1s and B 1s were 2.7, 8.9 and 1.0, respectively (Table S1), which are in close agreement with the theoretical ratios calculated from the assumed structure of the boronic acid monolayer, namely 2, 8 and 1, respectively.

The high-resolution C 1s peak can be decomposed into two components at 285.0 and 286.7 eV (Figure 1). The first one corresponds to unresolved contributions from both aromatic and alkyl carbon atoms, whereas the weaker peak at 286.7 eV is reasonably attributed to heteroelement-bound carbon (C-B). The C-C + C=C/C-B atomic ratio (determined from the areas under the C-C, C=C and C-B peaks) was estimated at 8.0, which is close to the theoretical ratio of 7. The high-resolution B 1s spectrum shows the presence of a peak located at 191.9 eV attributable to boronic acid moieties superimposed with the broad Si satellite peak at 186.2 eV. These values are in good agreement with published values for boronic acid monolayers on gold³ and chlorobenzeneboronic acid.⁴²

Moreover, the Si 2p signal shows as expected the presence of a principal component at 99.7 eV attributed to Si-Si bonds (Fig. 1c). Careful examination of the high binding energy side of Si 2p (near 104 eV binding energy) does not show the presence of Si-O bonds demonstrating no oxidation (or at a very low extent) of the underlying silicon surface and the formation of a boronic acid-terminated Si-C-linked monolayer.

The thickness of the organic monolayer (d) has been estimated from the attenuation of the angular averaged silicon signal according to

$$d = -\lambda_{\text{Si}}^{\text{mono}} \ln \left(\frac{I_{\text{Si,mono}}}{I_{\text{Si,clean}}} \right) \quad (1)$$

where $\lambda_{\text{Si}}^{\text{mono}}$ is the mean free path of the Si 2p electrons travelling through the organic layer ($\lambda_{\text{Si}}^{\text{mono}} = 3.6 \text{ nm}$)⁴³, $I_{\text{Si,mono}}$ and $I_{\text{Si,clean}}$ are the Si 2p signals collected on grafted silicon and clean, uncovered Si(111)-H surface, respectively.

The value of 1.4 nm found by this method is in close agreement with that deduced from calculations of energy minimization using the semiempirical PM3 method, considering the boronic acid units lifted up (1.0 nm). From XPS data and using the treatment reported by Cicero et al.,⁴⁴ the surface coverage can be estimated at 0.34 ± 0.04 ethylbenzene boronic acid chain per surface silicon atom, which is in perfect line with coverages determined for other UV-prepared monolayers derived from phenylacetylene.⁴⁴ A specific area of about 35 \AA^2 per bound ethylbenzeneboronic acid can thus be deduced, which is slightly greater than the maximum packing density of 25 \AA^2 per molecule calculated by considering a (2×1) arrangement of aryl units on Si(111).⁴⁵ This indicates a reasonably dense packing of the boronic acid-terminated chains. Our coverage value is however smaller than those estimated for other monoaromatic monolayers prepared under different conditions (~ 0.50), i.e. electrochemically⁴⁵ or thermally.⁴⁶

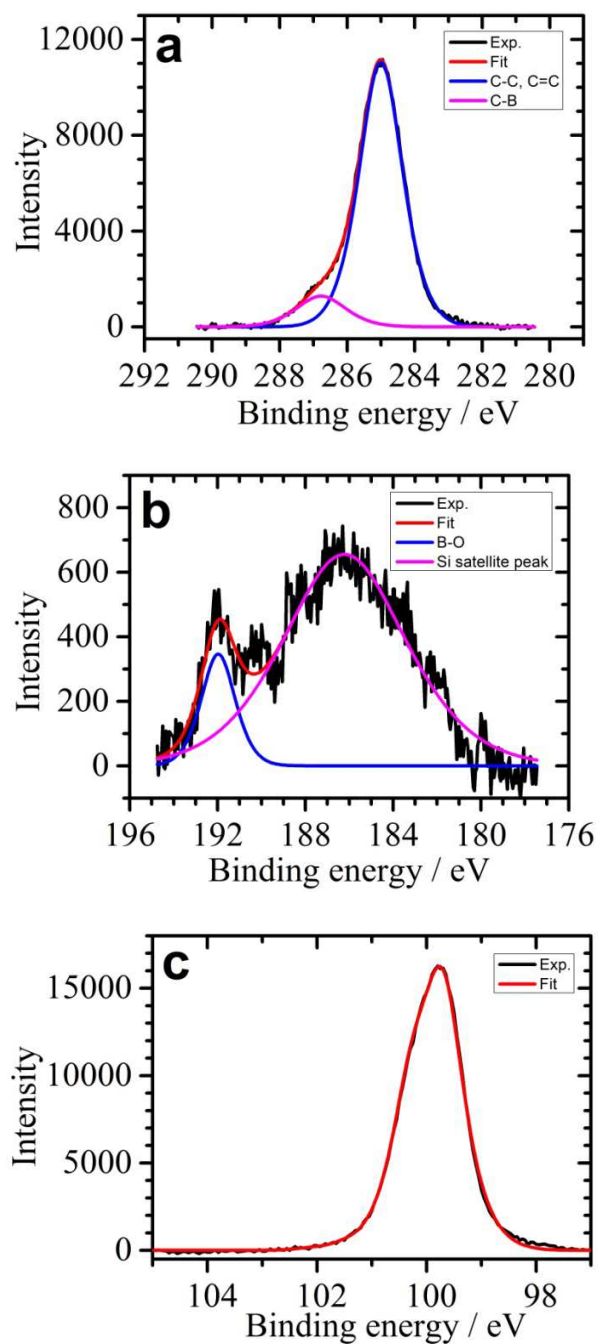


Figure 1. High-resolution XPS C 1s (a), B 1s (b) and Si 2p (c) spectra of **Si-EBBA**. Black lines are experimental data and red lines are the fitted curves that were generated using combinations of Gaussian-Lorentzian functions, each corresponding to a different contribution. Spectra have been recorded at 45° takeoff angle.

AFM. Atomic force microscopy (AFM) analysis of the **Si-EBBA** surface shows a similar structure to that of Si(111)-H, with atomically flat, almost defect-free terraces consisting of about 3-Å-high steps (Figure 2). This result indicates that the covalent attachment of the boronic acid-terminated monolayer did not alter the underlying surface morphology. Moreover, any contrast in friction images is not observed and any adventitious material is not removed upon a prolonged AFM tip scanning on the same place. Therefore, it can be concluded that the molecular monolayer continuously covers the substrate without forming isolated patches and islands.

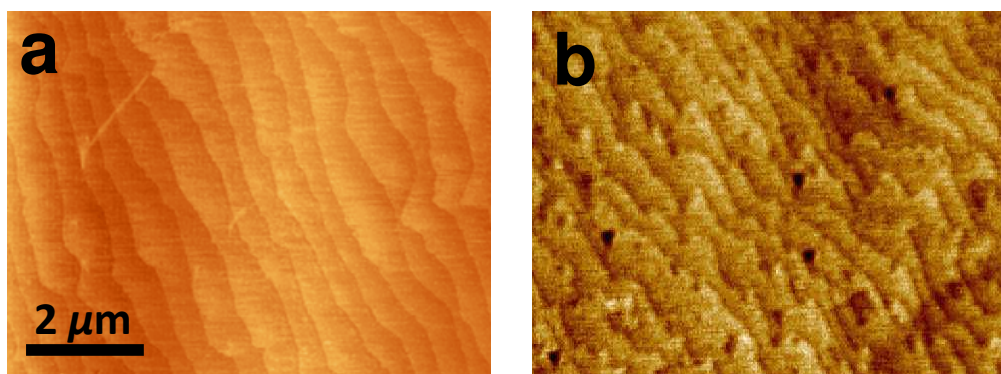


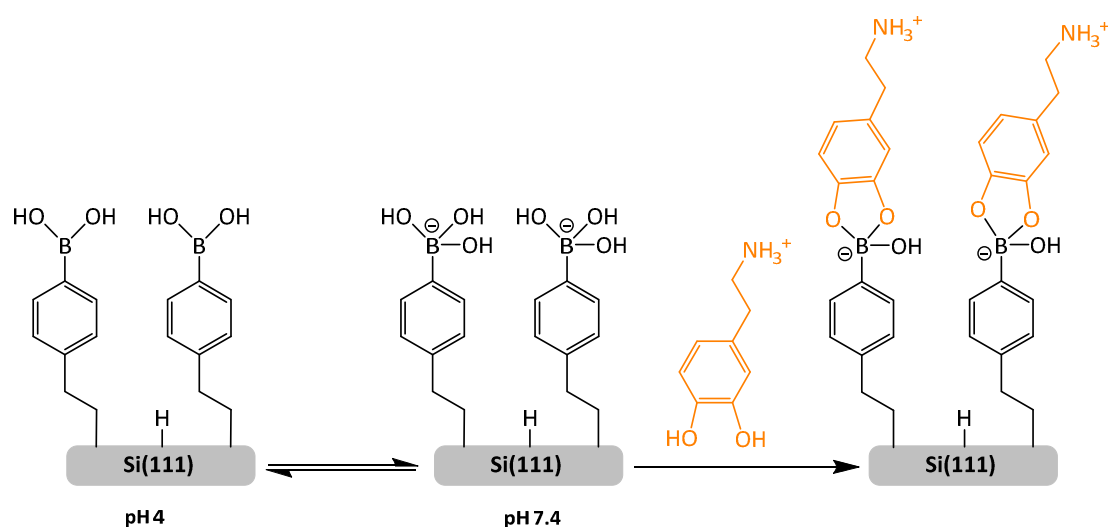
Figure 2. Contact-mode AFM images of Si(111)-H (a) and **Si-EBBA** (b) surfaces.

Electrochemical Sensing of Dopamine (DA) Using Impedance Spectroscopy. Electrochemical impedance spectroscopy (EIS) was used to quantitatively monitor the reaction of DA with the immobilized boronic acid moieties of the monolayer using the negatively charged $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$ as the reporter couple. Since the interaction of phenylboronic acid with DA is expected to be pH dependent, two different pH values were selected for sensing experiments, namely pH 4.0 and 7.4. The pH value of 7.4 has been preferentially chosen to detect the binding of DA to the functionalized surface but also to prevent the polymerization of DA occurring

usually at $\text{pH} > 8-8.5$.⁴⁷ Moreover, since the two protonation sites of dopamine have $\text{p}K_a$ values of about 8.7 and 10.0 for the catechol and amine moieties, respectively,⁴⁸ dopamine is predominantly in the form of a catechol molecule substituted by an ethylammonium chain at the two selected pH values.

The hydroxylation of phenylboronic acid occurs at weakly basic pH, and under these conditions the resulting tetrahedral benzenboronate ion can undergo covalent interactions with diols compounds (Scheme 2).^{1,49,50} This can be demonstrated by ^{11}B NMR measurements in solution. The formation of the boronate form upon the dopamine binding at pH 7.4 was supported by the presence of a signal at ca. 10 ppm (Fig. S2), which was characteristic of the tetrahedral sp^3 boron.⁵¹ In contrast, the ^{11}B NMR spectrum of a vinylbenzeneboronic acid solution at pH 4 showed the presence of one peak at ca. 29 ppm which was assigned to a trigonal sp^2 boron (Fig. S2). This signal was not significantly shifted after the addition of dopamine which indicates the absence of complexation at this pH value.

SCHEME 2. pH-Dependent Boronic Acid-Modified Silicon(111) Surface and Formation of the Boronate Ester Complex upon the Binding of Dopamine



Furthermore, it is worth mentioning that the EIS measurements were performed after stabilization of the modified silicon surface (**Si-EBBA** or **Si-EB**) in the tested electrolytic buffer for at least 2 h. This enabled to ensure a good reproducibility. Using such conditions, the relative uncertainty on the impedance values measured for equilibrated **Si-EBBA** or **Si-EB** in the absence or presence of dopamine was estimated at 10%.

Figure 3 shows the EIS spectra in the frequency range $10^5 - 10^1$ Hz of the **Si-EBBA** surface at two pH values in the absence and in the presence of DA and using the $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$ redox couple at the same concentration. The data are presented as plots of the imaginary part (Z'') vs the real part (Z') of the complex impedance ($Z = Z' + iZ''$) as a function of frequency f (Nyquist diagram). All Nyquist plots are characterized by a semicircle, the diameter of which is dependent on both the pH value and the presence or not of DA. The experimental data can be reasonably fitted assuming a Randles equivalent electrical circuit (Scheme 3). For this circuit model, the diameter of the semicircle on the Z' axis corresponds to the resistance of charge transfer R_{ct} . In the absence of DA, the R_{ct} value determined at pH 7.4 is almost two orders of magnitude higher than that estimated at pH 4, namely 4.9 M Ω against 60 k Ω (Fig. 3a). This demonstrates that charge transfer kinetics of the $\text{Fe}(\text{CN})_6^{3-/4-}$ couple at the modified electrode is much slower at pH 7.4 than at pH 4. Such results are not surprising and can be easily explained by electrostatic effects. At pH 4, the boronic acid units are globally neutral while at pH 7.4, the diffusion of the redox probe across the monolayer is totally impeded due to strong electrostatic repulsions between the formed boronate anion $\text{B}(\text{OH})_3^-$ and $\text{Fe}(\text{CN})_6^{3-/4-}$.

Now, in the presence of DA, Nyquist plots at pH 4 are not significantly modified because the guest is not able to bind to the immobilized boronic acid units (Fig. 3b). In contrast, a drastic decrease of R_{ct} upon increasing concentrations of DA is observed at pH 7.4 (Fig. 3c). The R_{ct} –

(DA concentration) calibration curve displayed in Fig. 3d is characterized by an initial linear plot at low DA concentrations (< 0.2 mM) followed by a plateau for larger concentrations indicating that all the accessible binding sites in the monolayer are occupied. The detection limit is found to be below $10 \mu\text{M}$. We have also checked that the addition of ascorbic acid, which is the principal interfering species in the dopamine detection, into a DA solution (at 1 mM) did not affect significantly the EIS response of the **Si-EBBA** surface until a tested concentration of 10 mM.

Also interestingly, the original state of this electrochemical sensor can be recovered after immersion of **Si-EBBA** in an acid medium ($\text{pH} \sim 1-2$) for few min. After this relatively simple chemical treatment, the EIS response is found to be similar to that before addition of dopamine, indicating that the free form of the immobilized boronic acid moieties has been regenerated.

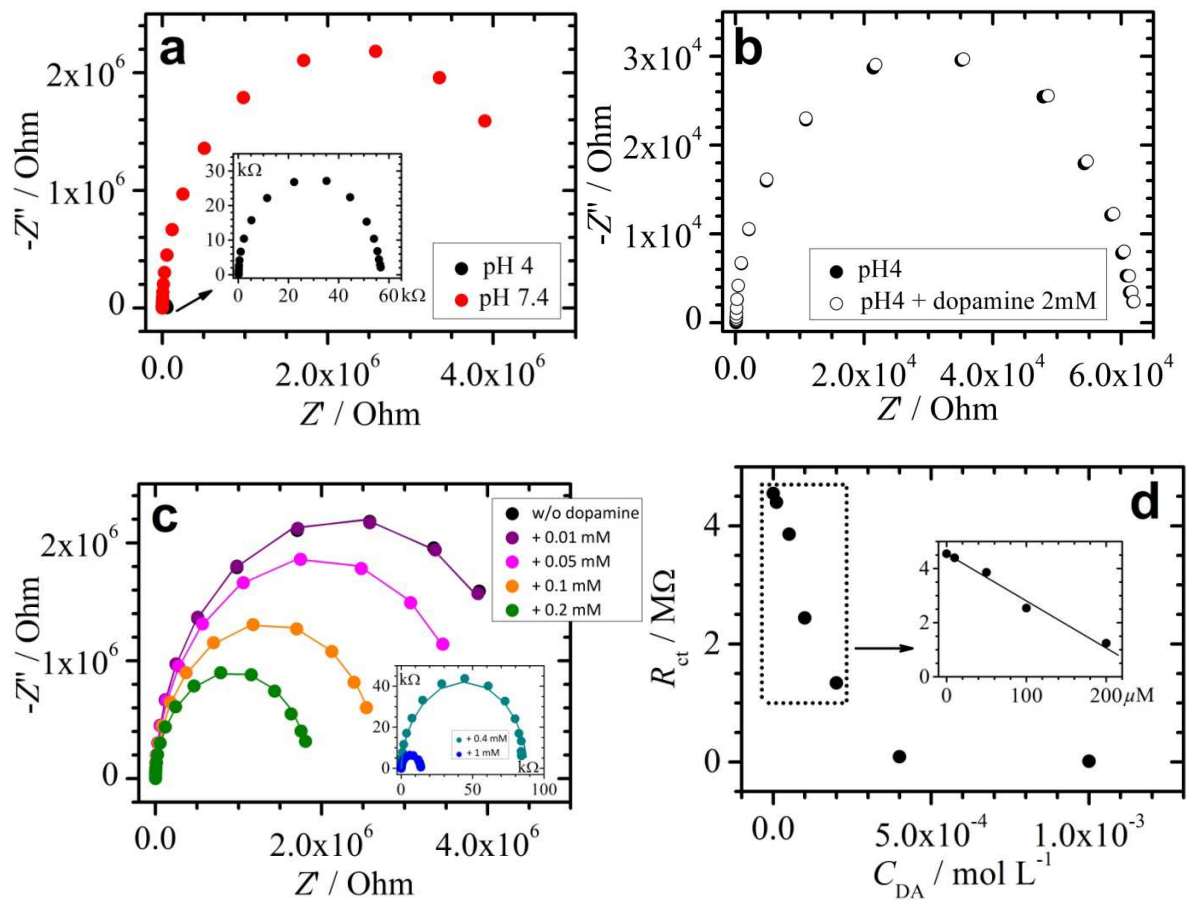
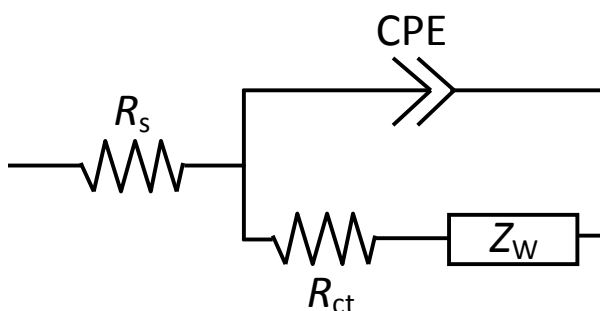


Figure 3. Nyquist ($-Z''$ vs. Z') plots for the **Si-EBBA** surface at pH 4 and 7.4 in the absence of dopamine (a), at pH 4 in the presence of dopamine at 2 mM (b), and at pH 7.4 with increasing concentrations of dopamine (c). For (a-c), the applied potential was 0.15 V vs SCE and $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$ were present at 2 mM. The filled dots in (c) are the experimental data and the solid lines correspond to the fitted curves assuming a Randles equivalent electrical circuit. (d) Corresponding R_{ct} – (DA concentration) plot.

SCHEME 3. Proposed equivalent electrical circuit for electron transfer of $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$ reporter couple upon the binding of dopamine to the **Si-EBBA** surface.^a



^a R_s is the solution resistance, CPE is a constant phase element, its impedance can be described as $Z_{\text{CPE}} = \frac{1}{(i\omega)^\alpha C}$ where C is the experimentally observed capacitance, ω is the angular frequency and α is generally a value between 0.9 and 1 ($\alpha = 1$ for an ideal capacitor). Z_w is the impedance due to mass transfer of the redox species to the electrode described by Warburg, R_{ct} is the charge-transfer resistance.

We have checked that the observed R_{ct} changes were ascribed to the electrochemical activity of the redox probe alone. In the absence of the redox probe, the voltammetric oxidation of dopamine in solution to *o*-dopaminoquinone gives a quasi-reversible redox process at a formal potential $E^{\circ'} = 0.35$ V and 0.26 V (average of the anodic and cathodic peak potentials) at pH 4 and 7.4, respectively (Fig. S3). Since the EIS measurements have been performed at 0.15 V vs

SCE, it can be considered that dopamine was not electrochemically oxidized under such conditions at both considered pH values.

Moreover, comparison with EIS data obtained with the **Si-EB** surface provides clear evidence that DA-induced changes are ascribed to the immobilized boronic acid units. Indeed, as shown in Figure 4, the addition of DA to the ethylbenzene monolayer does not lead to significant changes in the R_{ct} values.

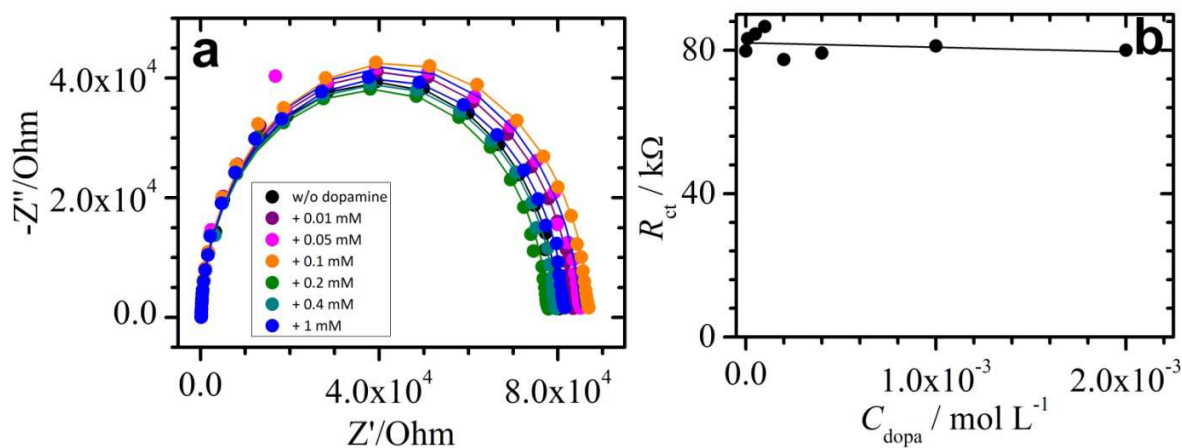


Figure 4. Nyquist ($-Z''$ vs. Z') plots for the **Si-EB** surface at pH 7.4 + 2 mM $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$ in the absence of dopamine and with increasing concentrations of dopamine (a). Applied potential: 0.15 V vs SCE. The filled dots are the experimental data and the solid lines correspond to the fitted curves assuming a Randles equivalent electrical circuit. (b) Corresponding R_{ct} – (DA concentration) plot.

The observed trend for the R_{ct} decrease at the **Si-EBBA** surface with DA concentration may seem counterintuitive and thus requires some clarification. As a matter of fact, electrostatic, electronic and steric effects can be involved to explain the facilitated charge transfer of $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$ upon binding of DA. At pH 7.4, the electrostatic repulsions between the negatively

charged surface and the redox probe are believed to be gradually lowered upon the binding of DA because the negative charge on the boron atom is in fact counterbalanced by the positive charge of the ammonium group of DA. Steric effects can be also considered to account for DA-induced impedance changes. The binding of DA to the immobilized boronate species may lead to some monolayer disordering caused by the steric hindrance of DA. Due to these steric constraints, the penetration of the redox probe between the molecular chains would therefore be made easier. Another plausible explanation could be related to a dopamine-induced modification of electron transfer kinetics. As elegantly demonstrated for dopamine-modified electrodes,⁵² the dopamine moieties bound to the silicon surface could lead to an electron transfer activation toward the redox probe in solution, acting as catalytic sites for adsorption or electron transfer or decreasing the hydrophobicity of the silicon surface.

4. Conclusions

In this work, boronic acid-terminated monolayers have been covalently bound to oxide-free hydrogen-terminated Si(111) surfaces using a one-step hydrosilylation route. This grafting procedure resulted in the formation of densely packed monolayers with a surface coverage of 0.34 ± 0.04 ethylbenzene boronic acid chain per surface silicon atom. The binding of dopamine to the immobilized boronic acid moieties at pH 7.4 led to an easier charge transfer of the $\text{Fe}(\text{CN})_6^{3-}/\text{Fe}(\text{CN})_6^{4-}$ reporter couple. Such functionalized surfaces showed a dopamine-sensitive impedimetric response with a detection limit below $10 \mu\text{M}$ and without interference of the ascorbic acid.

It is worth recalling that biologically meaningful concentrations of dopamine under *in vivo* conditions are usually in the range of 10-50 nM.⁵³ However, much larger concentration levels (until $50 \mu\text{M}$) have been detected when the dopamine release was electrically stimulated.⁵⁴

Consequently, the system proposed herein would be capable of detecting such concentrations. Moreover, it is believed that the detection sensitivity could be significantly improved using either a more sensitive electrochemical technique (e.g. differential pulse voltammetry with corrected background signals) or a miniaturized configuration (e.g. chemical field-effect transistor) which would be more appropriate for detection of dopamine within the mammalian brain.

Finally, some characteristics of the responsive monolayer could be also optimized in order to enhance the sensitivity parameter. Towards this goal, the dilution of the boronic acid-terminated chains with other molecular chains could be a promising strategy. Indeed, this should enable control of the surface coverage of the boronic acid units and improvement of both the quality and the packing density of the resulting mixed monolayer. In the case where the diluting chains are electrochemically active, we anticipate that the electrochemical detection of dopamine could operate without the need of a reporter couple in solution. In the future, silicon-based electronic devices incorporating such functional monolayers could be exploited for the real-time monitoring of levels of dopamine under physiological conditions.

Finally, it is obvious that the proposed approach that combines monolayer system and electrochemical detection method could be extended to other diols, such as sugars.

Supporting Information. Supplementary figures showing XPS survey spectrum of the boronic acid-modified silicon surface, ^{11}B NMR spectra of vinylbenzeneboronic acid in the absence and presence of dopamine and cyclic voltammograms of dopamine at a Pt working electrode. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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Author Contributions

The manuscript was written through contributions of all authors.

Notes

The authors declare no competing financial interest.

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