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HAL Id: cea-01478836
https://cea.hal.science/cea-01478836
Submitted on 28 Feb 2017

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Evidence of Mixed-Valence Hydrated Europium-Chloride Phase in Vacuum by Means of Optical and Electronic Spectroscopies

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1. INTRODUCTION

Inorganic phosphors used in mercury vapor lamps and cathode-ray tubes have been doped with lanthanides and especially europium ions. Actually, Eu$^{3+}$ ions, as luminescent species, are of strong interest, not only in fundamental but also in many photonic applications, in particular in organic light emitting diodes (OLEDs), lasers, optical communications, and as luminescent tags in biological systems. Their development at the origin of new design concepts in OLEDs and in liquid-luminescent tags in biological systems.5, 6 Their development is at the origin of new design concepts in OLEDs and liquid-crystal based display and lighting applications as well as for analytical applications as chemical and biological sensors,11 possibly associated with π-conjugated molecular ligands playing part of the antenna.8–11 Recently, rare-earth ions have been shown to present specific properties such as long coherence times of their optical transitions12 of possible interest for qubit applications.13–15

Eu$^{3+}$ luminescence is characterized by narrow atomic-like spectral lines. The emitted light mainly originates from the electronic transition from the $^{5}D_{0}$ excited energy level to the $^{7}F_{J}$ ($J = 1$ to $4$) levels inside the $4f$ electron shell.16 A drawback of europium in many application is its very low absorption cross-section at UV wavelengths with molar absorption coefficients $\varepsilon < 1 \text{ mol}^{-1} \cdot \text{L} \cdot \text{cm}^{-1}$. As a consequence, only a small amount of incident light is directly absorbed, leading to weak direct excitation of $4f$ levels. The excited Eu$^{3+}$ ions then relax radiatively in the visible spectral range. Ligands such as water molecules and hydroxyl ions behave as nonradiative quenchers of $^{7}D_{3}$ and $^{7}D_{4}$ excited states.17 More generally, in contrast with other lanthanide ions, Eu$^{3+}$ luminescence presents high sensitivity to environment. In particular, the $^{5}D_{0} \rightarrow ^{7}F_{2}$ so-called hypersensitive transition exhibits variations in terms of key characteristics such as intensity, quantum conversion efficiency, or luminescence decay depending on the chemical environment. Actually, although the partially filled $4f$ core shell responsible for the luminescence is expected to be weakly environment sensitive because of the shielding by the surrounding filled $5s$ and $5p$ shells, luminescent characteristics of Eu$^{3+}$ based complexes are known to depend significantly on the ligand-field symmetry and the nature of the ligands such as water or other solvents.18 Recent experiments have shown the direct implication of O, Cl, and H interaction with the europium ion in luminescence properties of EuCl$_3$·6H$_2$O.19 Europium possesses another substable oxidation state Eu$^{2+}$, the electronic configuration of which, for the lowest excited state, is changed to $4f^5$5d$^1$ rather than $4f^6$ for Eu$^{3+}$. This outer 5d subshell is much more sensitive to the surrounding environment sensitive because of the shielding by the surrounding filled $5s$ and $5p$ shells, luminescent characteristics of Eu$^{3+}$ based complexes are known to depend significantly on the ligand-field symmetry and the nature of the ligands such as water or other solvents.18 Recent experiments have shown the direct implication of O, Cl, and H interaction with the europium ion in luminescence properties of EuCl$_3$·6H$_2$O.19 Europium possesses another substable oxidation state Eu$^{2+}$, the electronic configuration of which, for the lowest excited state, is changed to $4f^5$5d$^1$ rather than $4f^6$ for Eu$^{3+}$. This outer 5d subshell is much more sensitive to the surrounding environment sensitive because of the shielding by the surrounding filled $5s$ and $5p$ shells, luminescent characteristics of Eu$^{3+}$ based complexes are known to depend significantly on the ligand-field symmetry and the nature of the ligands such as water or other solvents.18 Recent experiments have shown the direct implication of O, Cl, and H interaction with the europium ion in luminescence properties of EuCl$_3$·6H$_2$O.19 Europium possesses another substable oxidation state Eu$^{2+}$, the electronic configuration of which, for the lowest excited state, is changed to $4f^5$5d$^1$ rather than $4f^6$ for Eu$^{3+}$. This outer 5d subshell is much more sensitive to the surrounding.
environment, such as anions or ligands, than the deep-lying 4f electrons, and always split to several sublevels according to its crystal or ligand field symmetry. Therefore, Eu$^{2+}$ exhibits a broad emission band corresponding to 4f$^5$5d$^1 \rightarrow$ 4f$^6$ in the UV, with often poor quantum efficiencies.$^{20}$ Yet, partial reduction of Eu$^{3+}$ to Eu$^{2+}$ or the reverse oxidation is very likely to occur during material processing.$^{21,22}$ Since this may be detrimental to material performance, an accurate monitoring of the divergent to trivalent composition ratio is necessary. Because those compounds most often include both crystallization water and hydroxide ions, charge neutrality can hardly be exploited from chemical or elemental analysis. Moreover, because of poor absorption cross-section and strong quantum-yield dependence on environment, optical measurements yield large uncertainties.

X-ray photoemission from Eu ions, which is mainly sensitive to the inner shells,$^{23}$ can thus be anticipated to offer an alternative technique to monitor the oxidation-state composition of europium compounds in vacuum, during various material processing operations.

In this article, we report a combined photoluminescence and X-ray photoemission study of the chemical oxidation states of commercial hydrated EuCl$_3$ salts. Photoemission and X-ray absorption spectroscopies reveal the presence of trivalent but also divalent europium species in vacuum. The comparison of ratios between europium 2+ and 3+ species in salts of different origins by means of core level photoemission and X-ray absorption matches the differences observed in their chemical analysis of Eu and Cl species ratio.

2. EXPERIMENTAL METHODS

Experiments were performed using two commercial EuCl$_3$$\cdot$6H$_2$O salts (Acros and Sigma Aldrich, 99.9% purity). Eu content measurements in each sample were carried out using ICP-AES technique on a Liberty Series II (Varian); three replicates for each analysis run, and 443.556 nm detection wavelength. Calibration was performed using a Eu standard solution for ICP-AES and -MS (SCP Science). Eu samples were prepared by solubilization of EuCl$_3$ in water, and each sample was measured in duplicate. Chloride analyses were performed using ionic chromatography (Dionex). Sample preparation for spectroscopic analyses consisted in salt solubilization in dry ethanol, drop casting on gold substrate, and crystallization by evaporation of the solvent at 50 °C under N$_2$ gas flux. The substrates were 100 nm thick layers of gold epitaxially grown on mica. The $22 \times \sqrt{3}$ reconstructed surface of gold(111) is obtained by H$_2$ flame annealing at atmospheric pressure at a temperature around 450 °C. The substrates were prepared immediately before drop casting. Photoluminescence was measured using an Ocean Optics HR2000+ spectrometer. The sample was excited by a continuous-wave laser (Coherent, Cube 405). This compact semiconductor laser system emits up to 100 mW at 405 nm in a TEM00 mode. Photoemission spectroscopy experiments were performed at the French Synchrotron Soleil at the soft X-ray TEMPO beamline.$^{24}$ The spectral resolution of the beamline $\Delta E/ E$ is better than 10 000 over the whole photon energy range 40–1500 eV. A hemispherical electron analyzer Scienta 2002 equipped with a fast delay line detector is used to measure photoelectron spectroscopy spectra. Photoabsorption spectra were determined from sample current measured with picocamperimeters and normalized to gold mesh signal (60% transmission). Measurements were performed at a pressure lower than 3 × 10$^{-10}$ mbars.

3. RESULTS AND DISCUSSION

A chemical dosing of proportion of chlorine and europium ions for 99.9% ultrahigh purity EuCl$_3$$\cdot$6H$_2$O salt has been performed. Surprisingly, as Sigma Aldrich europium salt was presenting an expected proportion of three chlorines for one europium and six surrounding H$_2$O molecules, the dosing of the Eu and Cl species in Acros salt has shown a proportion of 2.85 ± 0.05 chlorines for one europium ion. This composition could be interpreted either in terms of a mixture of 85% of Eu$^{3+}$ species and 15% of Eu$^{2+}$ species or in terms of the presence of alternative negative counterions, such as oxide O$^{2-}$ or hydroxides OH$^-$. The latter seems more probable since, exposed to air, the less stable Eu$^{2+}$ species is known to react with O$_2$ and H$_2$O to be oxidized as Eu$^{3+}$ forming possible hydroxochlorides and oxychlorides. Spectroscopic investigation of the lanthanide salt may give insights in the interpretation of this unexpected lack of chlorine species and determine the equilibrium composition. In addition, a number of ten to 11 H$_2$O molecules surrounding the EuCl$_3$ (2 < x < 3) salt has been found. Lanthanide compounds present highly hygroscopic properties. In the rest of the discussion, hydrated EuCl$_3$ salt will refer to the Acros europium salt and Sigma Aldrich europium salt will be mentioned and used as EuCl$_3$$\cdot$6H$_2$O reference.

In air, the hydrated EuCl$_3$ salt luminescence spectra obtained after excitation at 405 nm through the $^5$D$_0$ level is shown in Figure 1. The spectrum exhibits a relatively sharp line emission structure. As for Eu(III) complexes in aqueous solution, major luminescence of hydrated EuCl$_3$ salt originates from non-degenerate $^7$D$_0$ level.

The number of observed line splittings for the different $^5$D$_0$ → $^7$F$_n$ transitions complies with the allowed 2J + 1 splitting peaks, i.e., at most 1, 3, 5, and 9 peaks for $^5$D$_0$ → $^7$F$_n$ transitions with n = 0, 1, 2, and 4, respectively. This result, associated with relatively sharp line splitting patterns and intense $^5$D$_0$ → $^7$F$_n$ transitions, is in good agreement with a homogeneous single type of Eu$^{3+}$ ion and surrounding noncentrosymmetric arrangement.$^{25}$ Luminescent properties of europium depend on its detailed chemical environment and electronic configuration. Compared to the luminescence of europium trichloride hexahydrate and thermal decomposition products,$^{27}$ the sharp
The asymmetric shape of the measured \(^{5}D_0 \rightarrow \ ^{7}F_2\) transitions well matches with the luminescence of a EuCl\(_3\)-6H\(_2\)O salt, with neither evidence of Eu(OH)\(_3\)Cl signal nor of divalent Eu\(^{2+}\) ion. However, the weak expected contributions from both types of species does not permit to definitely exclude one or the other of the hypotheses evoked to interpret the lack of chloride counterions in this salt.

In order to obtain further insights into electronic and chemical surroundings of europium ion in the salts, we have investigated the electronic properties of europium salt by means of core level photoemission. It is to be noticed that the necessary vacuum operation of this technique is very likely to influence the chemical composition, for instance, the hydration ratios, of the salts. Moreover, synchrotron radiation interaction with reactive materials is well-known to potentially reduce species under beam. Nevertheless, in our experiments, europium chloride appeared to be very stable under synchrotron radiation beam, and no modification of the core level spectra was observed during the acquisition. This does not preclude changes induced during application of vacuum, prior to photoemission measurements. However, the chemical sensitivity of the photoemission spectroscopy allows us to quantitatively establish the in-vacuum salt chemical composition. The Eu 4d core level spectrum is shown in Figure 2: the spectrum exhibits two main components located at higher binding energy attributed to trivalent europium and a second structure located at lower binding energy ascribed to Eu\(^{2+}\).\(^{23}\)

Although for pure EuCl\(_3\) cleaved crystals, the appearance of Eu\(^{2+}\) signature may be ascribed to modified bonding configuration at the surfaces,\(^{23,29,30}\) hydrated EuCl\(_3\) ionic crystal deposited on gold surface is expected to preserve its local crystal configuration even at the surface.\(^{31}\) Actually, the system does not consist in a truncated crystal leading to lower oxidized states at the surface but is an assembly of isolated europium ions carrying their own surrounding counterions (see already published crystal structure of [EuCl\(_3\)(H\(_2\)O)\(_j\)]\(\cdot\)Cl\(_3\)).\(^{22}\) Moreover, we chose a large photoexcitation energy to measure the 4d core level spectrum with a deeper penetration depth in order to minimize potential surface contribution. Additionally, Eu\(^{3+}\) is known to have the tendency to be reduced into Eu\(^{2+}\) species, which could explain the possible presence of divalent Eu after applying vacuum;\(^{33}\) in turn, this possible consequence of material processing procedures represents a well-known inconvieniency in practical applications. In order to further settle the existence of a mixed Eu\(^{3+}/Eu^{2+}\) oxidation states in the bulk, rather than a pure Eu\(^{3+}\) sample potentially exhibiting a Eu\(^{2+}\) contribution due to shake-off processes, a quantitative study of Eu\(^{3+}\) and Eu\(^{2+}\) species has been performed, beyond the lone Eu 4d core level spectrum analysis.

To evaluate the amount of Eu\(^{2+}\) species compared to Eu\(^{3+}\) states, the Eu\(^{2+}\) asymmetric contribution has been fitted with two Voigt doublets. We have thus determined a Eu\(^{2+}\) concentration of 20.7%. In order to clearly determine the location of the Eu\(^{2+}\) oxidation state species and discriminate between surface phenomenon and bulk Eu\(^{2+}\) and Eu\(^{3+}\) mixed ions, we have measured the Eu 3d near edge X-ray absorption fine structure spectrum (NEXAFS). Compared to X-ray photoemission, XAS is much more bulk sensitive, and as other works shown,\(^{34,35}\) the amounts of each oxidation state can be determined from XAS spectra. Figure 3 represents the Eu 3d–4f XAS experimental spectrum measured in total electron yield. The experimental XAS spectrum is analyzed and compared to calculated 3d–4f absorption spectra for Eu\(^{2+}\) and Eu\(^{3+}\) states. The atomic multiplet calculations were performed using the CTM4XAS 5.23 program.\(^{36}\) In the calculation, discrete energy states are broadened by a Lorentzian with a half width at half-maximum of 0.2 eV to take into account core–hole lifetime and by a Gaussian with a half width at half-maximum of 0.6 eV to account for instrumental resolution. A contribution of around 20% of Eu\(^{2+}\) states gives a good matching between calculated and experimental spectra. This ratio of Eu\(^{2+}\) and Eu\(^{3+}\) oxidation states is consistent with the proportion determined with the Eu 4d core level photoemission spectrum. This result shows that the balance of oxidation states is independent of probed sample depth and confirms a homogeneous repartition of the two oxidation states in the bulk of the sample. It is to be noticed that this ratio matches the results of the elemental analysis performed on the used europium salt sample, which gives a Cl/ Eu ratio of 2.85 ± 0.05, which corresponds to about 15% Eu(II) and 85% Eu(III) either as a composition of the pristine material or as a result of a partial Eu(III) reduction as Eu(II) following a vacuum elimination of the counterion other than chloride (e.g., following 4OH\(^{-}\) → O\(_2\) + 2H\(_2\)O + 4e\(^{-}\), if hydroxide counterions are to be considered). As a reference, the absorption spectrum of Sigma Aldrich europium salt exhibiting a EuCl\(_3\)-6H\(_2\)O chemical composition is presented in Figure 3b. The absorption spectrum presents a pure trivalent...
EuCl₃ salt, the oxidation state is supposed to be directly related to the nature of the divalent and trivalent europium counterions. Cl 2p core level has also been measured in order to ascribe the presence of divalent europium in Acros europium salt, after application of the vacuum.

Cl 2p core level has also been measured in order to ascribe the nature of the divalent and trivalent europium counterions. In lanthanide ion salt and, for instance, in the studied hydrated Cl₂p core level spectra of the two different Eu₃⁺ species associated to two and three chlorine ions, respectively. Cl 2p core level spectra are fitted with two components: the first one, located at higher binding energy, presents its maximum at 198.44 eV and represents in intensity area 82.3% of the total Cl 2p peak, whereas the second main component is centered at 199.83 eV, and its intensity area is equal to 17.7% of the total of the Cl 2p intensity. The component located at higher binding energy corresponds to chlorine linked to a more positively charged ion as a trivalent Eu species compared to chlorine situated at lower binding energy related to chlorine linked to a less charged Eu ion as a divalent species. The area intensity is in relatively good agreement, with respect to the uncertainty, with the intensity ratio of 1:3 leads to an expected percentage of 85.2% of the total Cl 2p peak.

The Cl 2p spectrum is somewhat broadened compared to previous work on chlorine adsorption on molecular systems and Cu(100). As on Cu(100), Cl 2p core level presents a sharp and resolved 2-component structure consistent with a single chlorine species. Cl 2p spectrum presents a broadened 2-component structure in the case of chlorine adsorption on molecular film deposited on Cu(100) surface showing more complicated interactions between Cl and molecules. To fit the Cl 2p core level, the use of two components is needed. This result is in clear contradiction with a pure Eu³⁺ composition of the salt.

The core level photoemission and X-ray absorption spectroscopy clearly show the spectroscopic signature of Eu²⁺ species indicating that after application of the vacuum on the deposited europium salt, the Eu²⁺ oxidation state in the adopted structural arrangement is extremely stable. Owing to the high sensitivity to oxygen species such as H₂O and O₂ of Eu²⁺ salts exposed to air, known to spontaneously react and oxidize as Eu³⁺, the Eu²⁺ divalent ions may appear during vacuum preparation of the salt.
sample, depending on the initial chemical composition of the pristine sample.

The matching between the Eu/Cl elementary ratio determined by chemical dosing, core level photoemission and X-ray absorption proves the correlation between initial chemical composition of the pristine salts and of the detected ratio of Eu$^{2+}$ and Eu$^{3+}$ species in the hydrated EuCl$_3$ salt, after application of vacuum. Finally, let us mention that after annealing hydrated EuCl$_3$ salt over 350 °C followed by hydration and reoxidation in air leading to hydration and reoxidation of the Eu$^{2+}$ species into Eu$^{3+}$. The measured Eu 4d core level spectrum, Figure 6, presents the behavior of pure Eu$^{3+}$ species without any noticeable Eu$^{2+}$ contribution. Thus, the possible initial counterion previously lost under vacuum and responsible for the appearance of divalent Eu, is not restored during exposition to air. The persistence of Eu$^{3+}$ character after thermal treatment followed by hydration and oxidation in air may indicate an involved counterion in pristine hydrated EuCl$_3$ salt other than hydroxide, in good agreement with measured photoluminescence (Figure 1).

In conclusion, we have shown that the photoluminescence spectrum of hydrated EuCl$_3$ salt presents a sharp-line structure in good agreement with the $^5D_0 \rightarrow ^7F_n$ ($n = 1$ to 4) radiative transitions. The sharpness and the number of radiative lines fit with a homogeneous single type of Eu$^{3+}$ ion and non-centrosymmetric surroundings. Under vacuum, core level photoemission spectroscopy revealed the presence of two types of europium ions. The asymmetric component of the spectrum is unambiguously attributed to Eu$^{2+}$ ion. Divalent species represents around 15–20% of the total amount of europium ions. This proportion correlates with the elementary composition found by chemical dosing. We have hence undoubtedly identified mixed valence hydrated EuCl$_{2.85}$ compounds, stable under vacuum.

### AUTHOR INFORMATION

Notes

The authors declare no competing financial interest.

### REFERENCES


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**Figure 5.** O 1s photoemission spectrum of hydrated EuCl$_3$ salt deposited on gold substrate ($h\nu = 1100$ eV).

**Figure 6.** Eu 4d photoemission spectrum of hydrated EuCl$_3$ salt deposited on gold substrate ($h\nu = 1100$ eV) after UHV annealing over 350 °C followed by air exposure.


(19) Ahlefeldt, R. L.; Smith, A.; Sellars, M. J. Ligand Isotope Structure of the Optical $F_0 \rightarrow \Gamma D_0$ Transition in EuCl$_3$·6H$_2$O. Phys. Rev. B 2009, 80 (20), 205106.


