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Fractionation of Suwannee River fulvic acid and Aldrich humic acid on α -Al₂O₃: spectroscopic evidence

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ABSTRACT

Sorptive fractionation of Suwannee River Fulvic Acid (SRFA) and Purified Aldrich Humic Acid (PAHA) on α -Al₂O₃ at pH 6 was probed in the supernatant using different spectroscopic techniques. Comparison of dissolved organic carbon (DOC) analysis with UV/Vis spectrophotometric measurements at 254 nm, including Specific UV absorbance (SUVA) calculation, revealed a decrease in chromophoric compounds for the non-sorbed extracts after a 24 h contact time. This fractionation, only observable below a certain ratio between initial number of sites of humic substances and of α -Al₂O₃, seems to indicate a higher fractionation for PAHA. C(1s) near-edge X-ray absorption fine structure spectroscopy (NEXAFS)

confirmed this trend and points to a decrease in phenolic moieties in the supernatant and to an eventual increase in phenolic moieties on the surface. Time-resolved luminescence spectroscopy (TRLS) of Eu(III) as luminescent probe showed a decrease in the ratio between the $^5D_0 \rightarrow ^7F_2$ and $^5D_0 \rightarrow ^7F_1$ transitions for the fractionated organic matter (OM) that is thought to be associated to a lower energy transfer from the OM to Eu(III) due to the loss of polar aromatics. These modifications in the supernatant are a hint for the modification of sorbed humic extracts on the surface.

1 Introduction

Sorption to mineral surfaces and complexation by natural organic matter (NOM) are important processes influencing the cycling of toxic metals in the environment (1, 2). Metal sorption onto oxides and clays and metal binding to NOM, e.g., humic substances (HS) composed mainly of humic and fulvic acids (HA and FA), have been extensively studied. NOM influences metal migration (3), organic contaminant transport (4), and colloids stabilities (5, 6).

However, the specific interactions between metal cations and NOM, both in interactions with a mineral surface, are still matters of debate. The use of linear combinations of binary systems to describe ternary systems was sometimes successful (2, 7-9), but also showed significant deviations under varying HA, FA, and metal concentrations (2, 10-14), or need a semi-empirical refinement of the models (15). More recently, Weng *et al.* proposed the ADAPT model to account for the differences in electrostatic and chemical binding of FA and HA (16). Explanations of these discrepancies can be modifications of NOM conformation and/or composition during sorption, which can be viewed as a modification of the system 'free energy' (2, 16-18).

A preferential adsorption of high molecular weight fractions on oxides or kaolinite occurs in the case of aquatic NOM (19-21), whereas the adsorption of low molecular weight fractions was proposed for Aldrich HA on hematite (22, 23). NOM fractionation on mineral surfaces seems to depend both on the NOM origin and on the mineral. The binding of HS to mineral surface is a combination of (i) ligand exchange reaction, via carboxylic and phenolic functions, (ii) electrostatic interaction between positively charged minerals and negatively charged humic samples, (iii) possible cation bridging, and (iv) hydrophobic interactions between more or less hydrophobic moieties. It has been shown that changes in electrostatic description of HS sorbed on minerals have to be accounted for the description of their sorption (16), and that sorption on minerals (24) and change in ionic strength (25) do influence the hydrophobicity of humic sample.

Various techniques have been used to identify NOM fractionation such as chromatographic techniques based on size separation. These techniques are somehow biased by additional fractionation where UV-Visible detection alone is not totally appropriate (26-28). Despite all these efforts, the identification at the molecular scale is still difficult (29, 30). There is a need of molecular level understanding of the sorptive fractionation processes.

HS have being used as a surrogate to estimate NOM properties. It was proposed that HS aggregation results from self-assembled molecules (31-35). The dimension of the smallest observable object is about some nanometres composing larger aggregates of some hundreds of nanometres (34). Fulvic acids only form aggregates at very high concentration (35).

In this work, we combine results of four different techniques to obtain complementary views on the HS sorptive fractionation on α -Al₂O₃ carrying aluminol sites: UV-visible absorption (UV-Vis), total organic carbon analysis (TOC), C(1s) near edge X-ray fine structure spectroscopy (NEXAFS) and time-resolved luminescence spectroscopy (TRLS). UV-visible absorption spectra of HS are supposed to give little structural information about the functionality due to the lack of detailed and characteristic bands (36). Its use for determining HS concentration remaining in solution after sorption is unreliable (20). However, a decomposition of UV-Vis spectra to monitor the properties of NOM has been proposed (37). Considering that HS contains ‘building block molecules’ that differ by their substituents (31, 32), one can see HS, in a first approximation, as a distribution of chromophores (38). TOC analysis is the most appropriate to determine accurately the concentration of humic substances (20). Specific UV-Vis Absorbance (SUVA=A_{254 nm}/TOC), is a useful approach for probing aromaticity of humic samples (39). Synchrotron-based C(1s) NEXAFS have being more recently applied to NOM study and provided semi-quantitative data on the carbon types or modification in NOM samples (40-43). TRLS have being used to study interactions of lanthanides (Ln) and actinides (An) with different organic media and to probe their chemical environment in HS (44-47). Both the spectra and luminescence bi-exponential decays of Ln-HS complexes are remarkable and permit to distinguish humic complexation from a more classical organic complexation (48, 49). Hence, using europium (III) as a luminescence probe allows obtaining a complementary view of fractionation changes in the metal binding environment.

2 Materials and Methods

2.1 Sorption Experiments.

Millipore deionized water (Alpha-Q, 18.2 M Ω cm) was used. SRFA is used as received from the International Humic Substance Society; its total site concentration Q_{SRFA}=7.4 mmol/g is taken from ref 50. Purified Aldrich Humic Acid (PAHA) was treated following the procedure describing in ref 51, with Q_{PAHA}=5.34 mmol/g (50). Q_{HS} were obtained in the framework of the NICA-Donnan model, which was applied afterward to determine the actual functionality under our conditions. Alpha alumina (Interchim, Montluçon, France) has a total site capacity for exchange of anions Q _{α -Al₂O₃}=24 μ mol/g, and the surface properties used hereafter to determine the functionality under our conditions are in ref 52. Carbonates were removed from α -Al₂O₃ surface as described in (52); the isoelectric point pH_{I_{EP}} 9.5 was determined by zeta potential measurement (data not shown). Europium (III) stock solution was obtained from the dissolution of Eu₂O₃ (Johnson Matthey, 99.99%) in HClO₄ 70% (Merck, GR for analysis).

Sorption studies were performed at pH 6.1 \pm 0.1 because of the high variation of sorption results awaited for this kind of mineral phase vs. HS concentration (16). All the suspensions, solutions, and dilutions were done by weighing. The ionic strength *I* was fixed with NaClO₄ to 0.1 mol/kg_{H₂O} (*m*) and pH was adjusted using freshly prepared NaOH and HClO₄. The pH measurements were done using a combined-glass electrode (Radiometer Analytical XC111) calibrated for its linear response with a 0.01 *m* HClO₄ solution, an equimolar 0.02 *m* NaH₂PO₄/Na₂HPO₄ solution, and an equimolar 0.02 *m* Na₂CO₃/NaHCO₃ solution, all containing NaClO₄ to keep [Na⁺] constant at 0.1 *m* (pH=2, 6.785, and 9.9 respectively). The electrode filling solution was modified with NaClO₄ 0.1 *m*, NaCl 10⁻² *m* to prevent KClO₄ precipitation in the frit of the electrode.

Batch sorption experiments were conducted at room temperature (22 ± 1 °C). Weighed aliquots of α -Al₂O₃, NaClO₄, and humic solution, were suspended in milli-Q water. The detailed results are presented in Table S1 in supplemental information (SI). The suspensions were shaken for 24 hours and pH was measured afterwards until a constant value was reached (usually 2 to 4 hours). This contact time is sufficient to obtain a quasi-equilibrium. Some works seem to indicate that exchange between the HS, or NOM sub-fractions, is still occurring up to 900 days (22, 53).

The solid phase was separated via a pre-centrifugation step performed in 30 mL Nalgene 3137-0030 centrifuge vials (50,000 g for 1 hour); the supernatant was then ultra-centrifuged in order to settle remaining colloidal particles in 10 mL Beckman 355603 vials (246,960 g for 1 hour). Modifications of the HS composition induced by this centrifugation steps can be ruled out as for the PAHA/ α -Fe₂O₃ system (30). No filtration step was done before absorbance measurement and particular caution was taken during the sampling to avoid resuspended colloidal particles. Supernatant spectra were consecutively measured using a Shimadzu UV-3150 spectrophotometer and TOC in solution was determined with a Shimadzu 5000 TOC analyzer after acidification and purging with argon. The uncertainties of the determinations were obtained from the regression analyses of the calibrations curves built from weighed aliquots of HS.

2.2 *C(1s) NEXAFS*

Carbon K-edge NEXAFS spectra were measured at the Scanning Transmission X-ray Microscopy (STXM) beamline X1A1 (NLSL), operated by the State University of New York at Stony Brook, USA. The principle of the methods is described elsewhere (54). STXM sample preparation was performed by drying 1 μ L of HS solution for supernatant, and after deposition of the HS-covered α -Al₂O₃, on a Si₃N₄ window (100 nm thick). For the supernatant, depending on the organic concentration the spectra was either (i) extracted from images taken at different energies across the absorption edge (stacks method) after aligning them using cross-correlation (55), or (ii) measured by the point spectra procedure (43). The two methods gave comparable results on the same samples. The spectra of the HS-covered α -Al₂O₃ samples had a very low signal to noise ratio. Energy calibration of the spherical grating monochromator was achieved by using the photon energy of the 290.74 eV CO₂ gas adsorption band (56). For comparison, all NEXAFS spectra were baseline corrected and normalized to 1 at 295 eV.

2.3 *Analysis of UV/Vis spectra.*

UV/Vis spectra of HS show a classical strong steady increase with decreasing wavelength. In HS, the majority of the chromophores that absorbs in UV are aromatic groups with various degrees of substitution (37, 57). In analogy to benzene, a simplified decomposition approach to allow a better understanding of the chromophoric group functionalities was proposed (37). The 200-400 nm region represents three simulated bands; a very intense local excitation band (LE) around 180 nm, a benzenoid band (Bz) around 203 nm due to vibrational perturbations in the π -electron system, and a low intensity electron transfer band (ET) around 253 nm due to strong quantum-mechanical prohibition but strongly affected by the presence of polar groups. The spectra are then fitted with the sum of the three Gaussians using a non-linear regression fitting procedure.

2.4 TRLS

Europium(III) was used to probe the luminescence properties in the supernatant. Eu(III) shows a luminescence spectrum corresponding to the $^5D_0 \rightarrow ^7F_j$ transition manifold (58). The luminescence properties of Eu(III) in the supernatant were compared to those in the original humic samples.

The excitation laser beam was generated by a 355 nm tripled output of a Continuum Nd-YAG laser, coupled to an optical parametric oscillator system (Panther II, Continuum, USA). The wavelength was tuned to 394.6 nm, providing about 1.6 mJ of energy in a 5 ns pulse with a repetition rate of 10 Hz. The luminescence signal is collected at 90° and focused into an Andor spectrometer (550-650 nm, resolution 0.2 nm). The signal is collected during a gate width of 300 μs, at a gate delay of 10 μs after excitation by the laser flash. To increase the signal to noise ratio, every spectrum was accumulated 1,000 times. Emission spectra were recorded between using a CCD camera cooled at -15°C. Only $^5D_0 \rightarrow ^7F_0$ (around 580 nm), $^5D_0 \rightarrow ^7F_1$ (around 593 nm), and $^5D_0 \rightarrow ^7F_2$ (around 615 nm) bands can be observed. Before TRLS measurement the pH was adjusted to 4 to limit the signal from EuOH^{2+} and EuCO_3^+ at pH 6.

3 Results and Discussion

3.1 UV/VIS, TOC, and SUVA analyses.

In Figure 1a, the distribution coefficients:

$$R_d \text{ (mL/g)} = \frac{[\text{HS}]_{\text{sorbed}} \text{ (mol/g)}}{[\text{HS}]_{\text{solution}} \text{ mol/mL}} = \frac{[\text{HS}]_{\text{sorbed}}}{[\text{HS}]_{\text{initial}} - [\text{HS}]_{\text{sorbed}}} \times \frac{V \text{ (mL)}}{m_{\alpha\text{-Al}_2\text{O}_3} \text{ (g)}} \quad (1),$$

of sorbed HS onto $\alpha\text{-Al}_2\text{O}_3$ are plotted vs. the ratio R_{HS} of initial available sites in HS and $\alpha\text{-Al}_2\text{O}_3$ defined as:

$$R_{\text{HS}} \left(\text{mol} \cdot \text{g}_{\text{HS}}^{-1} / \text{mol} \cdot \text{g}_{\alpha\text{-Al}_2\text{O}_3}^{-1} \right) = \frac{Q_{\text{HS}} \text{ (mol} \cdot \text{g}_{\text{HS}}^{-1})}{Q_{\alpha\text{-Al}_2\text{O}_3} \text{ (mol} \cdot \text{g}_{\alpha\text{-Al}_2\text{O}_3}^{-1})} \times \frac{C_{\text{HS}} \text{ (mg}_{\text{HS}} \cdot \text{kg}_{\text{H}_2\text{O}}^{-1})}{C_{\alpha\text{-Al}_2\text{O}_3} \text{ (mg}_{\alpha\text{-Al}_2\text{O}_3} \cdot \text{kg}_{\text{H}_2\text{O}}^{-1})} \quad (2).$$

The R_d notation reflects that under our conditions the final hypothetical equilibrium conditions is not reached and that reversibility is not ascertained (59). The concentrations of HS are either determined from the calibration in UV-Vis, which is known to be biased for sorbed samples (26), and in TOC. The adsorption isotherms are reported in SI (Figure S1a). The sorption is constant between 1 and 48 hours (data not shown) and there is no evolution of the SUVA vs. time for $R_{\text{PAHA}} = 10.4$ (Figure S1b). The latter figure confirmed that fast HS sorption was occurring on aluminum oxides as it had already been noticed for iron oxides (20).

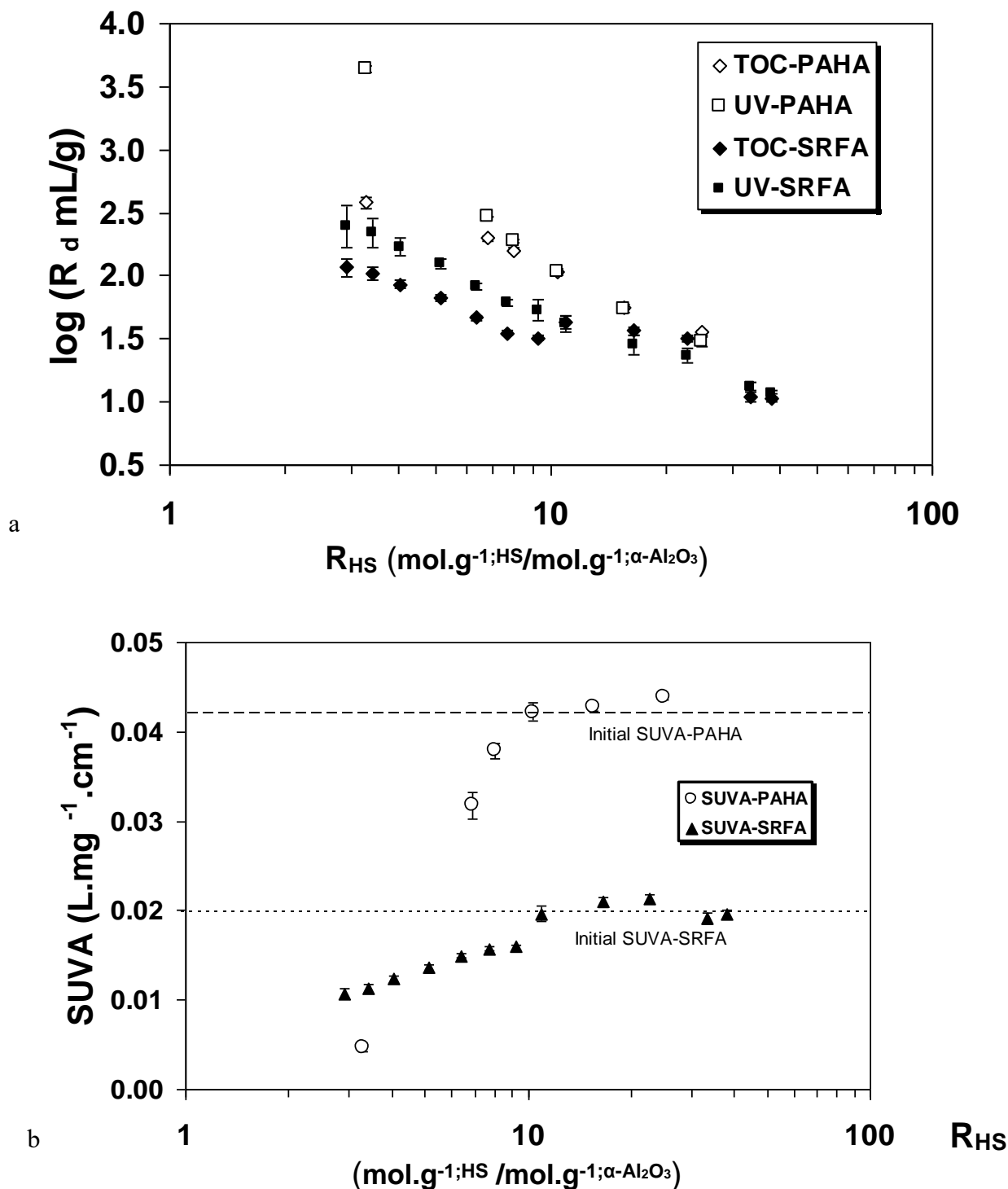


Figure 1. Evolution of R_d (mL/g) for SRFA (closed symbols) and PAHA (open symbols) adsorbed on $\alpha\text{-Al}_2\text{O}_3$ determined in UV-Vis (square) and TOC (diamonds) as a function of R_{HS} (a), and SUVA from the same experiments for PAHA (O) and SRFA (\blacktriangle) (b), pH 6, $I=0.1 \text{ m}$.

R_{HS} can be viewed as a dimensionless number representing the amount of humic sites present on the mineral surface that bears its own functionality. It could permit to directly compare different sets of experiments done under different concentration conditions, different functionalities (humic and fulvic, hydroxides and oxides) or different surface properties (silica, iron or aluminum oxides). It could also be

used to evidence an overcompensation of humic functionality compared to the available surface sites as in ref 60. Application to clays could be more difficult.

In Figure 1a, $R_d(\text{PAHA}) > R_d(\text{SRFA})$ as in ref 61. When $R_{\text{HS}} < 10$, R_d obtained by UV/Vis are higher than determined by TOC. This can also be seen on the normalized absorption spectra (Figure S2 and S3 of SI). This was observed for NOM sorption on oxides (20, 37). The SUVA in Figure 1b monitors the same trend. Initial SUVA for PAHA is more important than SRFA due to a more important specific absorbance. This representation results in a change in slope around $R_{\text{HS}} \approx 10$ corresponding to the point where UV/Vis and TOC do not correspond (Figure 1a). The greater decrease in SUVA for PAHA indicates a greater fractionation of aromatic moieties. The fractionation is thus only observable when $R_{\text{HS}} \leq 10$ but is occurring under all R_{HS} conditions. This loss of polar aromatic in the supernatant must imply an enrichment of aromatic moieties at the surface.

3.2 NEXAFS analyses and UV/Vis decompositions.

C(1s) NEXAFS measurements are presented on Figure 2a,b for HS supernatants. Only the HS-covered $\alpha\text{-Al}_2\text{O}_3$ sample having $R_{\text{HS}} \leq 10.9$ were analyzed. These spectra revealed very low signal to noise ratio compared to the quality of data for clays (62). The most plausible reason of these low signal is a more homogeneous covering of oxides and the occurrence of more dense and ‘strong spots’, on the Callovo-Oxfordian argillite. Only data for the most concentrated sample, i.e., $R_{\text{SRFA}} = 10.9$, could be analyzed (Figure S4 and Figure 2a).

The detailed band assignment of FA via correlation with ^1H and ^{13}C NMR is referred to ref 40, knowing that uncertainties exist when transferring the band assignment to other organic systems both concerning the absolute intensity response of electron transitions and assignment of functional groups due to overlapping resonance ranges of electron transitions (63).

The spectra of the initial HS show three main peaks at 285.2, 286.6, and 288.6 eV corresponding to the C(1s) $\rightarrow \pi^*$ transitions. The 285.2 eV band corresponds to aromatic C–C or C–H bonds (64). The 286.6 eV transition can be assigned to aromatic C–O (phenols) or C–X (halogen) bonds (64). The most intense band around 288.6 eV is related to the oscillation strength of the C=O bonds of carboxylic groups (65, 66). This latter peak is clearly asymmetric providing evidence that both the 289.5 eV transition, assigned to the O-alkyl group of alcohols/carbohydrates, and the mixed Rydberg valence band around 287.4 eV, generally assigned to aliphatic groups, are masked (67). The very small 290.2 eV peak can be assigned to remaining carbonates (68).

A steady decrease in aromatic and phenol-type groups occurs when $R_{\text{HS}} \leq 6$. This effect is pronounced for $R_{\text{SRFA}} = 2.9$. The same trend is observed for PAHA, but a strong fractionation of aromatics is still visible at $R_{\text{PAHA}} = 6.9$, whereas at $R_{\text{SRFA}} = 10.9$ the spectrum looks already similar to initial SRFA.

The C(1s) NEXAFS spectra confirm the SUVA results and give a direct spectroscopic proof that the observed difference between TOC and UV/Vis can be likely both attributed to a reduction of aromatic functional group (peak decrease at 285 eV) and chromophoric groups with polar aromatic functionalities, i.e. mostly phenolic.

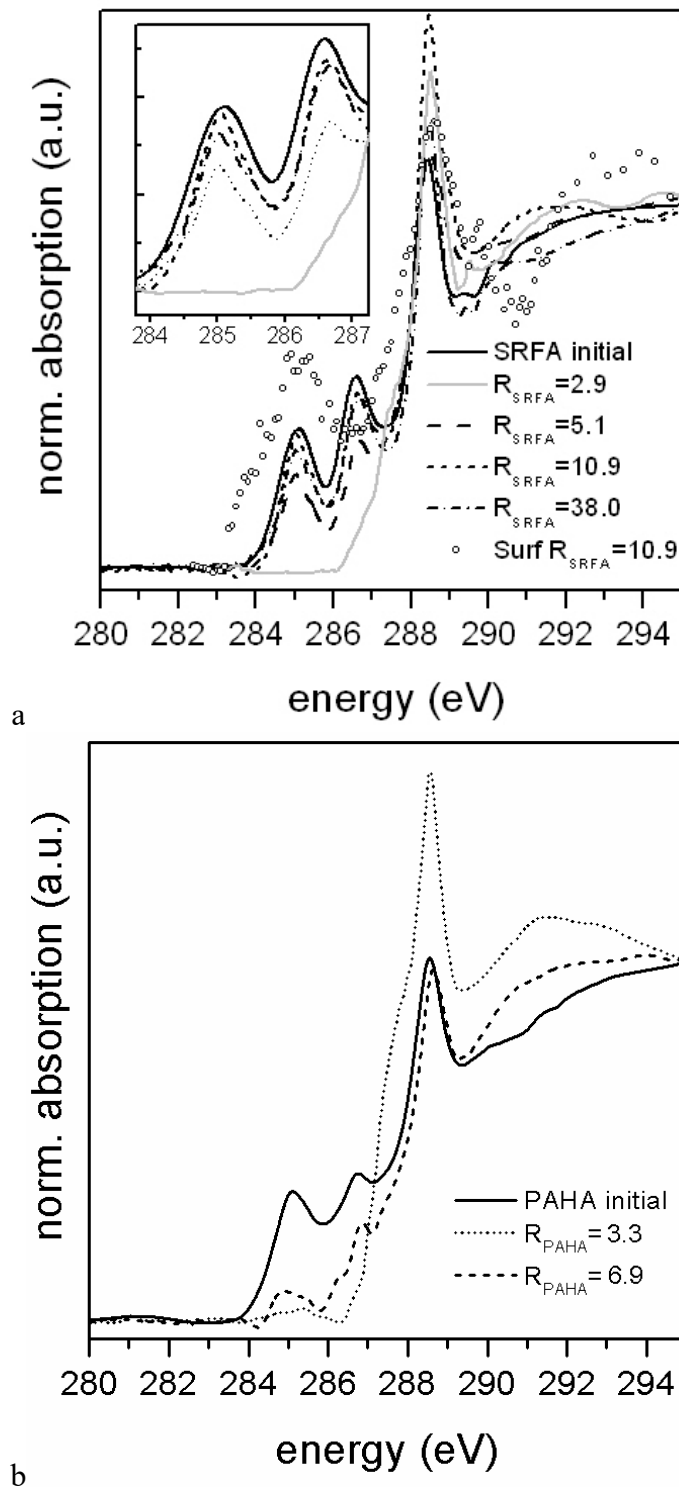


Figure 2. C(1s) NEXAFS spectra obtained before sorption (plain line), for the supernatant after adsorption on α -Al₂O₃ for different ratio R_{HS} for SRFA (grey $R_{SRFA} = 2.9$, dashed $R_{SRFA} = 5.1$, dotted $R_{SRFA} = 10.9$, dash-dot $R_{SRFA} = 38$), and smoothed spectrum for HS-covered α -Al₂O₃ at $R_{SRFA} = 10.9$ (original spectrum in Figure S4) (a), and supernatant PAHA (dashed $R_{PAHA} = 6.9$, dotted $R_{PAHA} = 3.3$) (b), pH 6.1, $I=0.1$ m.

Even if the spectrum of HS-covered α -Al₂O₃ sample at $R_{SRFA} = 10.9$ was exhibiting a low signal to noise ratio (Figure S4), two peaks can be distinguished, around 285 eV and 289 eV. A Savitzky-Golay smoothing (5 points, Origin 8.0) was applied to the spectrum giving the circles on Figure 2a. Other solid samples could not even be smoothed to extract reliable signal. From this smoothed spectrum, it seems that

the contribution around 285.2 eV (aromatic carbons) is higher than both the initial SRFA and the corresponding fractionated SRFA. The contribution around 288.6 eV (carboxylic carbons) is slightly higher than the original sample and comparable to the fractionated sample. It could be inferred that enrichment of aromatic moieties on the mineral surface is occurring, but considering the extensive smoothing, these data should be taken with great care.

UV/Vis spectra were decomposed following ref 37 (see equation S2, Figure S2 and S3, Table S1 and S2 in SI), giving $A_{0,LE}$, $A_{0,Bz}$, and $A_{0,ET}$ the maximum absorbance of the LE, Bz, and ET bands, respectively; Δ_{LE} , Δ_{Bz} , and Δ_{ET} the full widths at mid height of the decomposed bands; $E_{0,LE}$, $E_{0,Bz}$, and $E_{0,ET}$ are fixed at 180 nm, 203 nm, and 253 nm, respectively (37). The $A_{0,Et}/A_{0,Bz}$ ratios, which monitors the proportion of aromatics that contains polar groups, are plotted vs. R_{HS} (Figure 3). In the region of comparable UV/Vis and TOC ($R_{HS}>10$), $A_{0,ET}/A_{0,Bz}\approx 0.6$ for SRFA and $A_{0,ET}/A_{0,Bz}\approx 1.2$ for PAHA indicating that PAHA contains more polar aromatic groups (37). For $R_{SRFA}<10$, $A_{0,Et}/A_{0,Bz}$ is slightly but significantly decreasing to values around 0.5. This decrease is more pronounced for PAHA and is in line with SUVA evolution. This is an indication of a stronger fractionation of PAHA as seen in SUVA and STXM. This is also due to the increasing importance of the LE band in the fit (Table S1 and S2 in SI), which decreases the relative importance of $A_{0,Bz}$. The decrease of $A_{0,ET}/A_{0,Bz}$ and of Δ_{ET}/Δ_{Bz} (Figure S5 of the SI), and the increase of $A_{0,LE}$ (Table S1 in SI) indicate preferential removal of HS molecules that contain higher than average concentrations of aromatic units activated with oxygen-containing functional groups in the supernatant, i.e. phenolic, benzoic, aromatic esters and carbonyl groups (37, 57). This is in line with the loss of phenolic groups in NEXAFS, and with the PAHA/ α -Fe₂O₃ system (30). Another interpretation might be a hypochromism effect (69), which would imply a reversible self-association of molecules. A combination of the two phenomena is more likely.

The correlation between NEXAFS and $A_{0,Et}/A_{0,Bz}$ presented in Figure 3 confirms that the changes in slopes are occurring for $R_{HS}\approx 10$; we cannot fully ascertain the trend for PAHA. The phenomena underlying UV-Vis and NEXAFS are not directly comparable but the trends are strikingly similar. Direct correlations are useless for the moment being as when $R_{HS}\leq 3.5$ no aromatics can be detected in STXM.

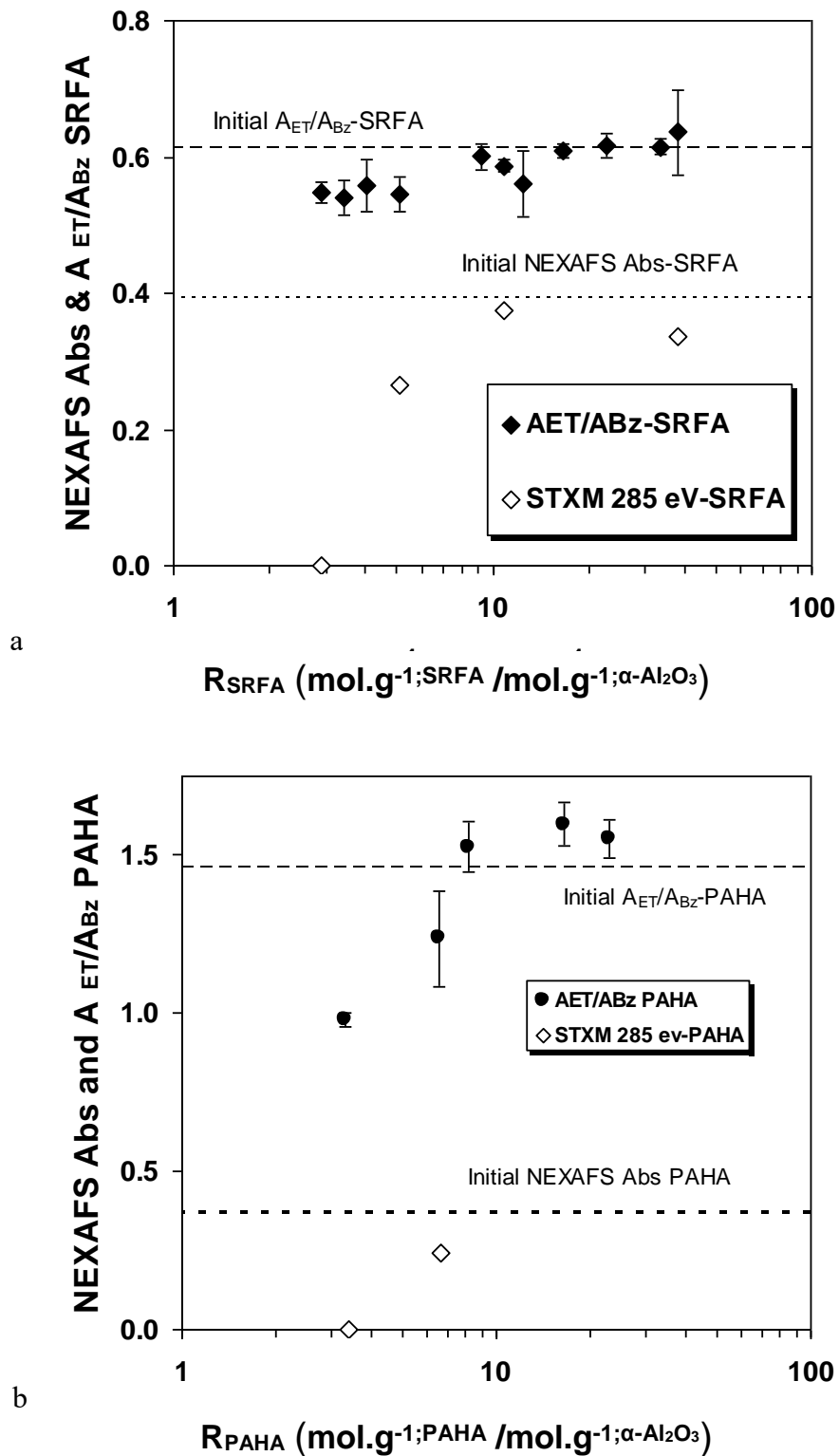


Figure 3. The $A_{0,ET}/A_{0,Bz}$ and NEXAFS peak height for SRFA (a) and PAHA (b) for different ratios of available sites, pH 6,1, $I=0.1$ m. The original data are reported in Figure S2 and S3, and in Table S1 and S2 of the SI

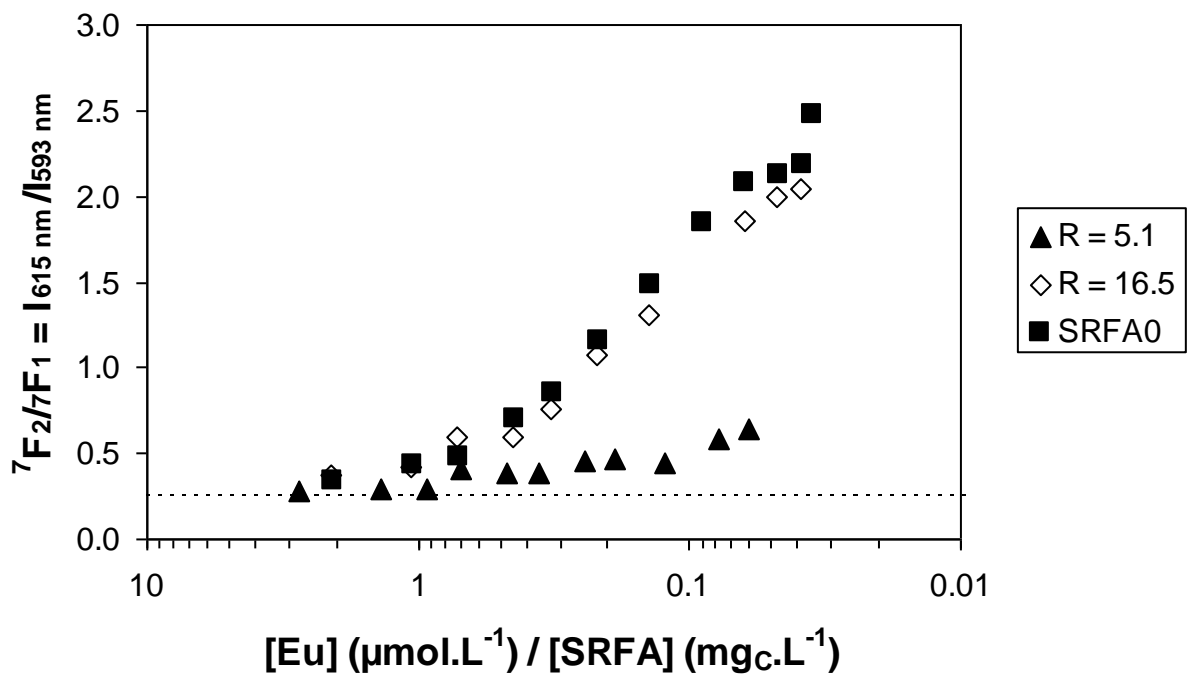
3.3 TRLS.

The ratio between ${}^5D_0 \rightarrow {}^7F_2$ and ${}^5D_0 \rightarrow {}^7F_1$ bands (under our conditions ${}^7F_2/{}^7F_1 = I_{615 \text{ nm}}/I_{593 \text{ nm}}$) is an indication of the chemical environment of Eu(III); the ${}^5D_0 \rightarrow {}^7F_2$ transition is hypersensitive to changes in the chemical environment. In Figure 4, ${}^7F_2/{}^7F_1$ are plotted vs. $[\text{Eu(III)}]/\text{TOC}$ ($\mu\text{mol}_{\text{Eu}}/\text{mg}_C$) for experiments with supernatants and the original HS. For aqueous Eu^{3+} , ${}^7F_2/{}^7F_1=0.25$ (dash-dotted line) (44, 49), whereas for a complete binding of Eu(III) to HS, *i.e.* $[\text{Eu}]/\text{TOC}=0.03$, ${}^7F_2/{}^7F_1 \approx 2.5$ as in ref 45. The increase is slightly steeper in the case of PAHA₀ when compared to the available functionality ($Q_{\text{PAHA}} = 1.75 \text{ mmol/g}$, and $Q_{\text{SRFA}} = 3.9 \text{ mmol/g}$ at pH 4, see figure S6 in SI), indicating that PAHA seems to have slightly stronger complexation capacity for lanthanides-actinides (III) than SRFA, which was also shown for other samples (70, 71).

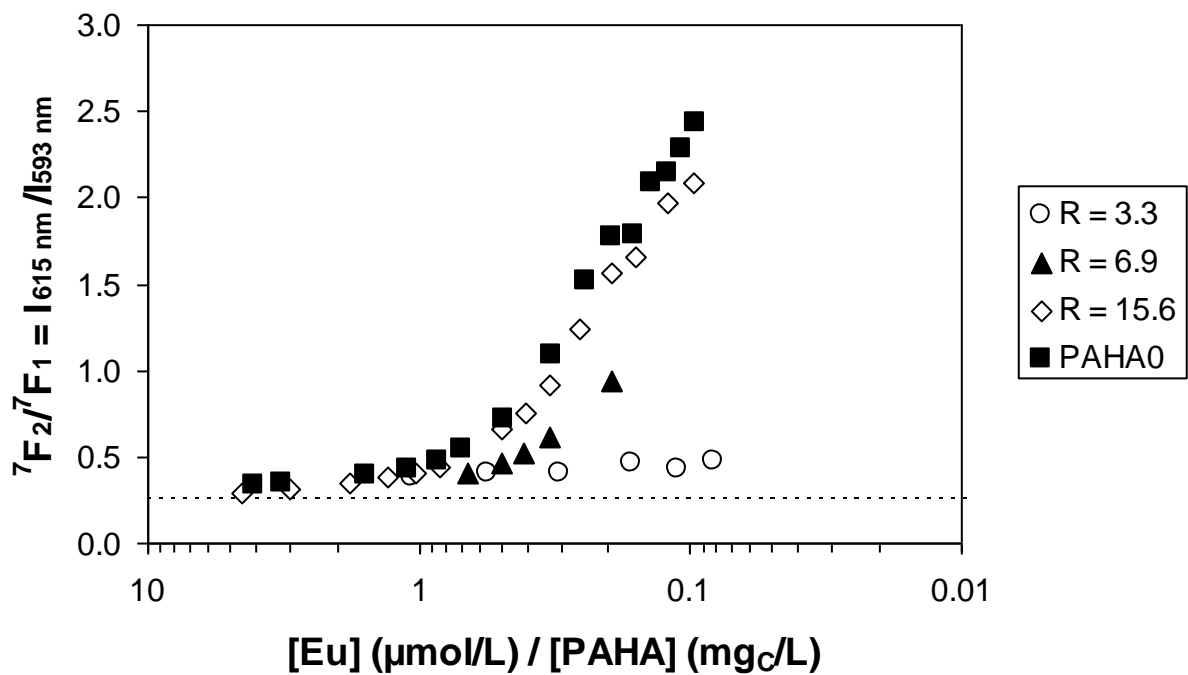
For the supernatant, when $R_{\text{HS}} \geq 10$, the ${}^7F_2/{}^7F_1$ evolution is close to original samples; conversely, when $R_{\text{HS}} \leq 10$, ${}^7F_2/{}^7F_1$ decreases for comparable $[\text{Eu}]/\text{TOC}$ ratio. One can remark that ${}^7F_2/{}^7F_1$ evolutions are similar for $R_{\text{SRFA}}=16.5$ and $R_{\text{PAHA}}=15.9$, and that the increase at $R_{\text{SRFA}}=5.1$ is slightly lower than for $R_{\text{PAHA}}=6.9$.

Energy transfer from the excited state of the complexed ligand to the 5D_0 level of Eu(III), leading to a luminescence enhancement, is generally taking place over the triplet levels in the chromophores via the ‘antenna effect’ (72). It could be inferred that this energy transfer is less efficient for fractionated organic matter due to the loss of polar aromatic moieties. For the supernatant at $R_{\text{SRFA}}=5.1$, ${}^7F_2/{}^7F_1 \approx 1$ when $[\text{Eu}]/\text{TOC} = 0.03 \mu\text{mol}_{\text{Eu}} \cdot \text{g}_C^{-1}$ as for complete complexation of glycolic acid (49); for $R_{\text{PAHA}}=3.3$, when $[\text{Eu}]/\text{TOC} = 0.08 \mu\text{mol}_{\text{Eu}} \cdot \text{g}_C^{-1}$, ${}^7F_2/{}^7F_1 \approx 0.5$ indicating a more pronounced decrease in energy transfer, in agreement with STXM, $A_{0,\text{ET}}/A_{0,\text{Bz}}$, and SUVA.

The evolutions of ${}^7F_2/{}^7F_1$ vs. $[\text{Eu}]/\text{TOC}$ on Figure 4 look like complexation isotherms. It was not possible to perform acid-base titrations of HS in the supernatant because of their low concentration. Furthermore, it was not possible to attain the plateau indicating the complete complexation, it is then impossible to calculate any interaction parameter related to a number of sites. Nonetheless, it seems that the supernatants have a lesser degree of interaction with Eu(III) than original samples as the half reaction point is shifted to higher carbon concentrations with decreasing R_{HS} values. It seems then that humic and fulvic acids in the supernatant have a lesser affinity with Eu^{3+} . It is up to now not possible to conclude on this matter as long as the functionality of non-sorbed fulvic fraction cannot be quantified. Should non-sorbed sample be different from the original samples, then sorbed samples should also have different properties than the original humic substances due to the relative enrichment of both polar aromatic groups at the surface, as postulated from STXM results, and modification of functionality (4).



a



b

Figure 4. Evolution of the ${}^7F_2/{}^7F_1$ peak height ratio as a function of $[\text{Eu(III)}]/\text{TOC}$ ($\mu\text{mol}_{\text{Eu}}.\text{L}^{-1}/\text{mg}_C.\text{L}^{-1}$) for SRFA (a) and PAHA (b), and for different values of R_{HS} , pH 4.0, $I=0.1 \text{ m}$: dash-dotted line is the ${}^7F_2/{}^7F_1$ ratio of Eu^{3+} at pH 4.

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The supplemental information gives: the sorption isotherm expressed as free HS concentration vs. total HS concentration; the kinetic evolution of SUVA and sorption % for $R_{\text{PAHA}} = 10.4$; the normalized UV-Vis absorbance spectra of initial SRFA, PAHA, and supernatants after sorption on $\alpha\text{-Al}_2\text{O}_3$ at pH 6.1 for different values of available sites ratio R_{HS} ; the initial spectrum of HS-covered $\alpha\text{-Al}_2\text{O}_3$ sample at $R_{\text{SRFA}} = 10.9$; illustration of the parallel evolution of the intensities of the ET and Bz bands, $A_{0,\text{ET}}/A_{0,\text{Bz}}$, and of the width ratio $\Delta_{\text{ET}}/\Delta_{\text{Bz}}$, and ${}^7\text{F}_2/{}^7\text{F}_1$ ratios vs. $[\text{Eu}]/[\text{HS}]$ ($\text{mol}_{\text{Eu}} \cdot \text{mol}^{-1}_{\text{site}}$) for original HS samples. The experimental data including pH, actual concentration of $\alpha\text{-Al}_2\text{O}_3$, SRFA, and PAHA, distribution coefficients are given as well as results of the fits and the fitting parameters and fitting uncertainties. Illustration of the parallel evolution of the intensities of the ET and Bz bands, $A_{0,\text{ET}}/A_{0,\text{Bz}}$, and of the width ratio $\Delta_{\text{ET}}/\Delta_{\text{Bz}}$ are also given. This information is available free of charge via the Internet at <http://pubs.acs.org>.

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